Full Factorial Design Synthesis of Silver Nanoparticles Using *Origanum vulgare*

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Abstract: Silver nanoparticles from *Origanum vulgare* (Oregano) have been synthesized in the past. However, no investigation has been performed on the combined effects of independent factors that affect the synthesis. Silver nitrate and Oregano concentrations, incubation temperature and time, as well as pH can play crucial roles in synthesis. In this regard Full Factorial Design was applied. A Voigt function was fitted on the measured UV – Vis spectra. The fitting parameters of the Voigt function (peak wavelength, area, and Full Width at Half Maximum) were used as the responses. A quadratic model was fitted for the peak wavelength and area, and the pH proved the dominant factor on nanoparticle synthesis.

Keywords: Silver nanoparticles; Green synthesis; full factorial design synthesis

1. Introduction

Silver Nanoparticles (AgNPs) have demonstrated excellent antimicrobial activities against a variety of microorganisms, which makes them suitable for materials used in medical products [1]. Nano silver also has been used in textiles, coated water filters, in treatment of mental illness and other areas [2].

Among various synthesis methods, green synthesis of AgNPs has received attention because it involves microorganisms and biomolecules, is eco-friendly and cheap [3,4]. Silver, gold and other metallic nanoparticles can be synthesized using plant extracts obtained from dry or fresh leaves, fruits, and seeds [4]. In a typical reaction, a metal salt solution is mixed with the extract, and a reduction of the metal ion occurs [4], with simultaneous capping of the formed nanoparticle [5]. A range of factors control the reaction such as metal salt and extract concentration, pH, incubation time and temperature [5,6].

Oregano (*Origanum vulgare L.*), is an aromatic herb rich in antioxidant compounds including phenolic acids, flavonoids and essential oils such as thymol, having as a result health benefits on human body [7]. Origanum mediated synthesized Ag NPs have been proposed as alternative antibacterial agents [8,9]. Silver nitrate ($AgNO_3$) aqueous solution with concentrations ranging from 0.5 mM up to 100 mM have been reported [1,9–12], used as the precursor for silver ions. The effects of reaction time and temperature were investigated [12].

In this work the combined effects of metal salt and extract concentrations, pH, incubation time and temperature were investigated using a full factorial design and response surface methodology. The UV – Vis absorption peak wavelength λ_0, the area under the absorption curve (A) and the Full Width at Half Maximum (FWHM) were used as the responses. To our knowledge this is the first time such a work is reported for AgNPs synthesized using Oregano.

2. Materials and Methods

2.1. Chemicals – Oregano Extract

AgNO₃ aqueous solution (0.1M) and sodium hydroxide (NaOH) pellets were purchased from Sigma Aldrich (Steinheim, Germany). Dried oregano (grown in the island of Lemnos) leaves were purchased from a local herb store in Lemnos Greece.

The Oregano Extract (OE) was prepared as follows: 3.76 gr of dried oregano leaves were boiled in 100 ml distilled water for 10 minutes. The final extract was obtained with filtering using cheesecloth. The extract was stored at 4 °C for future experiments.

2.2. Silver Nanoparticle Synthesis

A final volume of 2.5 mL aqueous solution was used in all synthesis unless otherwise stated. AgNO₃ at a concentration (C_{AgNO_3}) range 0.25 – 2 mM was used as the precursor of silver ions [13]. The reaction started by adding OE at a concentration (C_{ext}) in the range (0.8 – 20 % v/v). NaOH at a concentration (C_{NaOH}) range (0 – 7.9 mM) was also added. In a typical synthesis procedure, a mixture of AgNO₃, OE and NaOH at specific concentrations was heated in a water bath at a temperature (incubation Temperature) of 60 °C, for 1 h (incubation Time). All synthesis was performed in the dark without stirring.

2.3. Characterization techniques

UV – Vis spectroscopy measurements were performed using the Perkin Elmer – Lambda 25, UV – Vis spectrophotometer, in the spectral range 340 – 700 nm, using 1 mL cuvettes with 1 cm pathlength.

