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The Influence of Solid Heat Carrier Load of Char on Pyrolysis Characteristics of Pulverized Coal in a Fluidized Bed Reactor

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Abstract: Pulverized coal pyrolysis based on solid heat carrier has a huge advantage in high tar yield. In this study, pyrolysis experiments of pulverized coal were conducted in a lab-scale fluidized bed reactor at 650 °C, utilizing char as the solid heat carrier. The influence of mass ratio of char to coal (RATIO) was investigated. Results show that the incorporation of solid heat carrier of char significantly enhanced the initial reaction in coal pyrolysis, resulting in increasing yields of tar and gas but reducing one of char. The yield of tar maximally reached 148.80–262.22% of the Gray-King analysis value at the RATIO of 14.52 g/g. As the RATIO increased, the tar contained more light component content, indicating that incorporating solid heat carriers improved the tar quality. These findings offer significant insights for the design of fluidized bed pyrolysis unit utilizing char as solid heat carrier.

Keywords: pyrolysis; fluidized bed; pulverized coal; solid heat carrier; char; pyrolysis tar

1. Introduction

Pyrolysis, a widely-used chemical conversion technology, enables the conversion of coal into liquid products and other chemical raw material under mild conditions [1,2]. This process is crucial for achieving efficient and clean utilization of coal [3,4]. At present, pyrolysis of lump coal (25–80 mm) has achieved industrial-scale applications. With the advancement of mechanized coal mining technology, however, the proportion of lump coal in total quantities has been decreasing [5]. Consequently, various pyrolysis technologies with pulverized coal as raw materials have been developed. With continuous understanding of pyrolysis mechanisms, it is found that the size of coal particles plays a pivotal role in the heat and mass transfer within particles during pyrolysis [5,6] and fine particles can effectively shorten the release pathway for volatile precursors, enhancing the release of volatile products [7,8]. In order to maximize the extraction of high value-added hydrocarbon resources from coal, there exists a trend in the development of coal pyrolysis technology towards finer particle size of raw material.

Due to direct and critical impacts on pyrolysis characteristics, heating process of pulverized coal is the core of pyrolysis technology and has always been of great concern. Among numerous heating methods, the scheme of solid heat carrier has great advantages in rapid heating [9]. For such method, common solid heat carriers include sand, ceramic balls, char, and furnace slag [10,11]. Solid heat carriers are first preheated to a predetermined temperature and then heat up coal particles at a rate of hundreds or thousands of degrees per second by mixing. Compared to current slow pyrolysis process of lump coal, rapid thermal shock caused by solid heat carrier results in significant changes in pyrolysis performances of pulverized coal [1]. For instance, the initial temperature of gas precipitation decreases, the yield of primary tar and gaseous hydrocarbons increases substantially, and it is easier to regulate the pyrolysis products by controlling the pyrolysis process [12]. Fushimi et al. [13] investigated pyrolysis performances of Loy Yang brown coal in a quartz glass downer pyrolyzer with silica sand as solid heat carrier. It was found that an increase in the ratio of sand/coal was beneficial for the generation of CO₂ and CO and heavy tar was rapidly deposited on the surface

of solid particles, forming cokes by polymerization. Meng et al.[14] studied the effects of ash-ZnO heat carrier (A-ZHC) on the release and transformation of sulfur during the pyrolysis of one high-sulfur bituminous coal. Results showed that under the catalytic action of A-ZHC, sulfur was more inclined to transfer from tar to char in forms of sulfide and sulfone sulfur.

Recently, numerous technologies based on solid heat carrier have been utilized in the pyrolysis process [15–17], including rotary kiln [18], auger reactor [19], moving bed [20,21], and fluidized bed [22]. Notably, due to the features of high mass and heat transfer and strong mixing, fluidized bed reactor has been extensively employed. Such technology typically consists of two fluidized bed reactors, one used to provide solid heat carrier and the other used to achieve the pyrolysis process through fluidized mixing between pulverized coal and solid heat carrier. There exists an energy and mass balance between the two reactors, adjusted through solid circulation and reaction control. On the basis of actual needs, one reactor operates in a gasification or combustion state, accordingly providing the heat carrier of char [23,24] or ash [20,25]. Compared to ash, the heat carrier of char has advantages of significantly improving the yield of gaseous products from coal pyrolysis due to its superior surface area and pore volume [26]. Besides, char exhibits catalytic an effect on volatile components, increasing the light oil fraction in the tar [27].

While the impact of char as heat carrier on coal pyrolysis characteristics has been extensively examined, these studies primarily focus on the variation patterns of pyrolysis products and kinetics of pyrolysis reactions. Research aiming at achieving high yield and superior quality tar remains limited and not sufficiently in-depth [28], and the findings seem not align with practical engineering applications. For a self-sustaining pyrolysis process, the determination of mass ratio of solid heat carrier to coal is crucial for entire process operation such as pyrolysis temperature and circulation ration of solids, as well as coal pyrolysis behaviors including free radical fragment reaction and yield and quality of pyrolysis products [29]. Based on this, in this study, the impacts of mass ratio of char to coal (RATIO) on the pyrolysis characteristics of pulverized coal were systematically explored in a lab-scale fluidized bed pyrolyzer. The aim is to determine the optimal RATIO for efficient pyrolysis of pulverized coal and provide theoretical guidance for entire process design.

