The Structural and Electrochemical Properties of CuCoO₂ Crystalline Nanopowders and Thin Films: Conductivity Experimental Analysis and Insights from Density Functional Theory (DFT) Calculations

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Abstract: A novel manufacturing process is presented for producing Nano powders and thin films of CuCoO₂ (CCO) material. This process utilizes three cost-effective synthesis methods: hydrothermal, sol-gel, and solid-state reactions. The resulting delafossite CuCoO₂ samples were deposited onto transparent substrates through spray pyrolysis, forming innovative thin films with a Nano crystal powder structure. Prior to the transformation into thin films, CuCoO₂ powder was first produced using a low-cost approach. The precursors for both powders and thin films were deposited onto glass surfaces using a spray pyrolysis process, and their characteristics were examined through X-ray diffraction, scanning electron microscopy, HR-TEM, UV-visible, and Electrochemical impedance spectroscopy (EIS) analyses was conducted to determine the conductivity in transversal direction of this groundbreaking material for solar cell applications. On the other hand, the sheet resistance of the samples was investigated using the four probe-method to obtain the sheet resistivity and them calculate the in plane conductivity of the samples. We also investigated the aging characteristics of different precursors with varying durations. The functional properties of CuCoO₂ samples were explored by studying chelating agent and precursor solution aging periods. Resuming, this study thoroughly discusses the synthesis of delafossite powders and their conversion into thin films, which hold potential as hole transport layers in transparent optoelectronic devices.

Keywords: Delafossite; powder; films; spray pyrolysis; EIS; conductivity; relaxation time

1. Introduction

In recent decades, there has been a significant demand for advancements in the integration of high-efficiency p-type transparent conductive oxides (TCOs) into industrial applications, particularly in the past few decades[1-5]. However, the research community is keen on exploring alternative TCOs and enhancing their electrical and optical properties to improve device efficiency[6-8]. Among various delafossite oxides, copper-based delafossite oxides stand out as promising candidates due to their desirable electrical and optical characteristics. The copper-based delafossite oxide, specifically CuMO₂ (where M represents Al, Ga, Fe, Co, Mg, Fe, Cr...), has attracted considerable attention for diverse applications such as batteries[9-15], luminescent materials[16-21], thermoelectrics[22-29], solar energy conversion and photocatalysis[30-36], hydrogen production through water splitting using photocathodes[37-42], and gas sensors[43-50]. In this study, we focus on the CuCoO2 compound, since exhibiting intriguing structural, optical and electrical properties; it has only been studied in a limited

number of publications and in most of them, the CuCoO2 compound has shown interesting properties to be treated as a possible material in photovoltaic solar cells[51-59]. Extensive studies on delafossite powders and thin films reveal that $CuCoO_2$ is a p-type semiconductor that crystallizes in two distinct structures: rhombohedral (3R)[60-62] and hexagonal (2H)[63-65]. The successful synthesis of the $CuCoO_2$ phase has been reported in only a few papers to date. Moreover, various efforts have been made to manipulate the properties of $CuCoO_2$, such as characterizing it using different printing techniques or exploring its physical and chemical properties through doping processes involving a wide range of elements[66-68]. Considerable attention has been devoted to the integration of this delafossite material into various domains using different methodologies. Beekman et al. published a study on the synthesis of undoped delafossite through ion exchange[69], while Z. Du et al. synthesized it as an electrocatalyst for the oxygen reaction[70], and also investigated the solvothermal synthesis of $CuCoO_2$[71]. Isacfranklin et al. focused on $CuCoO_2$ electrodes for supercapacitor applications[72], whereas D. Xiong et al. conducted a study on the polyvinylpyrrolidone-assisted hydrothermal synthesis of $CuCoO_2$[73]. Other researchers explored hydrogen-related aspects, such as J. Ding et al., who investigated Co_3O_4-$CuCoO_2$ composites[74].