X- Ray Diffraction Patterns (XRD) and Fourier Transform Infrared (FTIR) experiments were carried out as described in a previous work [13]. The AgNPs investigated with these techniques (XRD and FTIR), were prepared with the following conditions: AgNO₃ concentration 1 mM, OE concentration 2 % v/v, and NaOH concentration 2 mM. The obtained nanoparticles were centrifuged at 20000 G, for 30 minutes at 4 °C, prior to these measurements.

2.4. Statistical Analysis and Experimental Design

As in a previous work [13], nanoparticle formation was observed by visual inspection for a colour change of the reaction solution and measuring the UV – Vis absorbance of an aliquot. The UV – Vis spectra were fitted using a Voigt profile [14] with fitting parameters [13]: the peak wavelength (λ_0) at maximum absorption, the Full Width at Half Maximum (FWHM), and the peak area under the UV – Vis curve (A). The best fit of the UV – Vis absorption spectra was found using the routine Peak Analyzer of the software Origin Pro (version 2020). The quality of fit was determined using the R square regression coefficient ($R^2 \geq 0.99$), and the fit significance was examined using ANalysis Of Variance (ANOVA) (p-value smaller than 0.05) [13].

The combined effects of the three independent factors: OE concentration (X₁), AgNO₃ concentration (X₂), and NaOH concentration (X₃), on the measured UV- Vis spectra using as responses the above fitting parameters were investigated using Full Factorial Design (FFD) under Response Surface Methodology (RSM). Three levels were used for each factor: (-1, 0 , and 1). Each of the factors was coded using the following equation [15]:

$$X_i = \frac{x_i - x_{oi}}{\Delta x_i}, i = 1,2,3 \qquad (1)$$

where X_i and x_i the coded and the actual value of independent factor i, x_oi the central value (level 0) of factor I, and Δx_i the step change of x_i corresponding to a unit change of the coded value [16].

The central values for the experimental design were: AgNO₃ concentration 1 mM, OE concentration 2 % v/v, and NaOH concentration 1 mM, with corresponding steps 0.5 mM, 1.2 % v/v, 1 mM respectively. The experimental design consisted of 27 runs (**Table 1)**, and each of the runs was repeated three times. These factors and their values were selected after preliminary experiments.

Table 1. FFD with three synthesis parameters (independent factors), with coded factor levels and actual values into parentheses and measured three mean responses (wavelength at peak maximum λ_0, peak area A, and FWHM) for AgNPs synthesis by OE.

Run	X_1 (C$_{ext}$ %v/v)	X_2 (C_{AgNO_3} mM)	X_3 (C$_{NaOH}$ mM)	Responses λ_0 (nm)	A (a.u.*)	FWHM (nm)
1	0(2)	0(1)	0(1)	412.9	129.4	93.7
2	-1(0.8)	1(1.5)	1(2)	405.8	190.5	94.2
3	-1(0.8)	0(1)	-1(0)	439.4	34.7	130.7
4	-1(0.8)	1(1.5)	-1(0)	442.7	31	97.4
5	-1(0.8)	-1(0.5)	-1(0)	438.1	33.5	140.2
6	0(2)	-1(0.5)	-1(0)	441.5	38.5	101.8
7	0(2)	-1(0.5)	0(1)	411.3	63.5	86.4
8	1(3.2)	1(1.5)	1(2)	412.2	214.2	98.6
9	1(3.2)	1(1.5)	0(1)	415.2	170.4	102.4
10	0(2)	0(1)	1(2)	407.5	141.9	93.5
11	-1(0.8)	-1(0.5)	0(1)	407.1	71.9	85.9
12	1(3.2)	0(1)	-1(0)	449.4	97.5	128.9
13	0(2)	-1(0.5)	1(2)	405.2	96.1	112.8
14	0(2)	1(1.5)	0(1)	414.4	147.9	98.5
15	1(3.2)	0(1)	1(2)	409.2	154.1	103.4
16	1(3.2)	1(1.5)	-1(0)	451.1	76.4	106.3
17	0(2)	0(1)	-1(0)	446.1	26.9	84.3
18	1(3.2)	-1(0.5)	1(2)	408.2	92.8	115.9
19	1(3.2)	-1(0.5)	0(1)	411.1	71.6	97.8
20	-1(0.8)	-1(0.5)	1(2)	408.2	97.8	133.7
21	-1(0.8)	1(1.5)	0(1)	412.1	119	95.3
22	1(3.2)	0(1)	0(1)	412.3	145.4	98.9
23	-1(0.8)	0(1)	0(1)	409	106.5	88.4
24	-1(0.8)	0(1)	1(2)	407.6	152.2	102.8
25	0(2)	1(1.5)	1(2)	409.1	194.9	94.2
26	1(3.2)	-1(0.5)	-1(0)	441.1	57.4	97.9
27	0(2)	1(1.5)	-1(0)	450.1	42.1	98.8