2. Experimental

2.1. Samples

Shenmu bituminous coal (SM) was used in this study. Table 1 gives the results of coal properties. Gray-King analysis indicates that the yield of tar in SM is 5.8%. The particle size of the coal, as shown in Figure 1, ranges from 0 to 0.1 mm, with a particle size for 50% cumulative volume fraction (d_{50}) of 71.58 μm .

The solid heat carrier of char employed was prepared from SM by pyrolysis at 700 $^{\circ}\text{C}$ for 4 hours, named SMC herein. For such char, it is believed that no tar will be generated at pyrolysis temperature below 700 $^{\circ}\text{C}$. Before the experiments, the SMC was sieved to a particle size range of 0.5–1.0 mm and the particle size of d_{50} was 741.98 μm . Table 1 and Figure 1 show corresponding property results of SMC.

Table 1. Ultimate and Proximate analysis and Gray-King analysis of SM and SMC.

Sample	Ultimate analysis (%), ad)						Proximate analysis (%), ad)			Gray-King analysis (%), ad)	
	C	H	O ^a	N	S	M	V	FC	A	Tar Yield ($Y_{\text{G-K, tar}}$)	
SM	70.30	4.47	13.00	0.81	0.40	4.45	29.43	59.55	6.57	5.60	
SMC	70.64	0.88	4.15	0.60	0.31	7.35	7.48	76.24	8.93	Not detected	

^a Calculated by difference; M-moisture; V-volatile matter; FC-fixed carbon; A-ash; ad- air dry basis.

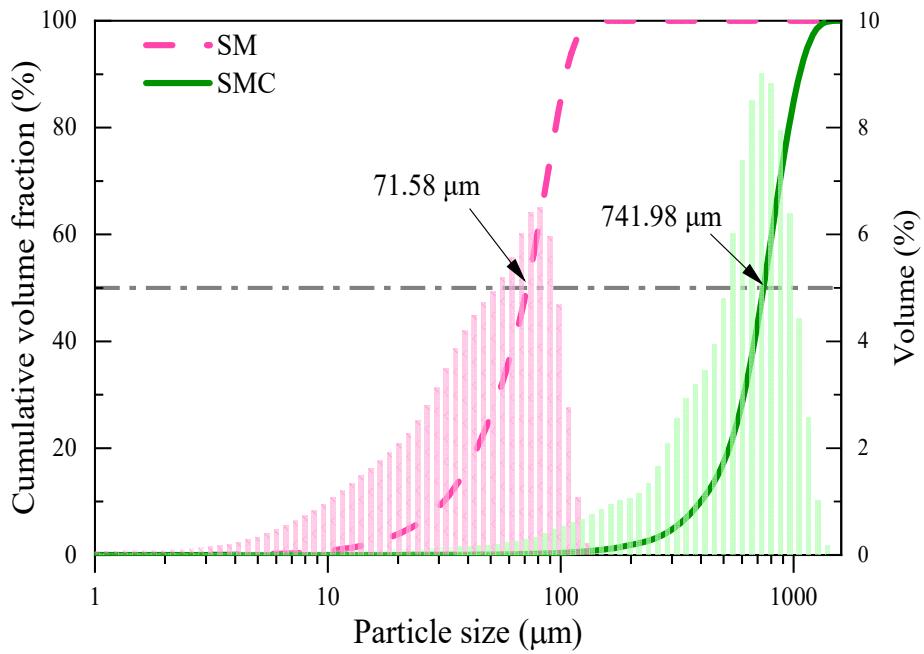


Figure 1. Particle size distribution of SM and SMC.

2.2. Experimental apparatus and procedure

The schematic diagram of pyrolysis apparatus is illustrated in Figure 2. The reactor, composed of stainless steel, has an inner diameter of 74 mm and an height of 530 mm. An integrated metal filter with a pore size of 2 um is installed at the upper part of the reactor, preventing the escape of fine solid particles from affecting subsequent gas and tar sampling. One K-type thermocouple is positioned above the air distribution plate to monitor the temperature of dense phase. Additionally, a pressure sensor is employed to oversee the pressure inside the reactor. The pulverized coal is loaded to the reactor bottom through a conveying pipe under the action of sowing coal wind of N₂. One 4 L/min N₂ flow is introduced into the reactor from the bottom, maintaining the fluidization state of solids inside the reactor. Gas flows are regulated by mass flow meters with an accuracy of 1%.

Before feeding the pulverized coal, the solid heat carrier of SMC was placed into the reactor and preheated by the heating furnace to 650 °C, the optimal temperature for maximum tar yield [27]. Such a design regarding the temperature of SMC would avoid the secondary cracking of tar due to higher temperature, as well as the release of volatiles from the solid heat carrier of SMC, interfering the results regarding the tar yield and pyrolysis gas composition. Once fed into the reactor, the pulverized coal would be rapidly heated by solid heat carrier, triggering a fast pyrolysis process. For collecting tar, the pyrolysis gas discharged from the reactor would successively undergo the processes of low-temperature cooling by refrigeration pump (-15 °C), two-stage extraction by isopropanol solution, and filtration by fiberglass, as illustrated by Figure 1. Accordingly, the tar produced was distributed in the cone-shaped collection bottle, isopropanol solution, and fiberglass filter. Following this, the gas free of tar was detected by an online gas analyzer for gas composition. When the concentration of combustible gas was lower than 5% and almost no longer changed, the pyrolysis reaction was considered finished. At this time, the experiment was terminated by turning off the heating furnace and gas flow. In this study, only the factor of the mass ratio of SMC to SM (RATIO) was investigated.