Several studies have reported on the effects of Ca^{+2} doping in $CuCoO_2$. Z. Du et al. studied the optical and electrical properties of the material[75-77], while M. Yang et al. examined the impact of nickel doping on the structure and morphology of delafossite[78-80]. Limited research has been conducted on the transformation of delafossite powder into thin films[81-83]. Notably, the literature currently lacks any research projects specifically investigating the structural and electrochemical properties of $CuCoO_2$ crystalline nanopowders and thin films deposited on glass substrates.

Several physical and chemical processes, such as spray pyrolysis[84-86] and spin-coating[87-88], have been employed to fabricate thin films of $CuCoO_2$. However, previous attempts to produce $CuCoO_2$ thin films have predominantly utilized complex and expensive techniques, posing significant practical limitations. Currently, only a few endeavors have focused on generating $CuCoO_2$ thin films in the hexagonal phase, particularly through chemical deposition methods. Therefore, the objective of this study is to develop a simpler and more efficient chemical synthesis using three different methods based on the utilization of $CuCoO_2$ powder, yields an excellent, ordered crystalline mixture between two structures: rhombohedral and hexagonal with and good band gap values aiming to achieve thin films of this material. Comprehensive investigations will be conducted to thoroughly analyze the optical and electrical characteristics of these thin films. In particular studies of electrochemical impedance spectroscopy (EIS) analysis and simulation of delafossites based solar cell where the CuCoO2 as HTL have been reported for the first time. The conductivity of samples, prepared using the three different techniques, was measured in direct current (named in-plane conductivity) using the four-point probe and alternating current (named transversal conductivity) by electrochemical impedance spectroscopy.

Additionally, Density Functional Theory will be employed to provide a theoretical understanding of the role played by each element in the electronic structure, as well as to accurately determine the bandgap of the material.

2. Experimental part:

2.1. Synthesis of CuCoO2 Powders:

$CuCoO_2$ was produced using three different methods. The grains of copper(II) nitrate trihydrate $(Cu(NO_3)_2,3H_2O; 99\%)$ were used as Cu^+ source; cobalt(II) nitrate hexahydrate $(Co(NO_3)_2, 6H_2O; 98\%)$ were used as Co^{3+} source and sodium hydroxide (NaOH) from Sigma–Aldrich. As a solvent, deionized water (DW) was utilized. At room temperature, all chemicals were added. The amount of each precursor was 2 mM of each of the copper and cobalt sources, 4,40g of sodium hydroxide, and 70 ml of deionized water (DW), respectively. All of the precursors were measured and mixed in a reasonable amount of solvent for three hours. For hydrothermal synthesis: method 1 (sample named CuCoO2_H), the liquid was placed in a 100 ml Teflon autoclave and autoclaved at 100 °C for 24 hours; we washed the obtained solution several times with distilled water. Method 2: solid state reaction

(sample named CuCoO₂_SSR); stoichiometric amounts of the above-mentioned powders had been ground with ethanol solution (95%) for 24 h. The ground powder had been calcined at 800°C for 5 h. Method 3: sol-gel (sample named CuCoO₂_SG); copper (II) nitrate and cobalt (II) nitrate were mixed in ethylene glycol. The solution was agitated at room temperature for 1 hour in a beaker before being dried at 150°C for 5 hours. On a heated plate with a magnetic stirrer, gelation took place until a purple color emerged. The amorphous powder was heated incrementally from 50 °C to 800 °C.

2.2. Synthesis of the CuCoO₂ Thin Films:

The glass substrate cleaning process was performed by ultrasonically cleaning and drying the substrate. The dissolved nanocrystal precursor solution is then converted into a thin film. CuCoO₂/glass films were produced using spray pyrolysis technology. First, disperse 1 mg of CuCoO₂ powder in 10 mL of ethanol/water mixture and sonicated for 30 min to form homogeneous slurry. The CuCoO₂ slurry was then deposited on the glass film (Figure 1). The resulting film was annealed in air at 350° C. for 40 minutes to form a CuCoO₂ -coupled glass film.