A second order polynomial was obtained for each response as function of the coded values (X_i) of the independent factors [13,17]:

$$Y_{(response)} = \beta_0 + \sum_i \beta_i X_i + \sum_{i,i} \beta_{i,i} X_i^2 + \sum_{i,j} \beta_{i,j} X_i X_j \qquad (2)$$

where β's are the regression coefficients [13,18].

Standard procedures were applied to assess the quality of the regression polynomials [19]. ANOVA was applied to assess the significance and adequacy of the model., as well as the significance of the regression coefficients appearing in the derived polynomial [20].

The magnitude and sign of the regression coefficients, and Pareto analysis [21], were used as a measure of the importance of the various independent factors and their interactions at significance level 5% (p-value < 0.05), unless otherwise stated. Not statistically significant terms were excluded from the polynomial models except of those required for a hierarchical mode [13].

3. Results and Discussion

3.1. UV – Vis spectra analysis

An absorption peak within the range of 400 – 450 nm, is indicative of AgNPs formation, which is attributed to a surface plasmon resonance [22–24]. Colour change of the reaction mixture also allows visual monitoring of AgNP formation [25].

In **Figure 1,** the combined effects of OE and NaOH concentrations on the UV – Vis spectra are plotted. A better absorption curve can be observed at the lowest extract concentration (OE 2 % v/v **Figure 1a**), with the presence of NaOH. A Voigt profile has been shown to fit adequate with UV – Vis

spectra of AgNPs solution in a previous study, providing a method of correlation between nanoparticle formation and synthesis parameters [13]. A single Voigt profile was suitable and in this work for the lowest extract concentration. Agglomerated or ill – formed nanoparticles [26], their crystallinity [27,28], and charge transfer between nanoparticles can influence the observed UV- Vis spectrum [29].
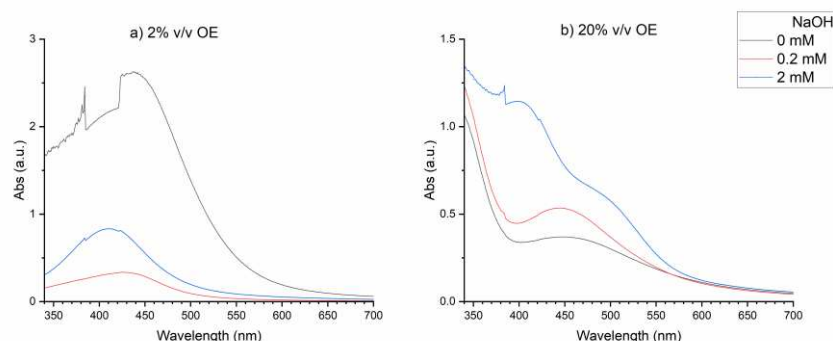
Figure 1. Effects of NaOH and OE concentrations on UV – Vis spectra. OE concentration a) 2% v/v, and b) 20% v/v.   NaOH concentration 0, 0.2 and 2 mM. AgNO$_3$ concentration 1 mM, incubation temperature 40 °C, and incubation time 10 min.

