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Article

Lactonization of α -Ferrocenyl Ketocarboxylic Acids via Nucleophilic Attack of Carbonyl Oxygen

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Abstract: The effects of the ferrocenyl moiety to enhance the nucleophilicity of the carbonyl group, situated at its adjacent position, have been explored in a series of α -ferrocenyl ketocarboxylic acids. In the presence of trifluoroacetic anhydride, 3-ferrocenylpropionic acid and 4-ferrocenylbutyric acid gave 5-ferrocenyl-4-trifluoroacetyl-2(3*H*)-furanone and 6-ferrocenyl-5-trifluoroacetyl-3,4-dihydropyran-2-one, respectively. Under similar reaction conditions, 2-ferrocenylcarbonylbenzoic acid, a keto carboxylic acid without a β -hydrogen, gave a dimerized lactone, 3,3'-diferrocenyl-3,3'-diphtalide possibly due to radical coupling. The nucleophilic attack of carbonyl oxygen, activated by the ferrocenyl moiety, to the carboxylic carbon, is assumed to be the crucial mechanistic step in forming these lactones. When the carbonyl group was reduced to an alcohol to break its conjugation with the ferrocenyl moiety, saturated lactones were isolated after the acidic workup. These results indicate that the α -ferrocenyl carbinols readily undergo solvolysis under acidic conditions, giving ferrocenylcarbenium ions, which are attacked by the carboxy oxygen to give lactones.

Keywords: ferrocenylcarbenium ion; nucleophilicity of carbonyl oxygen; valence tautomerization; dimerization; lactone; ferrocenyl ketocarboxylic acid; radical coupling; trifluoroacetic anhydride

1. Introduction

After the discovery of ferrocene in 1951 [1], research has been focused on understanding the role of the ferrocenyl moiety in stabilizing the positive charge at its alpha position. The stability of α -ferrocenyl carbenium ions has been established based on the enhanced rates of solvolysis of ferrocenyl substrates with the S_N1 mechanism. For example, ferrocenylmethyl acetate solvolysis in 80% aqueous acetone at 30 °C is 6.7 times faster than triphenylmethyl acetate [2]. Various research groups have studied the existence of α -ferrocenylcarbenium ions using the chemical shift values of protons and carbons at the carbocationic centers in NMR [3–8]. The ability of the ferrocenyl moiety to stabilize the positive charge at its α -position has generated further interest in understanding the basic nature of the oxygen atom in acylferrocenes. Protonation of oxygen in acylferrocenes influences their reactivity, particularly in electrophilic aromatic substitution reactions. The electronic effects induced by protonation at the oxygen atom can also significantly alter the electron density of ferrocene moiety, impacting the overall electronic properties of the molecule. Both experimental [9–11] and computational [12] studies have shown that carbonyl oxygen is the most basic site in acyl- or formylferrocene. The higher basicity of the carbonyl oxygen in acylferrocenes than that of aryl ketones and aldehydes has been attributed to the cooperative interaction of the proton with both oxygen and the iron center [9].

Understanding the basic nature of the carbonyl oxygen lying alpha to the ferrocenyl moiety, we wished to explore the nucleophilicity of carbonyl oxygen in acylferrocenes, since nucleophilicity correlates with basicity [13,14]. We realized that the α -ferrocenyl ketocarboxylic acid could offer both electrophilic and nucleophilic centers if the carboxy carbon is converted into a better electrophilic center. Moreover, separating ketone and carboxylic acid functions with 2 or 3 carbon atoms can give

thermodynamically stable 5- or 6-membered rings by intramolecular nucleophilic attack. We were also interested in studying the effect of breaking conjugation between the carbonyl group and the cyclopentadienyl ring on the reaction outcomes. The study on the nucleophilicity of carbonyl oxygen in acylferrocenes can help understand the competition between cyclopentadienyl carbons and carbonyl oxygen if the molecule contains an electrophilic center. Herein, we report two pathways of lactone formation: 1. intramolecular nucleophilic attack of carbonyl oxygen to the carboxy carbon in the presence of trifluoroacetic anhydride; 2. nucleophilic substitution of the hydroxy group in α -hydroxy carboxylic acids by carboxylic oxygen under acidic conditions.

2. Materials and Methods

2.1. General Procedures

All experiments were performed under nitrogen using standard Schlenk line techniques. All reagents and solvents were purchased from commercial suppliers. Ferrocene, succinic anhydride, magnesium sulfate (Alfa Aesar, Ward Hill, MA), glutaric anhydride (TCI America, Portland, OR), sodium borohydride, phthalic anhydride, aluminum chloride (Acros Organic, Waltham, MA), and trifluoroacetic anhydride (Oakwood Chemicals, SC) were used as received.

^1H NMR (400 MHz), ^{19}F NMR (376 MHz), and ^{13}C NMR (100.6 MHz) spectra were recorded on a JEOL-400 ESZ spectrometer at room temperature. The spectra were referenced to the residual protonated solvent (^1H) or the solvent signal (^{13}C). ^{19}F NMR spectra were referenced to the external reference of the spectrometer. IR spectra were recorded on a Bruker Alpha-E FTIR spectrometer using a diamond crystal ATR accessory in a range between 400 and 4000 cm^{-1} . ESI-MS were recorded on Agilent 6470 LC/TQ using 1260 infinity II HPLC system. Melting points were taken on DynalonTM AfonTM DMP100 Melting Point Device and were uncorrected.

X-ray diffraction data were measured at $T = 90$ K on a Bruker Kappa Apex-II diffractometer equipped with a microfocus $\text{CuK}\alpha$ source ($\lambda=1.54184$ Å) for **2a'** or a sealed-tube $\text{MoK}\alpha$ source ($\lambda=0.71073$ Å) for **2b** and **3b**. Structures were solved using SHELXT [15] and refined using SHELXL [16]. Hydrogen atoms were visible in difference maps, placed in idealized positions during refinement, and treated as riding. The crystal of **2a'** was an inversion twin, and the BASF parameter refined to 0.481(16). Because the two independent molecules in **2a'** are related by an approximate inversion center at 0.484, $\frac{1}{2}$, $\frac{3}{4}$, restraints were necessary to prevent C atoms from becoming non-positive definite. The crystal and refinement data are presented in Table 1.

Table 1. Crystal data and refinement parameters.