● Cu(NO₃)₂, 3H₂O
● Co(NO₃)₂, 6H₂O
● NaOH
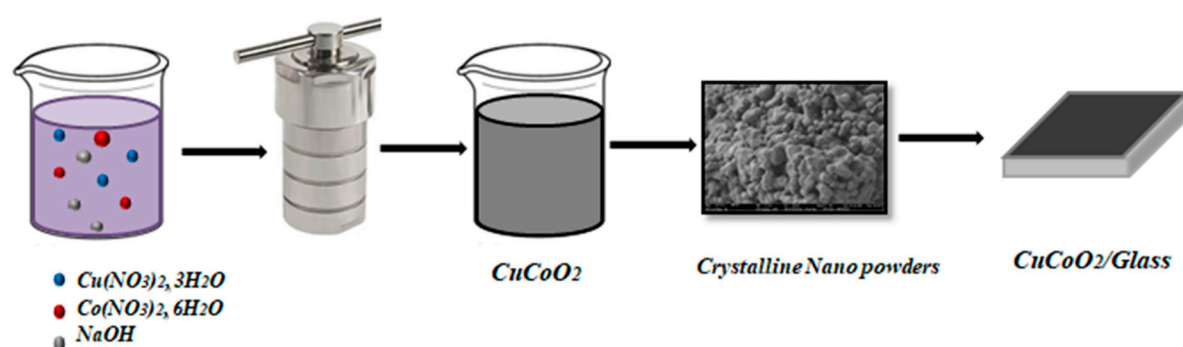
Figure 1. Schematic description of the synthesis process of the CuCoO₂/Glass film.

2.3. Electrochemical Impedance Spectroscopy (EIS) Measurements:

The conductivity was measured through the samples with a Novocontrol broadband dielectric spectrometer (BDS) equipped with an SR 830 lock-in amplifier and an Alpha dielectric interface in the frequency interval from 10^{-1} to 10^7 Hz with 0.1 V amplitude of the signal at temperatures ranging from 20 to 120°C in increments of 20°C. The samples had been airbrushed before being tested, and their thicknesses have been determined via a micrometer, averaging 10 readings from various sections of the surface. The samples were dried in a vacuum cell and placed between two gold round electrodes that served as blocking electrodes before being heated in the Novocontrol system in a neutral nitrogen-free environment. A temperature cycle from 20 to 120°C in 20°C increments was performed before collecting the dielectric spectra at every step to ensure uniformity and reduce interference from remaining water. During the testing, the electrodes were kept completely wet below 100°C and replicated a 100% relative humidity environment above 100°C in a BDS 1308 liquid device that was attached to the spectrometer and contained deionized water. To accurately control the temperature conditions, the temperature was kept constant throughout the conductivity measures (isothermal investigations) or shifted stepwise from 20 to 120°C using a nitrogen jet (QUATRO from Novocontrol), alongside a temperature error of 0.1 K throughout each frequency sweep.

The frequency dependence of complex impedance $Z^*(\omega) = Z'(\omega) + j \cdot Z''(\omega)$, yields the real component of conductivity as

$$\sigma'(\omega) = \frac{Z'(\omega) \cdot L}{\left[(Z'(\omega))^2 + (Z''(\omega))^2 \right] \cdot S} = \frac{L}{R_0 \cdot S}$$

(1)

where L and S represent the thickness and area of the sample in contact with the electrodes, and R_0 its resistance.