The addition of NaOH changes the pH of the solution. The pH of the reaction changes the chemical nature of the extract, which has an effect to change its performance and rate of reduction and therefore nanoparticle synthesis [30]. The size and shape of silver nanoparticles are affected from the pH of the reaction solution [31–34]. This was investigated further with the addition of hydrochloric acid (HCl), which did not produce any nanoparticles as shown in **Figure 2**.
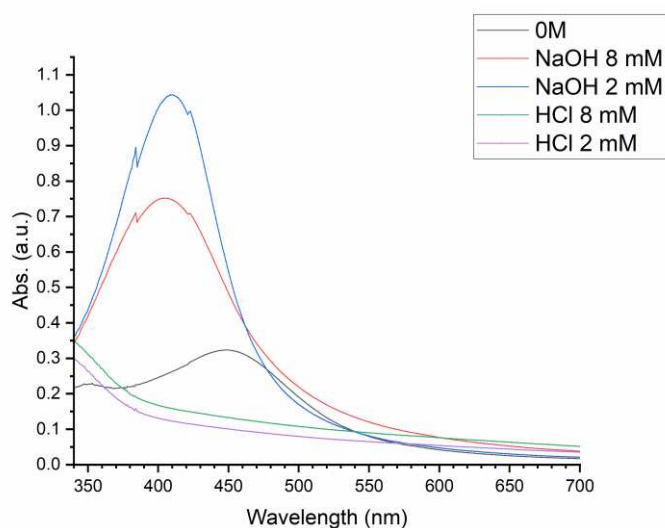
Figure 2. Effect of pH on UV – Vis spectra, with the addition of NaOH or HCl at different concentrations. AgNO$_3$ concentration 1 mM, OE concentration 2 % v/v, incubation temperature 60 °C and incubation time 1h.

The fitting parameters used in the Voigt profile are related to nanoparticle size, shape, yield and size distribution [13]. The aim was to determine the appropriate conditions that produce smallest peak wavelength with maximum peak area and narrowest FWHM. The investigation showed (see Supporting Information: **Figures S1–S3**, and **Tables S1 and S2),** that this is possible at incubation temperature of 60 °C, incubation time 1h, and NaOH concentration in the range (0 – 2 mM).

The corresponding ranges of OE and AgNO₃ concentrations were investigated in a similar manner as shown in **Figure 3.** The investigation (see Supporting Information: **Table S3**) determined a range (0.5 – 1.5 mM) for AgNO₃ and (0.8 – 3.2 % v/v) for OE concentrations.
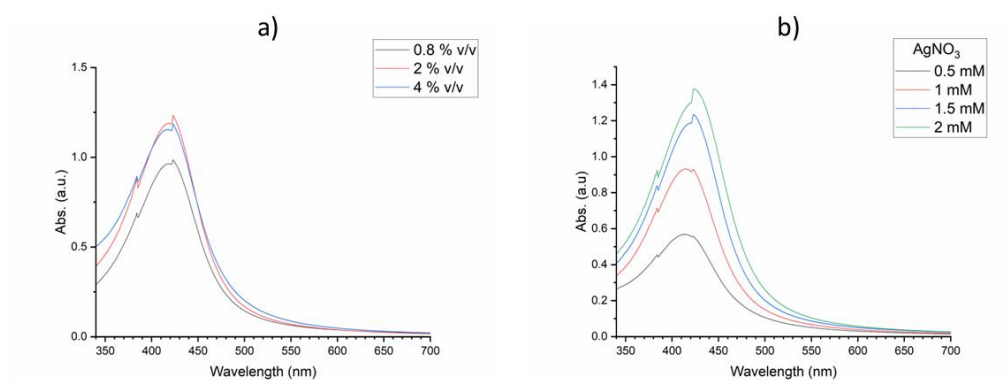
Figure 3. Effect of a) OE extract concentration, and b) AgNO₃ concentration on UV – Vis spectra. NaOH concentration 1 mM, incubation temperature 60 °C, incubation time 1h. a) AgNO₃ concentration 1 mM, and b) OE concentration 2 % v/v.

3.2. Statistical modelling of AgNPs synthesis

In order to elucidate on, the effect of all independent factors as well as their interactions, on the AgNPs synthesis, a quadratic regression model was applied (Equation 2), for each response. The ANOVA for each response is given in **Tables 2–4**.