	2a'	2b	3b
Chemical formula	$\text{C}_{25}\text{H}_{24}\text{Fe}_2\text{O}_2$	$\text{C}_{17}\text{H}_{13}\text{F}_3\text{FeO}_3$	$\text{C}_{36}\text{H}_{26}\text{Fe}_2\text{O}_4$
M_r	468.14	378.12	634.27
Deposition number	CCDC 2355187	CCDC 2355188	CCDC 2355189
Crystal system, space group	Monoclinic, $P2_1$	Monoclinic, $P2_1/n$	Monoclinic, $P2_1/c$
Temperature (K)	90	90	90
a, b, c (Å)	5.7838 (5), 11.8770 (9), 28.071 (2)	10.6270 (2), 9.7968 (2), 14.9960 (3)	9.8279 (8), 13.3787 (10), 10.8256 (8)
β (°)	93.773 (6)	110.1183 (9)	111.064 (4)
V (Å ³)	1924.2 (3)	1465.98 (5)	1328.29 (18)
Z	4	4	2
Radiation type	$\text{Cu K}\alpha$	$\text{Mo K}\alpha$	$\text{Mo K}\alpha$
μ (mm ⁻¹)	12.24	1.08	1.14
Crystal size (mm)	$0.15 \times 0.04 \times 0.01$	$0.38 \times 0.30 \times 0.13$	$0.22 \times 0.15 \times 0.14$
Diffractometer	Bruker Kappa APEX-II DUO	Bruker Kappa APEX-II DUO	Bruker Kappa APEX-II DUO

Absorption correction	Multi-scan	Multi-scan	Multi-scan
T_{\min}, T_{\max}	0.564, 0.887	0.764, 0.873	0.811, 0.857
No. of measured, independent and observed [$I > 2\sigma(I)$]	12160, 5225, 3851	74201, 13090, 12014	65909, 9349, 7963
reflections			
R_{int}	0.074	0.019	0.031
$(\sin \theta / \lambda)_{\max} (\text{\AA}^{-1})$	0.568	1.086	0.945
$R[F^2 > 2\sigma(F^2)], wR(F^2), S$	0.068, 0.178, 1.04	0.021, 0.059, 1.05	0.031, 0.086, 1.05
No. of reflections	5225	13090	9349
No. of parameters	524	244	190
$\Delta Q_{\max}, \Delta Q_{\min} (\text{e \AA}^{-3})$	1.01, -0.55	0.88, -0.48	1.02, -0.38

2.2. Experimental Procedures

3-(Ferrocenoyl)propionic acid (1a**).** The compound was synthesized by following the procedure reported in the literature [17].

5-ferrocenyl-4-trifluoroacetyl-2(3H)-furanone (1b**).** To the stirred solution of **1a** (500 mg, 1.74 mmol) in dichloromethane, trifluoroacetic anhydride (0.48 mL, 3.49 mmol) was added. The solution turned purple immediately. The reaction mixture was stirred at room temperature for 30 min. Volatiles were removed in vacuo. The crude product was purified by silica column chromatography using dichloromethane as eluent, and the purple band on the column was collected. Yield (221 mg, 35%). Melting Point: 108–110 °C. IR (ATR, cm^{-1}): 1825.93, 1726.00, 1690.86, 1553.54. ^1H NMR (400 MHz, CDCl_3 , ppm): δ 3.69 (br, 2H), 4.23 (s, 5H) 4.71 (br, 2H), 5.39 (br, 2H). $^{19}\text{F}\{^1\text{H}\}$ NMR (Acetone- d_6 , ppm): δ -76.5. $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, acetone- d_6 , ppm): δ 33.2, 68.9, 70.0, 70.5, 71.7(cp), 73.2, 103.0, 116.9 (q, $^1\text{J} = 290.4$ Hz, CF_3), 171.0, 172.8 (q, $^2\text{J} = 34.5$ Hz, C-O). ESI-MS: m/z 364.1.

(4-ferrocenoyl)butyric acid (2a**) and 1,3-diferrocenoylpropane (**2a'**):** To the stirred suspension glutaric anhydride (6.13 g, 53.8 mmol) and anhydrous aluminum chloride (14.3 g, 107 mmol) in 1,2-dichloroethane (80 mL), a solution of ferrocene (12.0 g, 64.6 mmol) in 1,2-dichloroethane (80 mL) was added dropwise. The reaction mixture was stirred at room temperature for 1 hr and poured into ice-cold water. The organic phase was separated, and the aqueous phase was extracted with dichloromethane (3 x 30 mL). The product was extracted in 1.0 M NaOH (3 x 40 mL). The aqueous phase was acidified with conc. HCl until precipitation was complete. The precipitate was separated by filtration, and the residue was dried in air to give **2a** (2.34 g, 14.5%) as an orange solid. ^1H NMR (400 MHz, CDCl_3 , ppm) : δ 2.03 (quint, $^3\text{J} = 7.2$ Hz, 2H), 2.50 (t, $^3\text{J} = 6.8$ Hz, 2H), 2.81 (t, $^3\text{J} = 6.8$ Hz, 2H), 4.18 (s, 5H, Cp), 4.49 (br, 2H, Cp), 4.80 (br, 2H, Cp). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, acetone- d_6 , ppm): δ 19.4, 32.5, 38.1, 68.4, 69.3, 69.6, 79.5, 173.7, 202.3. The melting point and spectroscopic data of the complex were in good agreement with the reported compound [18]. The organic phase was collected, volatiles removed in vacuo, and the products separated by silica column chromatography using a mixture of hexane and ethyl acetate (3:1). The second band of the column, after evaporation of solvents, gave **2a'** (1.26 g, 8.3%) as an orange solid. Analytically pure product was obtained by slow solvent evaporation from the diethyl ether and hexane mixture. Melting Point: 131–132.5 °C (Lit.[19] 125 – 127 °C). IR (ATR, cm^{-1}): 1664 (C=O). ^1H NMR (400 MHz, CDCl_3): δ 2.11 (p, $^3\text{J} = 7.2$ Hz, 2H), 2.83 (t, $^3\text{J} = 7.6$ Hz, 4H), 4.19 (s, 10H, Cp), 4.49 (t, $^3\text{J} = 2.0$ Hz, 4H), 4.82 (t, $^3\text{J} = 2.0$ Hz, 4H). $^{13}\text{C}\{^1\text{H}\}$ (100 MHz, CDCl_3 , ppm): δ 19.5, 38.9, 69.5, 69.9, 72.3, 79.0, 204.3. ESI-MS: m/z 469.0 [M + H]⁺. The structure of **2a'** was analyzed using single-crystal X-ray analysis.