3. Results and Discussion

3.1. Structural analysis

Figure 2 presents the XRD diffractograms of the as-prepared samples a) $CuCoO_2\_H$, b) $CuCoO_2\_SG$, and c) $CuCoO_2\_SSR$ with a mixture of two structures: rhombohedral and hexagonal. This was confirmed by checking the database: JCPDS Map No. 074-1855 and JCPDS Map No. 021-0256. The three patterns observed in Figure 2 were identified as pure phases of $CuCoO_2$ without secondary phases. The main peak of Delafossite $CuCoO_2$ in the rhombohedral structure (110) is at $2\theta = 37.92°$, and the hexagonal structure is at $2\theta = 38.27°$.
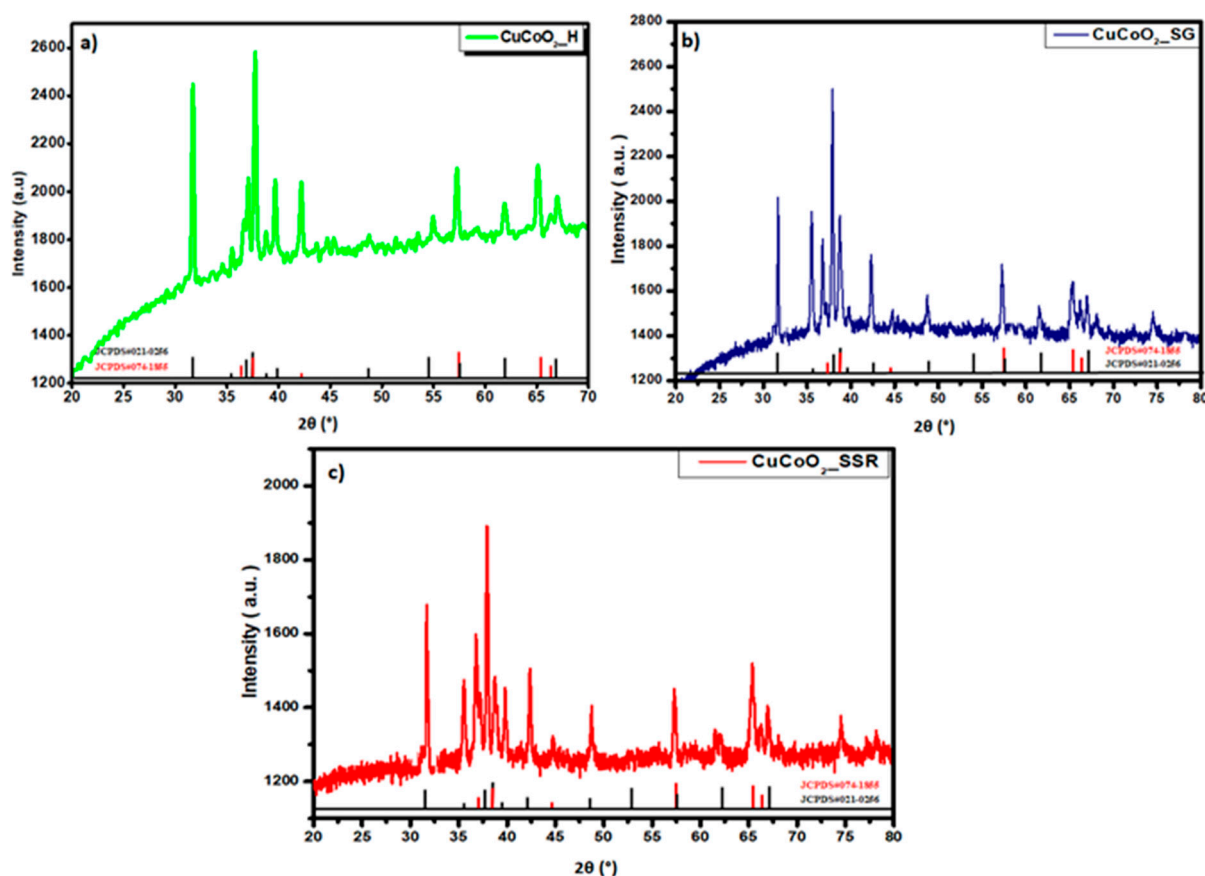
Figure 2. XRD patterns of $CuCoO_2$ nanocrystals prepared by a) hydrothermal ($CuCoO_2\_H$), b) sol-gel ($CuCoO_2\_SG$), and c) solid-state reaction ($CuCoO_2\_SSR$).