Table 2. ANOVA for quadratic model for wavelength at peak maximum λ_0.

Source	df	SS	MS	F	p-value
Model	9	22322.92	2480.32	138.19	<0.0001
Linear					
X_1	1	264.48	264.48	14.74	
X_2	1	278.24	278.24	15.50	
X_3	1	17798.83	17798.83	991.68	
Square					
X_3^2	1	3813.3	3813.3	212.46	
Interaction					
X_2X_3	1	77.74	77.74	4.33	
Error	71	1274.32	17.95		
Lack of fit	17	117.17	6.89	0.3217	
Pure error	54	1157.15	21.43		
Total	80	23597.24			

df = degrees of freedom, SS = Sum of Squares, MS = Mean Square, F statistic.

Table 3. ANOVA for quadratic model for peak area A.

Source	df	SS	MS	F	p-value
Model	9	231000	25662.02	76.05	<0.0001
Linear					

	df	SS	MS	F
X_1	1	9827.11	9827.11	29.12
X_2	1	52886.07	52886.07	156.73
X_3	1	133900	133900	396.83
Square				
X_2^2	1	1556.99	1556.99	4.61
X_3^2	1	4321.69	4321.69	12.81
Interaction				
X_1X_2	1	2595.25	2595.25	7.69
X_1X_3	1	3111.7	3111.7	9.22
X_2X_3	1	21421.51	21421.51	63.48
Error	71	23958.08	337.44	
Lack of fit	17	6523.82	383.75	1.19
Pure error	54	17434.25	322.86	
Total	80	254900		

df =degrees of freedom, SS = Sum of Squares, MS = Mean Square, F statistic.

Table 4. ANOVA for quadratic model for FWHM.

Source	df	SS	MS	F-value	p-value
Model	9	8440.95	937.88	1.56	0.1435
Linear					
X_3	1	230.88	230.88	0.3847	
Square					
X_3^2	1	3220.51	3220.51	5.37	
Error	71	42616.33	600.23		
Lack of fit	17	8700.41	511.79	0.8149	
Pure error	54	33915.92	628.07		
Total	80	51057.28			

df =degrees of freedom, SS = Sum of Squares, MS = Mean Square, F statistic.

A p-value smaller than 0.001, can be observed for the quadratic model for both the peak wavelength at maximum λ_0 (**Table 2**), and the peak area A (**Table** 3) indicating that the quadratic model was suitable for both responses. This is further supported by the large value of the F statistic. In contrast, a pure model can be observed for the FWHM (**Table 4**) (p-value = 0.1435) and F = 1.56. The coefficient of determination (R^2) of the models for the peak wavelength was 93.92%, whereas for the peak area was 89.41%, which indicates a good correlation between observed and predicted values [15,17,35]**.** The model adequacy was tested by applying standard diagnostic tools (See Supplementary Material) [19,20]

The regression coefficients for the wavelength at peak maximum λ_0 are given in **Table 5**. The wavelength is affected mainly from the NaOH concentration (X_3), followed by equal contribution from OE (X_1), and $AgNO_3$ (X_2) concentrations. A large contribution appears from the quadratic term of NaOH concentration. A small contribution also appears from the interaction term X_2X_3 of $AgNO_3$ and NaOH concentrations. Terms were NaOH concentration appears, contribute negatively to the peak wavelength.

Table 5. Regression coefficients for wavelength at peak maximum λ_0.

Model term	CE	Std. Error
Intercept	412.4	1.25
X_1	2.21	0.5765
X_2	2.27	0.5765
X_3	-18.16	0.5765
X_2X_3	-1.47	0.7061
X_3^2	14.56	0.9986

CE = Coefficient Estimate, Std. Error = Standard Error, R^2 = 93.92%.

Table 6. Regression coefficients for peak area A.

Model term	CE	Std. Error
Intercept	114.44	5.40
X_1	13.49	2.5
X_2	31.29	2.5
X_3	49.80	2.5
X_1X_2	8.49	2.39
X_1X_3	-9.30	-15.40
X_2X_3	24.39	18.29
X_2^2	-9.30	-17.93
X_3^2	-15.49	-24.13

CE = Coefficient Estimate, Std. Error = Standard Error, R^2 = 89.41%.