6-ferrocenyl-5-trifluoroacetyl-3,4-dihydropyran-2-one (2b**).** We synthesized this compound by following procedures similar to **1b** but with a longer reaction time (12 h). Yield 49%. Melting Point: 86.1–87.9 °C. IR (ATR, cm^{-1}): 1790, 1703 (C=O), 1652 (C=C). ^1H NMR (400 MHz, Acetone- d_6): δ 2.82 – 2.84 (m, 2H), 2.87 – 2.90 (m, 2H), 4.27 (s, 5H), 4.53 (t, $^3\text{J} = 1.6$ Hz, 2H), 4.65 (t, $^3\text{J} = 1.6$ Hz, 2H). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, Acetone- d_6 , ppm): δ 20.8, 27.7, 70.1, 70.9, 75.5, 108.1, 116.3 (q, $^1\text{J} = 291.3$, CF_3), 163.9, 165.9, 181.5 (q, $^2\text{J} = 34.5$ Hz, CO). ^{19}F NMR (Acetone- d_6 , ppm): -73.8. The X-ray quality crystals were obtained by cooling its solution in hexane.

2-(Ferrocenoyl)benzoic acid (**3a**). The compound was synthesized by following the procedure reported in the literature [20].

3,3'-diferrocenyl-3,3'-diphthalide (**3b**). To the stirred solution of **3a** (500 mg, 1.49 mmol) in dichloromethane (10 mL), trifluoroacetic anhydride (1.05 mL, 7.48 mmol) was added. The reaction mixture was stirred at room temperature for 30 min, and then water (20 mL) was added. The product was extracted with dichloromethane (2 x 20 mL). The organic layer was dried with anhydrous MgSO_4 and filtered, and the filtrate evaporated to dryness. The crude product was purified by silica gel column chromatography using a mixture of dichloromethane and ethyl acetate (1:1) to give **3b** (0.236 g, 48%) as a pale-yellow solid. The ^1H NMR analysis of the product showed the mixture of two isomers in a ca. 1:1 ratio. The mixture was suspended in diethyl ether and filtered. The ether-insoluble product was further purified by recrystallization from chloroform by diffusion of ethyl ether. The single-crystal X-ray analysis of the product revealed the formation of a *Meso* isomer. **Meso**: Melting Point: decomposes above 200 °C. IR (ATR, cm^{-1}): 1767 (C=O). ^1H NMR (400 MHz, CDCl_3 , ppm): δ 3.48 (s, 10H), 3.69 – 3.70 (m, 2H), 3.73 – 3.75 (m, 2H), 4.04 – 4.05 (m, 2H), 4.36 – 4.37 (m, 2H), 7.50 (t, 3J = 7.6 Hz, 2 H), 7.71 – 7.75 (m, 4H, Ar), 7.99 (d, 3J = 7.6 Hz, 2H). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3 , ppm): δ 67.4, 67.8, 68.0, 68.6, 68.9, 84.5, 89.8, 124.0, 125.2, 126.6, 129.7, 133.6, 150.6, 169.6. **Racemic**: Melting Point: 124 °C. IR (ATR, cm^{-1}): 1728 (C=O). ^1H NMR (400 MHz, CDCl_3): δ 3.47 (s, 5H, Cp), 3.85 (br, 1H), 3.99 (br, 1H), 4.19 (br, 1H), 4.42 (br, 1H), 7.40 (t, 3J = 7.2 Hz, Ar), 7.56 – 7.63 (m, 2H, Ar), 7.73 (d, 3J = 7.2 Hz, Ar). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3 , ppm): 67.2, 68.5, 68.6, 68.8, 69.4, 85.5, 89.9, 123.4, 125.5, 126.7, 129.7, 133.8, 150.0, 169.7. **Monoferrocenyl dimerized lactone**. This product was collected as the first fraction of the column. ^1H NMR (400 MHz, CDCl_3): δ 2.16 (s, 1H), 3.76 (s, 5H), 4.08 (br, 1H), 4.20 (br, 1H), 4.27 (br, 1H), 4.49 (br, 1H), 4.09 (t, 3J = 2.4 Hz, 1H), 6.37 (dd, 3J = 5.6 Hz, 3J = 2.4 Hz, 1H), 7.03 (dd, 3J = 5.6 Hz, 3J = 1.6 Hz, 1H), 7.35 – 7.37 (m, 1H), 7.50 – 7.59 (m, 4H), 7.69 (t, 3J = 7.2 Hz), 7.84 (d, 3J = 7.6 Hz, 1H), 7.98 (d, 3J = 7.6 Hz, 1H), 7.18 – 7.20 (m, 1H). The reaction was performed by adding two-mole equivalents of ferrocene as **3a**. Under similar reaction and workup conditions, the yield of the product was found to be 85%.

5-Ferrocenyldihydro-2(3*H*)-furanone (**4b**). To a stirred solution of **1a** (1.00 g, 3.49 mmol) in aqueous 1.0 M NaOH (10 mL), NaBH_4 (0.661 g, 17.5 mmol) was added. The reaction mixture was refluxed for 1 hr. A brisk gas formation was observed during the reflux, and the compound changed its color from red to orange. The mixture was cooled to room temperature and layered with ethyl acetate. Concentrated HCl was added dropwise with constant stirring until the compound moved from the aqueous to the organic layer. The organic phase was collected, dried with MgSO_4 , and evaporated to dryness to give a viscous mass. The products were further purified by column chromatography in silica using DCM as eluent, and an orange band was collected. Yield: 0.583 g, 58%. Melting point: 129 -130 °C. IR (ATR cm^{-1}): 1758 (C=O). ^1H NMR (400 MHz, CDCl_3): δ 2.19 – 2.32 (m, 1H), 2.52 – 2.66 (m, 3H), 4.18 (s, 5H), 4.20 – 4.22 (m, 4H), 5.31 – 5.34 (m, 1H). ^1H NMR (400 MHz, Acetone- d_6): δ 2.30 -2.38 (m, 1H), 2.49 - 2.66 (m, 3H), 4.18 (s, 5H), 4.20 (br, 2H), 4.25 (br, 1H), 4.32 (br, 1H), 5.32 – 5.35 (m, 1H). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3 , ppm): 29.3, 29.7, 67.9, 70.1, 78.8, 78.9, 87.8, 176.9.