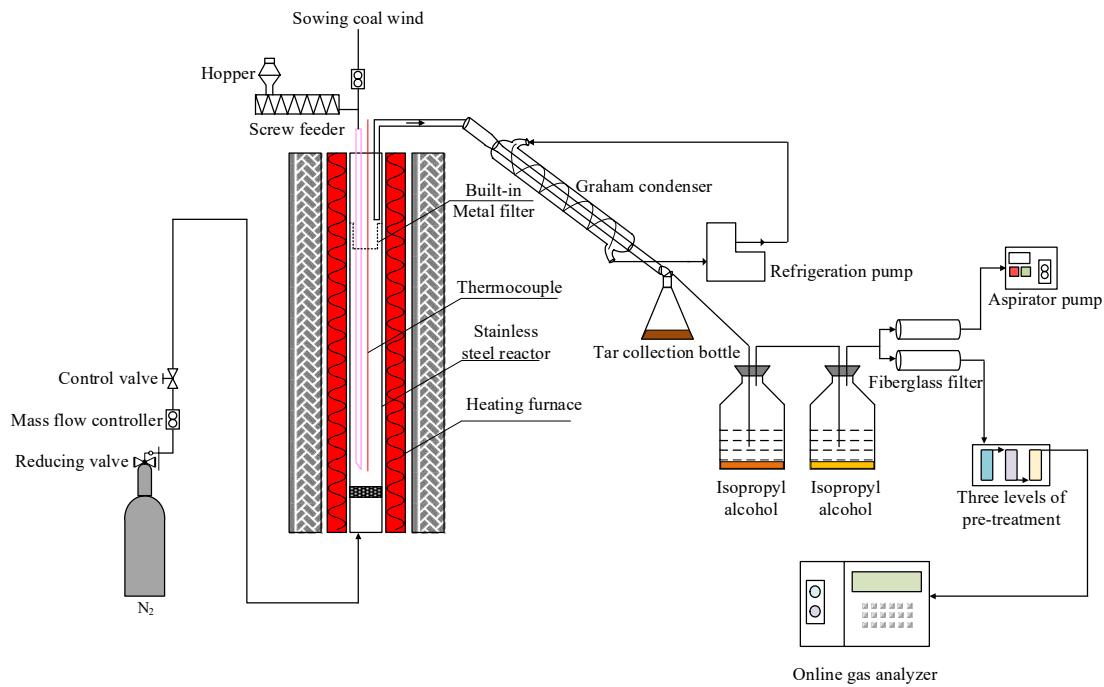


Figure 2. The schematic diagram of the SM pyrolysis experimental apparatus.

2.3. Sample characterization

A gas chromatography mass spectrometry (GC-MS, Agilent 6890A GC, Agilent 5375 MS) was employed to detect the primary composition and structure of organic compounds in tar samples. The instrument was set at the conditions of the solvent delay time of 3 minutes, the ratio of the injector to the detector of 50:1, the ion source temperature of 200 °C, and the mass-to-charge ratio m/z ranging from 50-550. During the detection, the tar sample was heated at a rate of 6 °C/min to 280 °C, at which was maintained for 10 minutes, using helium as carrier gas. The peak area integral method was adopted to quantify the high concentration compounds typically found in tar. Each peak in the chromatogram represents a specific substance, determined by comparing with standard spectra of compounds in the NIST database.

To obtain the fraction composition of tar, the gas chromatographic simulation distillation experiment was conducted using the 7890A GC manufactured by Agilent Company. The analysis method was according to the standard of ASTM D2887-01a and SH/T 0558-95.

The pyrolyzed char was characterized by N_2 physical adsorption test using a physical adsorption analyser (Quantachrome, Autosorb-IQ, USA). Prior to the test, all samples underwent a drying and degassing process at 100°C for 12 hours. According to the test data, the microscopic morphology of char was obtained. The specific surface area (S_{BET}) was determined through the multi-point BET method; the total pore volume (V_{TOT}) was derived from the equilibrium adsorption capacity at the ratio P/P_0 of ~1.00; and the pore size distribution was computed using the BJH method.

The surface functional groups of char were examined using a Fourier Transform Infrared Spectrometer (FT-IR). Before the examination, the char was subjected to drying at a temperature of 105 °C for 12 hours to eliminate external water. In the tablet pressing, the mass ratio of KBr to char was approximately 100:1. The spectra from 4000 to 400 cm^{-1} were recorded with a resolution of 4 cm^{-1} .