The response surface (peak wavelength) for AgNPs synthesis therefore is provided by the following second order polynomial:

$$Y_{(peak\ wavelength)} = 412.4 + 2.21X_1 + 2.27X_2 - 18.16X_3 - 1.47X_2X_3 + 14.56X_3^2 \quad (3)$$

The regression coefficients for the peak area A are given in **Table 6**. The area is affected almost equally from NaOH (X_3) and AgNO$_3$ (X_2) concentrations, followed by the OE (X_1) concentration. A large contribution appears from the interaction term X_2X_3 of AgNO$_3$ and NaOH concentrations, followed by a smaller contribution from the interaction terms X_1X_2 of OE and AgNO$_3$ concentrations, and the interaction X_1X_3 with negative sign of OE and NaOH concentrations. A large negative contribution appears for the quadratic terms of both AgNO$_3$ and NaOH concentrations.

The response surface (peak area) for AgNPs synthesis therefore is provided by the following second order polynomial:

$$Y_{(peak\ area)} = 114.44 + 13.49X_1 + 31.29X_2 + 49.8X_3 +$$
$$+8.49X_1X_2 - 9.3X_1X_3 + 24.39X_2X_3 - 9.3X_2^2 - 15.49X_3^2 \quad (4)$$

Pareto analysis percentages are shown in **Table 7**, for the wavelength at peak maximum λ_0, and the peak area A quadratic model terms. In both cases the largest contribution comes from the NaOH concentration (X_3) with values 59.53% (wavelength) and 52.53% (area). A total of 97.8% for the peak wavelength is influence by the linear and quadratic term of NaOH concentration, in contrast to 57.61% for the peak area. The peak area is correlated to the quantity of formed nanoparticles which should depend on both AgNO$_3$ (20.74%) and OE concentrations (3.85%). This is not sufficient however, for the synthesis of nanoparticles as the presence of NaOH plays a crucial role, in both the number of formed nanoparticles (area) and their size (wavelength). This agrees, with previous work reporting the effect of pH on nanoparticle size [30,32,34].

Table 7. Pareto analysis [21] for significant terms in quadratic model for both the peak wavelength λ_0 and peak area A.

Peak wavelength λ_0		Peak area A	
Term	Per. Ef (%)	Term	Per. Ef (%)
X_1	0.88	X_1	3.85
X_2	0.93	X_2	20.74
X_3	59.53	X_3	52.53
X_2X_3	0.39	X_1X_2	1.53
X_3^2	38.27	X_1X_3	1.83
-	-	X_2X_3	12.6
-	-	X_2^2	1.83
-	-	X_3^2	5.08

Per. Ef = Percentage effect.

Examination of the data (**Table 1**) revealed that the best UV- Vis spectra in terms of minimum wavelength λ_0 at peak maximum, minimum FWHM and maximum area A, can be obtained for the factor combination corresponding to experimental run 23. This run corresponds to an actual OE

concentration of 0.8 %v/v, AgNO₃ concentration 1 mM, and NaOH concentration 1 mM. The measured peak wavelength 409 nm is smaller than the reported values by other groups [1,9–12], who have not investigated pH effects.

3.3. XRD

XRD pattern of dry powder obtained from the formed AgNPs was achieved. The analysis exhibited poor crystallinity of the sample with a broad peak formed ~25° possibly caused by the Oregano, although Ag contributions superimposed on the board pattern are visible (**Figure 4**). This is rational since the characteristic peaks of AgNPs are not easily recognizable when synthesized using natural products [36–39].
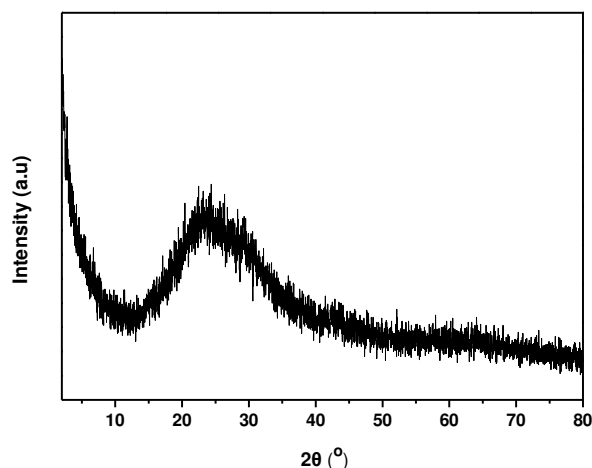
Figure 4. X-ray diffraction pattern of AgNPs.