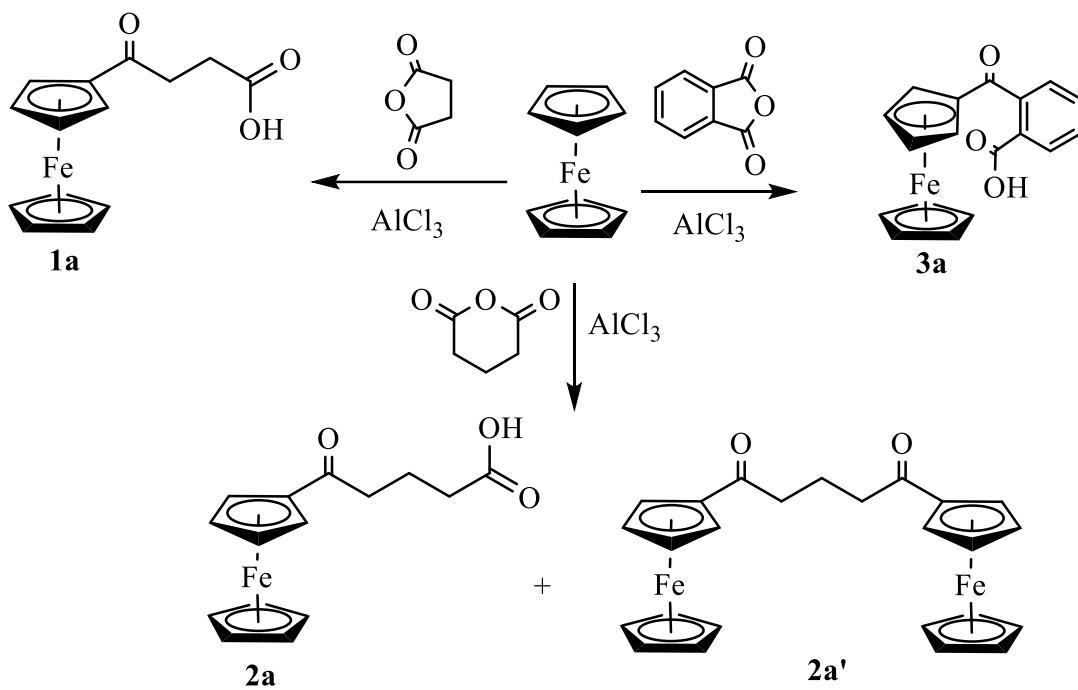
6-Ferrocenyl-2*H*-tetrahydropyran-2-one (**5b**). We synthesized this compound by following procedures similar to **4b** but with a longer reaction time (2 h). Yield: 28%. Melting point: 106 – 107 °C. IR (ATR, cm^{-1}): 1727 (C=O). ^1H NMR (400 MHz, CDCl_3): δ 1.81 – 2.01 (m, 3H), 2.18 – 2.23 (m, 1H), 2.46 – 2.54 (m, 1H), 2.61 – 2.68 (m, 1H), 4.17 – 4.18 (m, 7H), 4.20 – 4.23 (m, 1H), 4.24 – 4.26 (m, 1H), 5.16 (dd, 3J = 10.4 Hz, 3J = 3.6 Hz). $^{13}\text{C}\{^1\text{H}\}$ NMR (100 MHz, CDCl_3 , ppm): 18.7, 29.0, 29.6, 66.1, 67.3, 68.3, 68.4, 68.9 (Cp), 78.9, 87.1, 171.5.

3. Results and Discussion

Synthesis and Structural Elucidation of Compounds 2a, 2a', 1b – 5b and Proposed Mechanisms of Their Formation

As shown in Scheme 1, the Friedel-Crafts acylation of ferrocene with succinic anhydride, glutaric anhydride, and phthalic anhydride in the presence of AlCl_3 using dichloroethane as solvent gave 3-

(ferrocenoyl)propionic acid, **1a**, 4-(ferrocenoyl)butyric acid, **2a** and 2-(ferrocenoyl)benzoic acid, **3a**, respectively. Unlike **1a** and **3a**, the synthesis of **2a** yielded a neutral byproduct that displayed the presence of an unsubstituted cyclopentadienyl ring, two types of methylene protons (in a 1:2 integration ratio), and two sets of substituted cyclopentadienyl protons (in a 1:1 integration ratio) (Figure S1). FTIR (1664 cm^{-1}) and ^{13}C NMR (204 ppm) (Figure S2) indicated the presence of ketone function in the molecule. ESI-MS analysis of the product showed a strong signal at m/z 469 (Figure S3) corresponding to two ferrocenyl moieties, two ketones, and three methylene groups in the molecule. The single crystal X-ray structure of the product unambiguously indicated the formation of 1,3-diferrocenylpropane, **2a'**. Although the compound **2a'** was synthesized from the reaction of ferrocene and glutaryl chloride [21,22], to our knowledge, the formation of such a binuclear product by Friedel-Crafts acylation of ferrocene with glutaric anhydride has not been previously reported.



Scheme 1. Synthesis of ferrocenyl keto-carboxylic acids.

The compound **2a'** crystallizes in monoclinic space group $P2_1$ with two independent molecules in the asymmetric unit. These two molecules are related by an approximate inversion center (Figure 1). The Cp rings display nearly eclipsed conformation as in the closely related crystal structure of 1,2-diferrocenylethane [23]; the torsional angle C – Cg1 – Cg2 – C in four ferrocenyl moieties ranges from 1.21° to 21.42°. In the crystal, the compound displays weak intermolecular $\pi\cdots\pi$ interactions between two Cp rings, C-H \cdots π interactions between Cp-H and the centroid of Cp, and C-H \cdots O interactions between Cp-H and the carbonyl oxygen of two adjacent molecules (Figure S4).