2.4. Data processing

According to the concentration of pyrolysis gas component measured by the online gas analyzer and assuming that no N_2 is generated in the pyrolysis, the volume yield of pyrolysis gas component i ($y_{g,i}$, L/g), the volume of pyrolysis gas component produced per unit mass the SM fed, can be calculated by N_2 balance, as follows.

$$y_{g,i} = \frac{Q_{N_2} \times \int_0^t \frac{v_{i,t}}{v_{N_2,t}} dt}{m_{coal}} \quad (1)$$

where i represents CO, CO₂, CH₄, H₂, and C_nH_m (referring to the sum of C₂H₄, C₂H₆, C₃H₆, and C₃H₈); Q_{N_2} is the flow rate of N₂ introduced, L/min; t is the time spent on the pyrolysis experiment, min; $v_{i,t}$ is the volume fraction of pyrolysis gas component i at the t minutes of pyrolysis experiment, %; $v_{N_2,t}$ is the volume fraction of N₂ at the t minutes of pyrolysis experiment, %; m_{coal} is the mass of SM used, g.

The yield of pyrolysis gas (Y_g , %), the mass percentage of pyrolysis gas produced to the SM fed, can be calculated via Equation (2).

$$Y_g = \frac{Q_{N_2} \times \int_0^t \frac{v_{g,t} \times MW_{g,t}}{v_{N_2,t}} dt}{22.4 \times m_{coal}} \times 100\% \quad (2)$$

where $MW_{g,t}$ is the average molar mass of dry pyrolysis gas, g/mol.

The yield of char (Y_{char} , %), the mass percentage of char produced to the SM fed, can be calculated by following equation.

$$Y_{char} = \frac{m_{char}}{m_{coal}} \times 100\% \quad (3)$$

where m_{char} is the mass of char produced, g. The m_{char} can be determined by subtracting the mass of solid heat carrier from the total one of solids remained in the reactor after the experiment.

The yield of tar (Y_{tar} , %), the mass percentage of tar collected to the SM fed, can be calculated as follows.

$$Y_{tar} = \frac{m_{tar}}{m_{coal}} \times 100\% \quad (4)$$

where m_{tar} is the mass of tar collected during the pyrolysis process, g.

3. Results and discussion

3.1. Pyrolysis process of SMC

As mentioned in Section 2.1, the solid heat carrier of SMC was prepared at 700 °C. In theory, no tar would be released at the pyrolysis temperature of 650 °C. However, unexpected results that a small amount of tar is produced from the solid heat carrier cannot be ruled out in the experiments. In order to accurately evaluate the interference of solid heat carrier on the yield of tar under experimental conditions, a blank experiment of SMC was conducted without SM. According to the results of blank experiment, the tar yield of SMC was found to be 0.27 %. Consequently, in the spirit of rigor, the tar produced from SMC should be considered despite of very low.

3.2. Products distribution

The distribution of pyrolysis products is depicted in Figure 3. As evidenced by the yield distribution diagram, without the solid heat carrier, the yields of char, tar, and gas amount to 58.35%, 5.10% (91.05% of the Gray-King analysis yield), and 4.19%, respectively. It is evident that the addition of solid heat carrier significantly elevates the tar yield, reaching a maximum value of 14.68% at the RATIO of 25.32 g/g (262.22% of the Gray-King analysis yield). According to the results of blank experiment, however, this value incorporates the tar produced from the solid heat carrier of SMC. Deducting this portion of tar from the total yield, the maximum tar yield reduces to 8.33% (148.80% of the Gray-King analysis ratio), which is achieved at the RATIO of 14.52 g/g. In theory, under an optimal RATIO, the maximum tar yield ranges between 148.80% and 262.22% of the Gray-King analysis yield. Furthermore, at the RATIO below 9.67 g/g, it is observed that a higher RATIO results in a decreasing char yield but an increasing gas yield. This may be due to the intensification of primary pyrolysis reaction under the rapid heating of solid heat carrier [30]. However, when the RATIO surpasses 9.67 g/g, the char yield slightly rebounds with the increasing RATIO. In this case,

besides of intensification of primary pyrolysis reaction, the catalytic effect of such solid heat carrier on the secondary reaction of volatile matter released should be considered. Such catalytic effect will enhance the disproportionation of tar released (especially the heavy tar), generating smaller molecular gas and larger molecular coke [31]. Therefore, it is inferred that the solid heat carrier can regulate the pyrolysis reaction through thermal and catalytic effects.

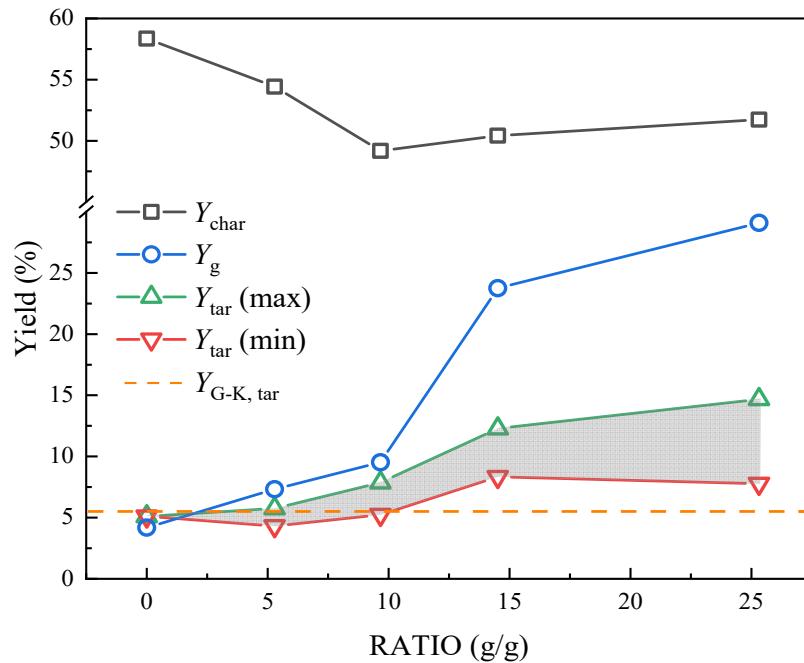


Figure 3. The effect of RATIO on product yields.