3.4. FTIR

FTIR was used to identify the ligands that surround the nanoparticle surface. Seven absorption bands appear in the spectrum (**Figure 5**). The weak band at 3749 cm^{-1}, indicates the presence of polyphenols due to the binding of silver ions with hydroxyl group [13,40]. The band at 3438 cm^{-1} originates either from O-H or NH stretching vibration [41,42]. The band at 2961 cm^{-1} indicates C-H stretching for an alkane [43]. The small shoulder formed at 2850 cm^{-1} and 2918 cm^{-1} arises from the C-H stretching vibrations either from -C-H- or -C-H₂ – aliphatic compounds [41,42]. The band at 1636 cm^{-1} is attributed to the C-N and C-C stretching [13,40]. The sharp band at 1384 cm^{-1} corresponds to the N=O symmetry stretching typical of the nitro compounds [13,40]. The bands at 1056 cm^{-1} and 1112 cm^{-1} are assigned to stretching vibration C-O [44] . The band around 836 cm^{-1} could be attributed to out-of-plane C-H wagging vibrations, the most significant signal used in distinguishing between different types of aromatic ring substitution [45].
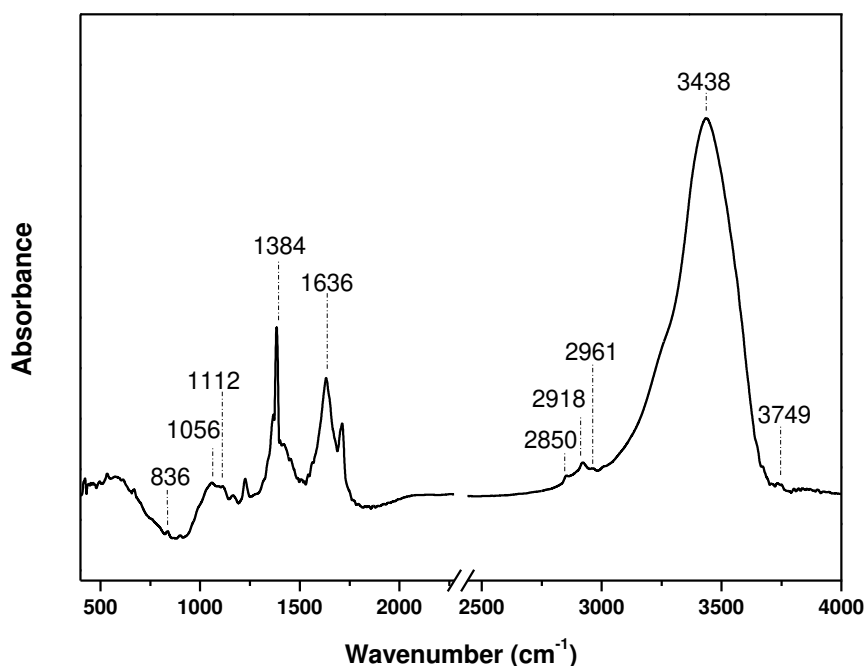
Figure 5. FTIR spectrum of AgNPs synthesized by oregano extract.

4. Conclusions

The NaOH concentration or equivalent the pH of the solution, is the main factor that influences the nanoparticle synthesis both in terms of size and quantity. The proposed model opens the possibility for size tunning of silver nanoparticles synthesized by Oregano extract. A Voigt function could be fitted on the UV-Vis spectra, with parameters the wavelength at peak maximum λ_0, the peak area A under the curve, and the Full Width at Half Maximum. A quadratic model can be fitted on both the wavelength λ_0, and the area A.

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