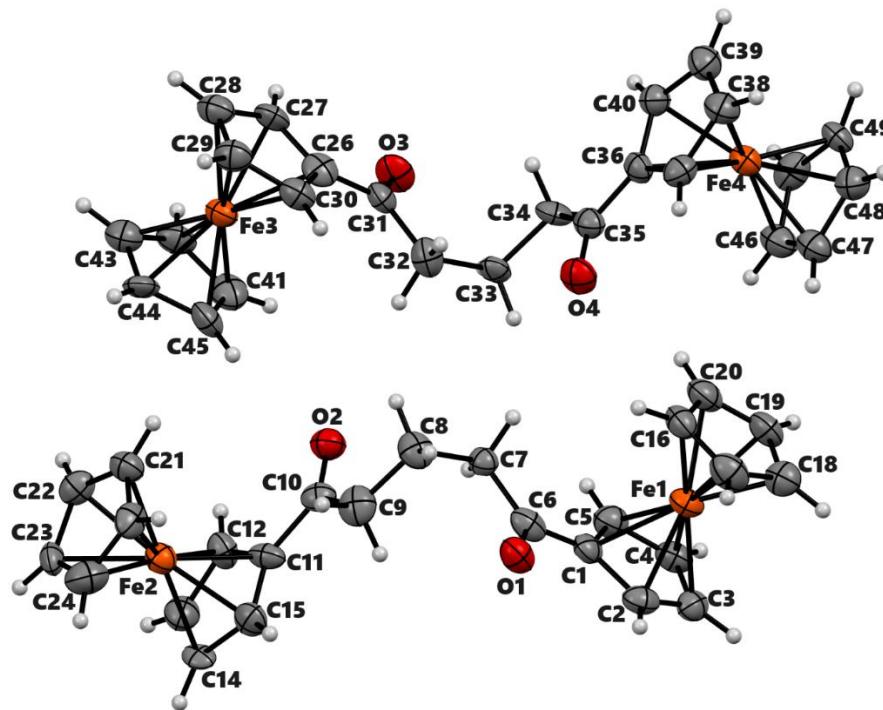
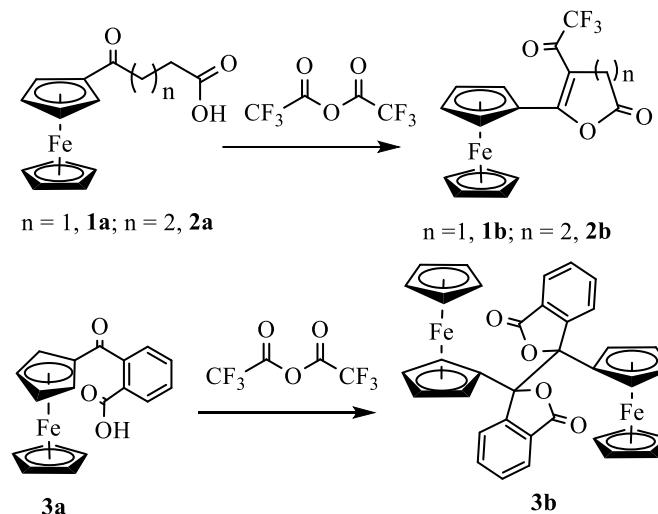


Figure 1. ORTEP diagram of solid-state structure showing the atom-numbering scheme of compound **2a'**. Displacement ellipsoids are drawn at the 50% probability level. Selected bond lengths (Å) for the complex: Fe1–C1 2.035(16), Fe1–C2 2.022(16), Fe1–C3 2.037(17), Fe1–C4 2.055(17), Fe1–C5 2.023(16), Fe1–C16 2.017(17), Fe1–C17 2.054(17), Fe1–C18 2.016(18), Fe1–C19 2.027(18), Fe1–C20 2.029(18), C1–C6 1.48(2), O1–C6 1.218(18), C6–C7 1.52(2), C7–C8 1.54(2).

Since the primary goal of this study was to investigate the nucleophilicity of the carbonyl group next to the ferrocenyl moiety, we treated 3-(ferrocenoyl)propionic acid, **1a** with trifluoroacetic anhydride (Scheme 2). On adding trifluoroacetic anhydride (TFAA), the compound immediately changed its color from red to deep purple. We separated the product by silica gel column chromatography using DCM as eluent. The ¹H NMR analysis of the product shows only four signals with an integration ratio of 2:2:2:5 (Figure S5), indicating the loss of two methylene protons compared to the starting keto-carboxylic acid, **1a**. The FTIR showed two strong signals at 1825 cm⁻¹ and 1726 cm⁻¹, indicating the presence of two carbonyl groups in the molecule. A high-energy IR band (1825 cm⁻¹) was consistent with a trifluoroacetyl group in the molecule, which was further confirmed by two quartets at 116.9 (¹J = 290.4 Hz) ppm and 172.8 (²J = 34.5 Hz) ppm in ¹³C NMR and a singlet at -76.5 ppm in ¹⁹F NMR (Figures S6, S7). ESI-MS analysis of the product showed a signal at m/z 365.0, but determining the structure of the compound was still a challenge. Finally, the compound yielded dark purple crystals when it was crystallized by slow vapor diffusion of hexane into its solution in dichloromethane under nitrogen. Although a complete structure refinement by X-ray crystallography was not possible due to the tiny size of the crystals and their twinning, the analysis was sufficient to identify the compound. The preliminary crystallographic results showed the formation of an α , β -unsaturated cyclic lactone with a trifluoroacetyl group attached to the β -position of the ring (Figure S8).



Scheme 2. Lactonization of ketocarboxylic acids in the presence of TFAA.

To understand the chemistry further, we treated 4-(ferrocenoyl)butyric acid, **2a**, with TFAA in DCM (Scheme 2). As expected, the reaction yielded the desired product **2b** in 49% isolated yield after chromatographic separation. The spectroscopic data (Figures S9 – S11) of the product were in agreement with the formation of the lactone in a similar way to the formation of **1b**. The single crystal X-ray analysis of the product unambiguously confirmed the formation of 6-ferrocenyl-5-trifluoroacetyl-3,4-dihydropyran-2-one (Figure 2). The compound crystallizes in the monoclinic space group $P2_1/n$ with a single molecule in its asymmetric unit. The bond length between C11 and C12 is 1.3599 (5) Å, confirming a double bond between α and β carbons. The six-membered lactone ring has a half-boat conformation with the Cremer & Pople [24] puckering parameters $Q_T = 0.4736(5)$ Å, $\theta = 117.34(6)^\circ$, and $\varphi = 30.74(6)^\circ$. The two Cp rings in the molecule acquire roughly eclipsed conformation with an average torsional angle C-Cg1-Cg2-C of 11.61°. The Cp rings of the ferrocenyl moiety are almost parallel, the dihedral angle between their planes being 2.73°.