Figure 4 shows the variations of the volume yield of pyrolysis gas component with RATIO. It is evident that the pyrolysis gas mainly includes CH_4 , CO_2 , C_nH_m , H_2 , and CO . All yields of these components basically show an increasing trend with RATIO, contributing to the rise of pyrolysis gas yield. At the RATIO below 9.67 g/g, these yields follow an order of CO , H_2 , CH_4 , CO_2 , and C_nH_m (from high to low), and the yield of C_nH_m fluctuates without obviously increase. When above 9.67 g/g, the yield of H_2 exceeds that of CO , becoming the highest, and the yield of C_nH_m shows a clear upward trend. Continuing to increase the RATIO actually reduces the yield of CO_2 . These results indicate an important role of RATIO in regulating the pyrolysis gas components, especially in the range of 9.67–14.52 g/g. In this range, both thermal and catalytic effects of solid heat carrier promote an increase in pyrolysis gas component yield. Furthermore, C_nH_m is primarily produced by the demethylation or ethylation of methyl ether bridges, aromatic side chains, and chain hydrocarbons in tar. According to the increasing trend of C_nH_m yield with RATIO, it can be inferred that the catalytic effect of solid heat carrier gradually intensifies the secondary pyrolysis reaction when the RATIO is above 9.67 g/g.

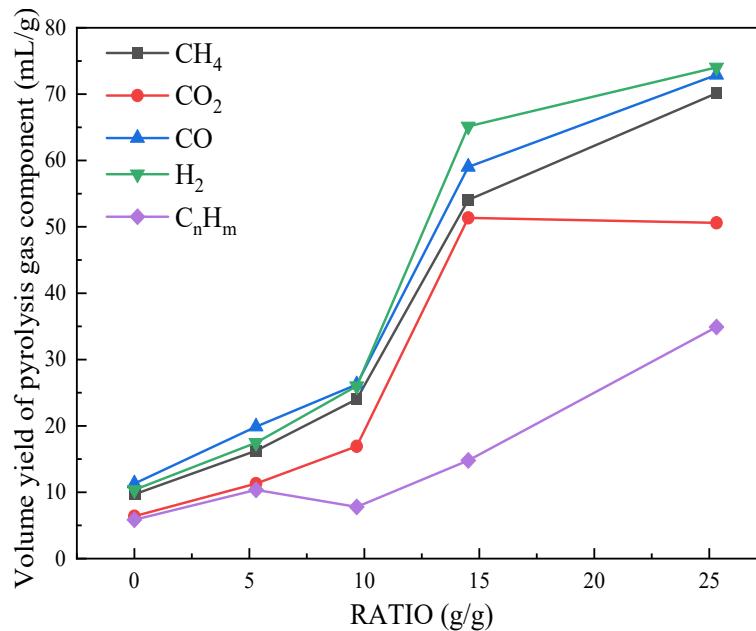


Figure 4. The effect of RATIO on pyrolysis gas components.

3.3. Composition of tar

In the GC-MS analysis of tar, dozens of hydrocarbons have been identified. For better analysis of tar chemical components, these hydrocarbons are sorted in an order of peak area percentage content from high to low, and those components with the same molecular formula but different molecular structure are merged. Following this method, 60 distinct hydrocarbon compounds are identified and analyzed, each of which possesses a unique structure. These hydrocarbons are categorized into aliphatic compounds, aliphatic ring compounds, monocyclic aromatic hydrocarbons, polycyclic aromatic hydrocarbons, and others. Figure 5 illustrates the percentage content of these five types of tar components at different RATIOS. Obviously, the tar mainly contains monocyclic aromatic hydrocarbons, polycyclic aromatic hydrocarbons, and aliphatic compounds, the sum content of which exceeds 96%. It is also found that there is a strong correlation between tar component (including the type and content) and RATIO. Specifically, in the absence of solid heat carriers, monocyclic aromatic hydrocarbons constitute the majority of the pyrolysis tar at 54.32%, while in the cases with solid heat carrier the content of such hydrocarbons decreases to varying degree and meanwhile the proportion of polycyclic aromatic hydrocarbons increases. This indicates that with the addition of solid heat carrier, the tar secondary reactions intensify, leading to the polymerization of monobenzene ring aromatic hydrocarbons to form polycyclic aromatic hydrocarbons. In this case, the aromaticity of the tar is enhanced. For instance, at the RATIO of 9.67 g/g, there exist 33 types of monocyclic aromatic hydrocarbons in tar, accounting for 43.1% of the tar.