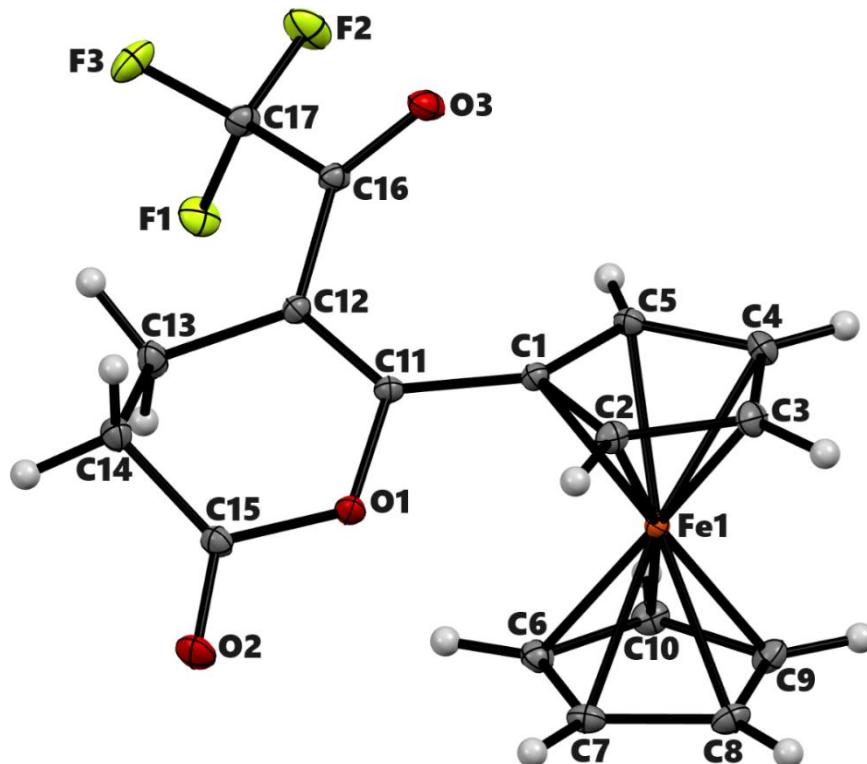
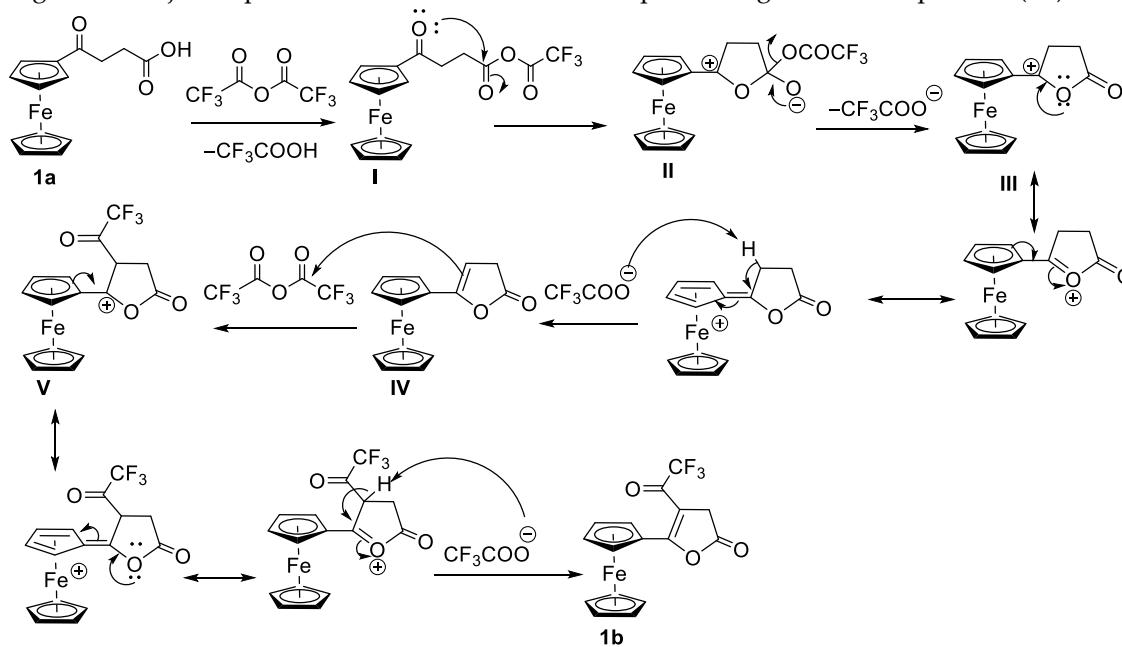


Figure 2. ORTEP diagram of solid-state structure showing the atom-numbering scheme of compound **2b**. Displacement ellipsoids are drawn at the 50% probability level. Selected bond lengths (Å) for the complex: Fe1–C1 2.0338(4), Fe1–C2 2.0310(4), Fe1–C3 2.0510(4), Fe1–C4 2.0607(4), Fe1–C5 2.0495(4), Fe1–C6 2.0515(4), Fe1–C7 2.0431(4), Fe1–C8 2.0528(4), Fe1–C9 2.0597(4), Fe1–C10 2.0621(4), C1–C11 1.4593(5), C11–C12 1.3599(5), O1–C11 1.3837(5), O1–C15 1.3823(5), O2–C15 1.1996(5), C13–C14 1.5262(6), C12–C16 1.4630(5), O3–C16 1.2149(5), C16–C17 1.5579(5), F1–C17 1.3377(5), F2–C17 1.3284(5), F3–C17 1.3458(5) Fe–Cp (centroid, substituted) 1.643, Fe–Cp (centroid, unsubstituted) 1.657.

The mechanism of intramolecular cyclization of the keto-carboxylic acid **1a** to form an α , β -unsaturated lactone **1b** can be proposed as shown in Scheme 3. The reaction between the carboxylic acid and trifluoroacetic anhydride gives a mixed anhydride (**I**) on which the trifluoroacetate group acts as an excellent leaving group. The carbonyl oxygen of the ketone makes a nucleophilic attack on the carboxy carbon to form a tetrahedral intermediate (**II**). Such an unusual nucleophilic attack of carbonyl oxygen might be possible due to the electron-donating effects of the ferrocenyl moiety at its alpha position. The tetrahedral intermediate loses trifluoroacetate, forming a resonance-stabilized carbocation (**III**). The stability of such carbonium ions is attributed to the electron-releasing effects of the ferrocenyl moiety as well as the nearby lone pair of electrons on the oxygen atom. The carbocation loses a proton from the β -position to create a double bond in conjugation with the cyclopentadienyl moiety (**IV**). The double bond attacks TFAA to give another resonance-stabilized carbocation with a trifluoroacetyl group in the ring (**V**). The inductive effects of trifluoroacetyl group and a positive charge at the adjacent position both facilitate the loss of proton to give the final product (**1b**).



Scheme 3. Proposed mechanism of the formation of the lactone.