To further evaluate the tar quality, a detailed ion flow chromatography of the tar is given in Figure 6, with the case of the RATIO of 14.52 g/g as an example. The primary compounds identified in this tar sample include aliphatic compounds (C₁₄-C₄₀ alkanes: 6.79%, alcohols: 1.44%), benzene (8.38%) and its derivatives (methyl phenol: 21.32%, phenol: 11.28%, hydrazine: 4.37%), and polycyclic aromatic hydrocarbons (naphthalene and its derivatives: 30.1%, hydrazine: 3.7%, hydrazine and its derivatives: 3.39%). Based on the tar compositions, it can be found that the tar predominantly consists of aromatic compounds and the vast majority are the aromatic hydrocarbons with two or more rings. Furthermore, the tar is also rich in low-grade phenolic compounds. Given these characteristics, this tar is apt for the extraction of phenolic chemical raw materials or for hydrogenating gasoline and other fuels, making it an excellent choice of raw materials for high-quality hydrotreating applications [32].

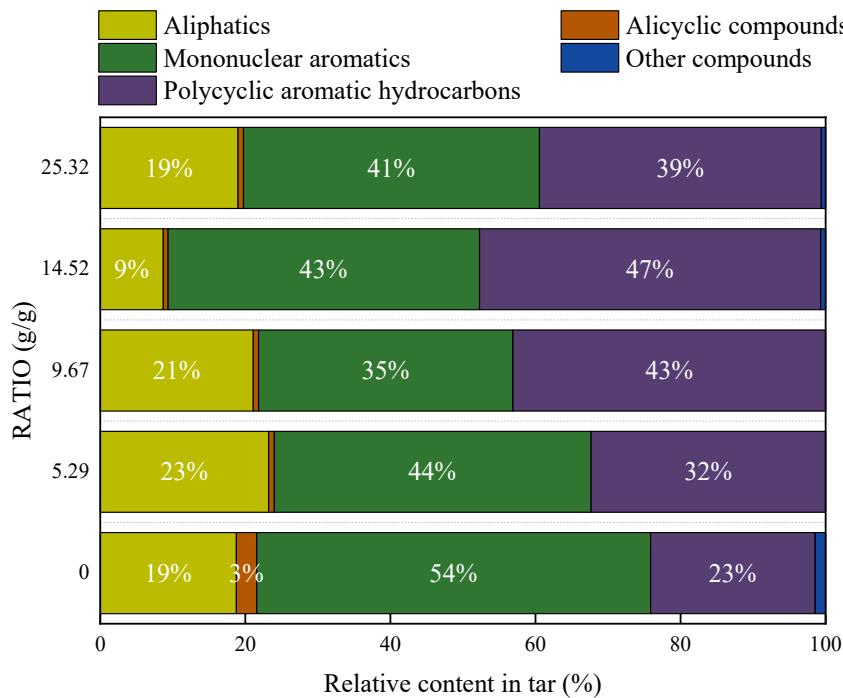


Figure 5. The effect of RATIO on chemical composition of pyrolysis tar.

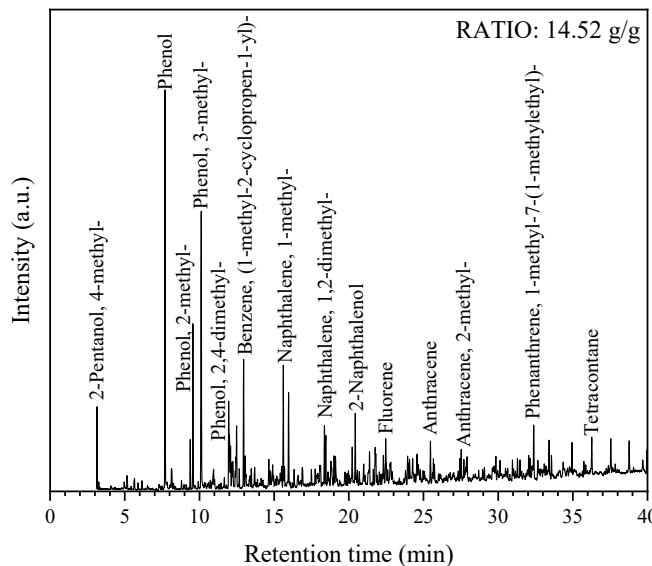


Figure 6. Gas chromatogram of pyrolysis tar at the RATIO of 14.52 g/g.

Figure 7 illustrates the fraction distribution of pyrolysis tar at different RATIOs. According to the boiling points of different fractions, the tar components can be divided into six categories: light oils (below 170 °C), phenol oils (170-210 °C), naphthalene oils (210-230 °C), washing oils (230-300 °C), anthracene oil (300-360 °C), and asphalt (above 360 °C) [33]. Generally, the fractions with a boiling point below 360 °C are collectively referred to as light components. As seen from Figure 7, without solid heat carrier, the light component content is 30.75% and the heavy asphalt content stands at 68.75%. As the RATIO escalates, there is an increase in the light component content and a decrease in the asphalt content. This suggests that the heavy asphalt is transformed into lighter tar. Such transformation is obviously related to the aforesaid catalytic effect of solid heat carrier of SMC on heavy tar. In this transformation process, the aromatic ring structure of macromolecular asphalt breaks down into light tar with smaller aromatic ring structure, gas, and coke. At the RATIO of 14.52

g/g, the light component content seems to reach the maximum, constituting 54.67% of the tar content. Further increasing the RATIO, the light component content slightly diminishes and the asphalt content accordingly rebounds. Overall, relative to the pyrolysis tar devoid of solid heat carrier, the light component content in tar augments with RATIO, underscoring the potential benefits of solid heat carrier for enhancing tar quality.

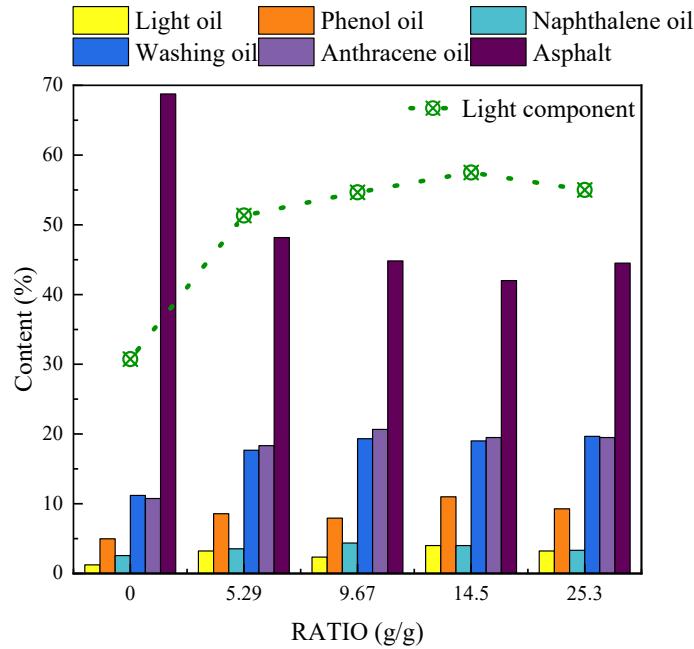


Figure 7. The effect of RATIO on fraction distribution of pyrolysis tar.

3.4. Characterization analysis of char

The results of porous structure and proximate analysis of pyrolysis char at different RATIOs are listed in Table 2. There seems to be a trend that with the increase of RATIO, the porous structure of char develops first and then declines. For instance, at a relatively low RATIO of 5.29 g/g, the char is characterized with a S_{BET} of 7.204 m^2/g and V_{TOT} of 0.02655 cm^3/g , while the S_{BET} and V_{TOT} of the char respectively reduce to 3.253 m^2/g of 0.02181 cm^3/g at a relatively high RATIO of 25.32 g/g. In a general, the more developed porous structure of the char, the more active sites on the char surface, which can promote the secondary cracking of tar [34]. However, the secondary cracking of large tar molecules will conversely form coke to block the pores inside the char [35]. Hence, the porous structure of char is the consequence of mutual constrain among different pyrolysis reactions. In these experiments, it is evident that negative effects of secondary cracking reaction on porous structure of char seems stronger at a higher RATIO. According to proximate analysis results, it is found that with the increase of RATIO, the volatile matter content of char shows a downward trend as a whole. At the RATIO of 9.67 g/g, the removal ratio of volatile matter reaches a maximum of 83.34 %. These results regarding proximate analysis also seems to be attributed to the more intensified secondary cracking reaction of tar at higher RATIO, leading to more volatile matter being converted into coke.

Table 2. Porous structure and proximate analysis of char at different RATIOs.

No.	RATIO (g/g)	S_{BET} (m^2/g)	V_{TOT} (cm^3/g)	Proximate analysis (%, ad)			
				M	V	FC	A
1	0	3.648	0.02606	0.00	11.29	77.46	11.26
2	5.29	7.204	0.02655	0.00	9.94	78.00	12.07
3	9.67	7.138	0.02675	0.00	9.97	76.67	13.36
4	14.52	3.207	0.02236	0.00	10.94	76.03	13.03
5	25.32	3.253	0.02181	0.00	9.82	77.48	12.70

Figure 8 shows the FTIR spectra of SM and chars at different RATIOS. Due to the obvious differences of FTIR spectra between SM and chars, especially in the range of 1750-400 cm⁻¹ and 3100-2700 cm⁻¹, it is confirmed that the surface chemical properties of SM have undergone significant change in the pyrolysis process. Similarly, such a conclusion also applies to the pyrolysis process of SM at different RATIOS because of visible differences of FTIR spectra in the range of 1750-400 cm⁻¹.