The proposed mechanism involves the formation of 5-ferrocenyl-2(*3H*)-furanone, **IV** as an intermediate product from the β -elimination of carbocation **III**. We were interested in exploring the reaction outcomes in the absence of hydrogen at the β -position of ferrocenyl moiety. For this purpose, we treated 2-(ferrocenoyl)benzoic acid, **3a**, with trifluoroacetic anhydride in DCM. On adding TFAA, the compound changed color from dark red to amber at first and then to green. This color change is attributed to the formation of ferrocenium ions [25,26]. During aerobic workup, the reaction product changed to a lighter red color. The ^1H NMR analysis of the product indicated the formation of two isomers in a ca. 1:1 ratio. We separated the isomeric mixture by taking advantage of their different solubilities in diethyl ether. The substituted cyclopentadienyl rings of both isomers showed four multiplets, indicating the formation of a chiral center at the alpha position of the ferrocenyl moiety. A single signal in ^1H or ^{13}C NMR for the unsubstituted cyclopentadienyl ring for both isomers

indicated the presence of either a mirror plane or a rotational axis of symmetry in the molecule (Figures S12 – S15). The ether-insoluble isomer was crystallized from chloroform by vapor diffusion of diethyl ether. The single-crystal X-ray analysis of the product revealed the formation of *meso* isomer of dimeric lactone, 3,3'-diferrrocenyl-3,3'-diphthalide (Figure 3). The compound crystallizes in monoclinic space group $P2_1/c$ with the molecule lying on an inversion center. The lactone rings bend inward, making a dihedral angle of 48.42° with the cyclopentadienyl ring. The two Cp rings adopt a nearly eclipsed conformation with C-Cp1-Cp2-C dihedral angle of 2.18° . The C-C bond connecting the two monomers is $1.5751(12)$ Å. The ether-soluble product also contained the same number of signals in NMR as the *meso* isomer. Therefore, the stereochemistry of the product, by analogy, was assigned to be *racemic* (*R, S*). The *meso* isomer decomposes above 250°C while the *racemic* one melts at 124°C . The isomeric mixture of the complex **3b** was reported by Nesmeyanov et al. with limited characterization data [27].

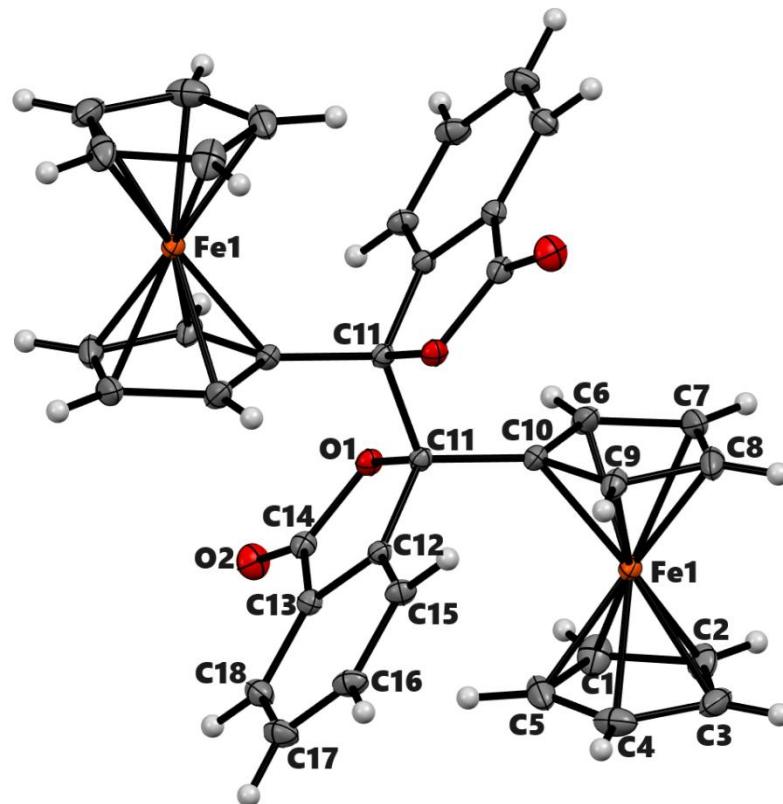
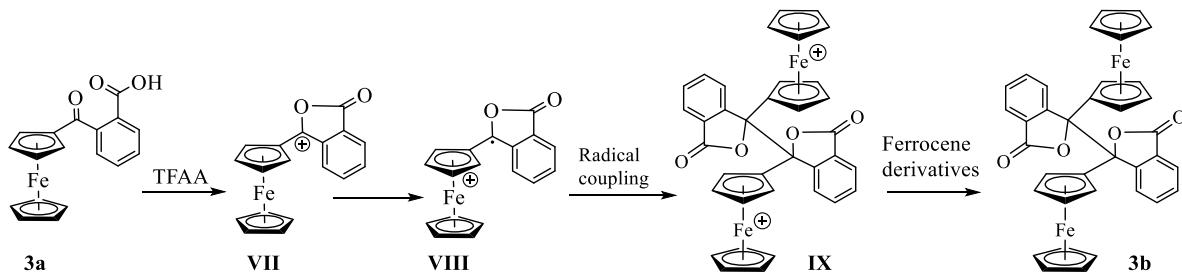


Figure 3. ORTEP diagram of solid-state structure showing the atom-numbering scheme of compound **3b**. Displacement ellipsoids are drawn at the 50% probability level. Selected bond lengths (Å) for the complex: Fe1–C1 2.0494(8), Fe1–C2 2.0576(8), Fe1–C3 2.0567(8), Fe1–C4 2.0500(9), Fe1–C5 2.0491(8), Fe1–C6 2.0502(8), Fe1–C7 2.0587(8), Fe1–C8 2.0542(8), Fe1–C9 2.0475(7), Fe1–C10 2.0405(7), O1–C14 1.3711(9), O1–C11 1.4578(8), O1–C14 1.3711(9), O2–C14 1.2084(9), Fe–Cp (centroid, substituted) 1.650, Fe–Cp (centroid, unsubstituted) 1.658.

The formation of the dimerized lactone from **3a** in the presence of TFAA further supports our hypothesis of intramolecular nucleophilic attack of carbonyl oxygen on the carboxylic carbon. As shown in Scheme 4, the nucleophilic attack of the carbonyl oxygen on the carboxy carbon, followed by a loss of the leaving group (*vide supra*), forms the α -ferrocenylcarbenium ion (**VII**). The carbenium ion undergoes internal oxidation-reduction (valence tautomerization), whereby the non-bonding electron from the d_{2g} set of iron is transferred to the positively charged carbon atom, giving a radical cation (**VIII**) [28]. The two radical cations undergo coupling to give a dication dimer (**IX**). The intermolecular oxidation-reduction between the dimer and ferrocene derivative in the solution produces the neutral dimer **3b**. Similar reaction mechanisms have been proposed earlier by Cais and

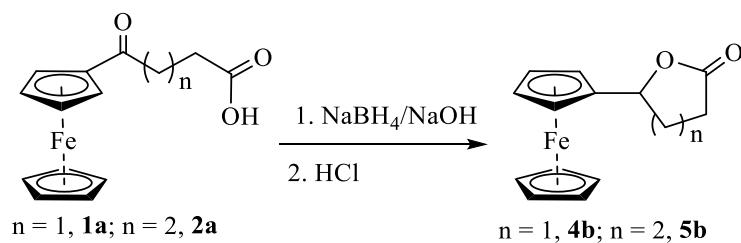
others [25,29–33] to explain the formation of dimerized products obtained from the solvolysis of α -ferrocenylcarbinols under acidic conditions.