To quantitatively compare the changes in surface chemical property, in this study, seven infrared structural parameters (A_{ar} , A_{al} , $\frac{H_{al}}{H_{ar}}$, $A_{1750-1500}$, $I_{1750-1500}$, $A_{900-700}$, $I_{900-700}$) are defined according to Ref. [36,37], and the peak area of each characteristic peak is obtained by piecewise curve fitting to characterize the chemical structure of organic matter in char. A_{ar} is the absorbance integral area in the aromatic hydrogen band of infrared spectrum (3100-3000 cm⁻¹). A_{al} is the absorbance integral area in the aliphatic hydrogen band of the infrared spectrum (3000-2700 cm⁻¹). $\frac{H_{al}}{H_{ar}}$ is the absorbance integral area ratio of aliphatic hydrogen to aromatic hydrogen, representing the degree of thermal evolution of organic matter. A smaller $\frac{H_{al}}{H_{ar}}$ indicates a higher degree of thermal evolution of char. $A_{1750-1500}$ is the absorbance integral area in the band of aromatic ring C=C double bond and C=O double bond in the infrared spectrum (1750-1500 cm⁻¹). $I_{1750-1500}$ is the relative change of char macromolecule C=C double bond, characterizing the degree of fragmentation of char macromolecule. $A_{900-700}$ is the absorbance integral area in the band of benzene ring substitution in the infrared spectrum (900-700 cm⁻¹). $I_{900-700}$ is the relative change of benzene ring substitution in the char. The corresponding calculation results are shown in Table 3.

It can be inferred that in the pyrolysis process, the benzene ring structure in coal molecules and C=O side chain in macromolecules fracture, reducing the contents of C=C and C=O on the solid surface. Compared with the case without solid heat carrier, the intervention of solid heat carrier reduces the value of $I_{1750-1500}$ from 0.42 to within a range of 0.24-0.27. This indicates that the solid heat carrier intensifies the fracture of macromolecular benzene ring structure and the fragmentation seems unaffected by the RATIO. At the same time, $\frac{H_{al}}{H_{ar}}$ shows a decreasing trend from 1.45 to 0.32 with RATIO. Hence, under the action of more solid heat carrier, the char undergoes a stronger thermal evolution process. Similarly, the monotonously decreasing trend of $I_{900-700}$ from 1.52 at the RATIO of 0 g/g to 1.05 at the RATIO of 25.32 g/g is due to the fact that the aromatic groups of char macromolecules have been broken.

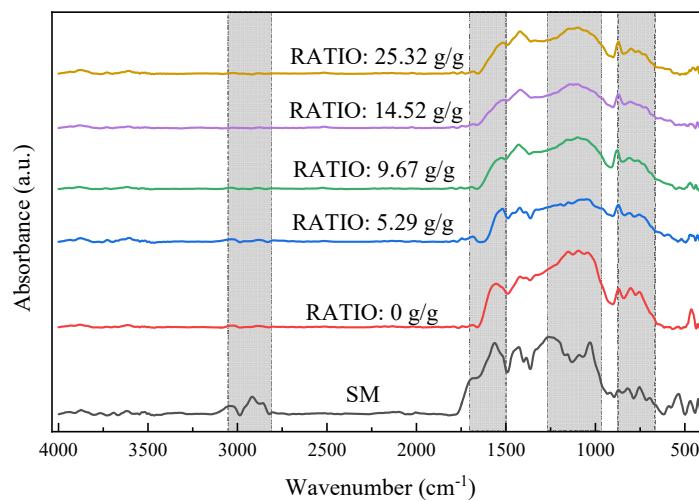


Figure 8. FTIR spectra of coal and char at different RATIOS.

Table 3. FTIR spectra parameter of raw coal and char.

Sample	RATIO (g/g)	A_{ar}	A_{al}	$\frac{H_{al}}{H_{ar}}$	$A_{1750-1500}$	$I_{1750-1500}$	$A_{900-700}$	$I_{900-700}$
SM	-	88.78	231.41	2.40	1635.93	1.00	600.26	1.00
Char	0	11.84	18.63	1.45	688.00	0.42	905.33	1.51

5.29	26.74	35.71	1.23	389.44	0.24	785.09	1.31
9.67	8.52	11.01	1.19	447.26	0.27	810.38	1.35
14.52	1.04	2.08	1.84	434.16	0.26	701.09	1.17
25.32	2.42	0.84	0.32	413.93	0.25	631.08	1.05

4. Conclusions

In this study, the pulverized SM with a particle size of 0-0.1 mm, was pyrolyzed at 650°C in a lab-scale fluidized bed reactor, utilizing SMC as the solid heat carrier. The influence of mass ratio of char to coal (RATIO) was investigated. Primary conclusions drawn are as follows.

(1) As the RATIO increased, there was a rise in the yields of gas and tar but a decrease in the yield of char. Notably, at the RATIO of 14.52 g/g, the tar yield reached the peak in a range of 8.33-14.68%, accounting for 148.80-262.22% of the Gray-King analysis value.

(2) A higher RATIO intensified the polymerization of monobenzene ring aromatic hydrocarbons, leading to a swift rise in both concentration and quantity of polycyclic aromatic hydrocarbons and thereby enhancing the aromaticity of the tar. Specifically, at the RATIO of 14.52 g/g, the tar mainly consisted of aromatic compounds, characterized with two or more aromatic rings and enriched in low-grade phenolic components, such as naphthalene and its derivatives (30.1%) and methyl phenol (21.32%). Besides, the presence of solid heat carrier of char improved the tar quality because of more light components in the tar.

(3) Solid heat carrier promotes the release of volatile matter by intensifying the primary pyrolysis reaction, while suppressing it by enhancing the secondary cracking of tar. The removal ratio of volatile matter reaches the maximum of 83.34 % at the RATIO of 9.67 g/g. The addition of solid heat carrier fractures the macromolecular benzene ring structure and the aromatic groups of char macromolecules, enhancing the thermal evolution process of char.

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