Scheme 4. Proposed mechanism of the formation of dimerized lactone.

We attempted to isolate and characterize all the reaction products and their byproducts to validate the proposed mechanism. We could not isolate more than 48% of the dimeric products in our multiple synthetic efforts. We observed the formation of a black intractable residue, possibly due to the decomposition of cationic intermediates caused by the sacrificial oxidation of ferrocene derivatives. Careful separation of the reaction products by column chromatography indicated the product with one Fe(Cp) loss from dimer **3a**, as evidenced by ^1H NMR. The compound contains only one unsubstituted cyclopentadienyl group and 17 other chemically quinque protons, indicating the loss of CpFe moiety from the dimer (Figure S16). We added two equivalents of ferrocene in the reaction mixture to prevent the loss of the starting compound during reductive dimerization of the radical cation **VIII**. Under these conditions, the yield of **3b** increased to 85%, and no decomposition product was detected. We assume that ferrocene preferentially oxidizes to provide the electron to the carbenium ion.

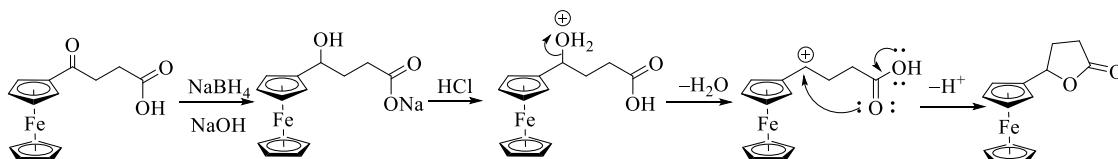
To get a better insight into the participation of the ferrocenyl moiety in activating the carbonyl oxygen for a nucleophilic attack, we decided to reduce the ketone to secondary alcohol and treat the hydroxycarboxylic acid with TFAA. The compound changed from red to yellow when compounds **1a** and **2a** were treated with NaBH_4 in $\text{NaOH}(\text{aq})$. The product remained soluble in NaOH , indicating the presence of the carboxylic acid group in the molecule. However, when the solutions were acidified using HCl , both compounds gave orange precipitates no longer soluble in aqueous NaOH (Scheme 5). The neutral product showed a strong signal at 1758 cm^{-1} (**4b**) and 1727 cm^{-1} (**5b**) in IR and a doublet of doublet corresponding to one proton at 5.31 ppm (**4b**) and 5.16 ppm (**5b**) in ^1H NMR indicating the formation of lactones (Figures S17 – S20). The lactone **5b** was crystallized from a diethyl ether and hexane mixture by slow evaporation. A single-crystal X-ray analysis of the product showed the formation of the desired product (Figure S21); however, its complete refinement was difficult due to weak scattering and crystal twinning. Formation of lactone **4b** under similar reaction conditions has been reported in the literature with limited characterization data [34,35].



Scheme 5. Reduction of ketone in keto-carboxylic acids and isolation of lactones after acidic workup.

These results indicate that isolating hydroxycarboxylic acids from reducing ferrocenyl keto-carboxylic acid is difficult after acid workup. The instability of such hydroxycarboxylic acid in the presence of acid might be due to the facile solvolysis of the 2° alcohol to give a stable α -ferrocenyl

carbenium ion. The carbenium ion is intramolecularly attacked by the oxygen atom of the carboxylic acid, giving a lactone as shown in Scheme 6.



Scheme 6. Mechanism of the formation of lactone from α -ferrocenylhydroxycarboxylic acids.

4. Conclusions

We have studied the uncommon lactonization in α -ferrocenyl ketocarboxylic acids in the presence of trifluoroacetic anhydride. We believe the important mechanistic step of lactone formation is the nucleophilic attack of ketone oxygen to carboxy carbon. The electron-donating effect of ferrocenyl moiety increases the electron density of carbonyl oxygen to facilitate the attack. These reactions yielded 5-ferrocenyl-4-trifluoroacetyl-2(3*H*)-furanone from 3-(ferrocenyl)propionic acid and 6-ferrocenyl-5-trifluoroacetyl-3,4-dihydropyran-2-one from 4-(ferrocenyl)butyric acid with structurally rare but biologically and pharmacologically important furanone [36] and pyranone [37] motifs. In these reactions, trifluoroacetic anhydride behaves as an electrophile, and its fragmentation product, trifluoroacetate, acts as a base. The 2-(ferrocenyl)benzoic acid, on the other hand, gave a dimerized lactone, 3,3'-diferrocenyl-3,3'-diphtalide caused by the valence tautomerization of α -ferrocenylcarbenium ion followed by dimerization of two ferrocenylphthalide units. The current work explains the necessity of reducing ketone to methylene before nucleophilic ring-closing in ferrocene-bound ketocarboxylic acids to synthesize homoannular ferrocene derivatives [17,38,39]. Our attempts to synthesize α -ferrocenyl hydroxycarboxylic acids by reducing the ketone function in 3-(ferrocenyl)propionic acid and 4-(ferrocenyl)butyric acid were unsuccessful. We believe that the incipient hydroxycarboxylic acid cyclizes by an intramolecular S_N1 -type mechanism. We are currently exploring the effects of acids on the redox potentials of the iron centers in acylferrocenes by electrochemical methods.

Supplementary Materials: The following supporting information can be downloaded at the website of this paper posted on Preprints.org. Figure S1: ^1H NMR of **2a'**; Figure S2: ^{13}C NMR of **2a'**; Figure S3: ESI-MS of **2a'**; Figure S4: Crystal packing of **2a'** along the crystallographic *a* axis; Figure S5: ^1H NMR of **1b**; Figure S6: ^{13}C NMR of **1b**; Figure S7: ^{19}F NMR of **1b**; Figure S8: Molecular structure of **1b**; Figure S9: ^1H NMR of **2b**; Figure S10: ^{13}C NMR of **2b**; Figure S11: ^{19}F NMR of **2b**; Figure S12: ^1H NMR of **3b** (*meso*); Figure S13: ^{13}C NMR of **3b** (*meso*); Figure S14: ^1H NMR of **3b** (*racemic*); Figure S15: ^{13}C NMR of **3b** (*racemic*); Figure S16: ^1H NMR of demetallated product from lactone **3b**; Figure S17: ^1H NMR of **4b**; Figure S18: ^{13}C NMR of **4b**; Figure S19: ^1H NMR of **5b**; Figure S20: ^{13}C NMR of **5b**; Figure S20: Molecular Structure of **5b**.

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