After characterization of $CuCoO_2$ nanoparticles by three methods: hydrothermal method, sol-gel method and solid-state reaction method, we examined their structure following deposition onto glass substrates and thin film synthesis using spray pyrolysis. Figure 3 shows the XRD graphs of the thin films $CuCoO_2$ on the glass substrates. We observed some of the peaks shown in the XRD pattern, which now appear in the diffractograms of the deposited film. However, the structure did not change and the material showed two structures containing Delafossite: $3R-CuCoO_2$ (JCPDS#21-0256) and $2H-CuCoO_2$ (JCPDS#74-1855). Furthermore, it is clear that each XRD pattern begins with a tablet. It belongs to the glass substrate.
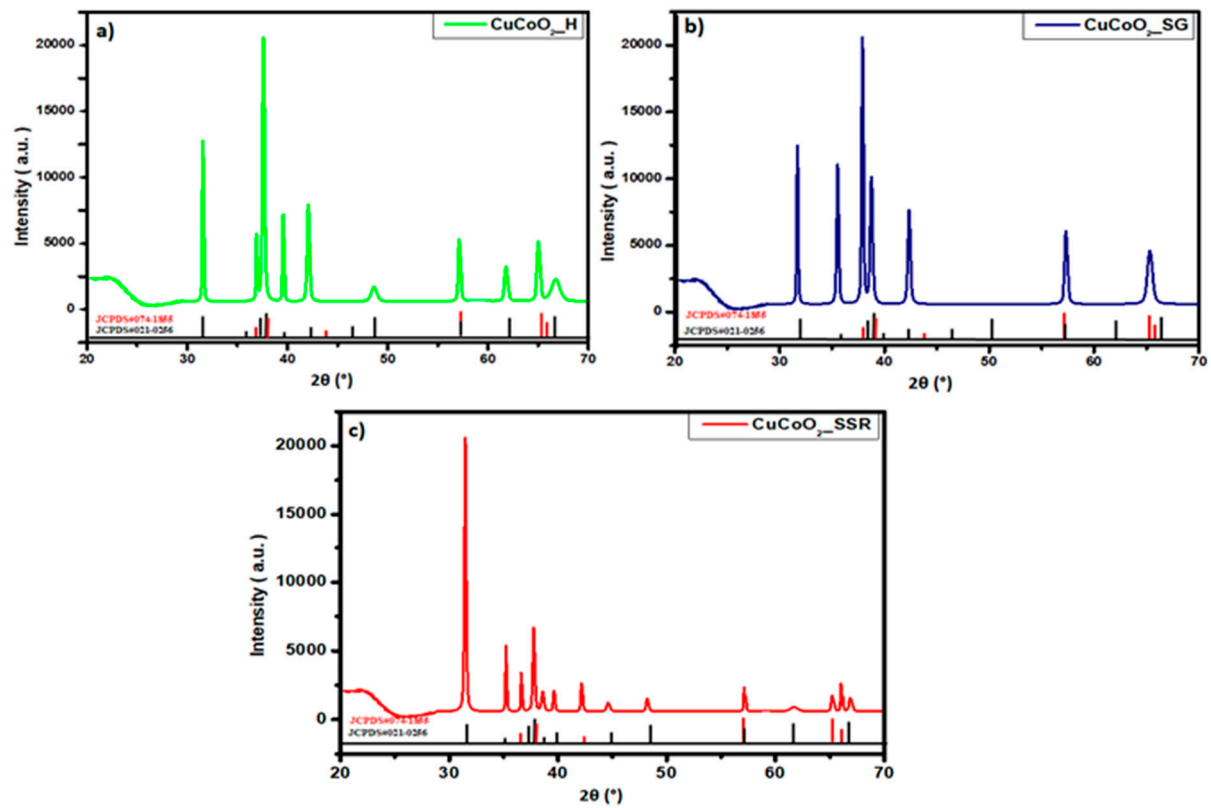
Figure 3. XRD diffractograms of films deposited on glass substrates of a) CuCoO₂_H, b) CuCoO₂_SG, and c) CuCoO₂_SSR.

3.2. FE-SEM Analysis

We are interested in the morphology and particle size distribution of our compound CuCoO₂. Observations by scanning electron microscopy (SEM) were carried out on a submicron scale on the three prepared CuCoO₂ nanocrystal powders as shown in Figure 4. A 1µm magnification is given.
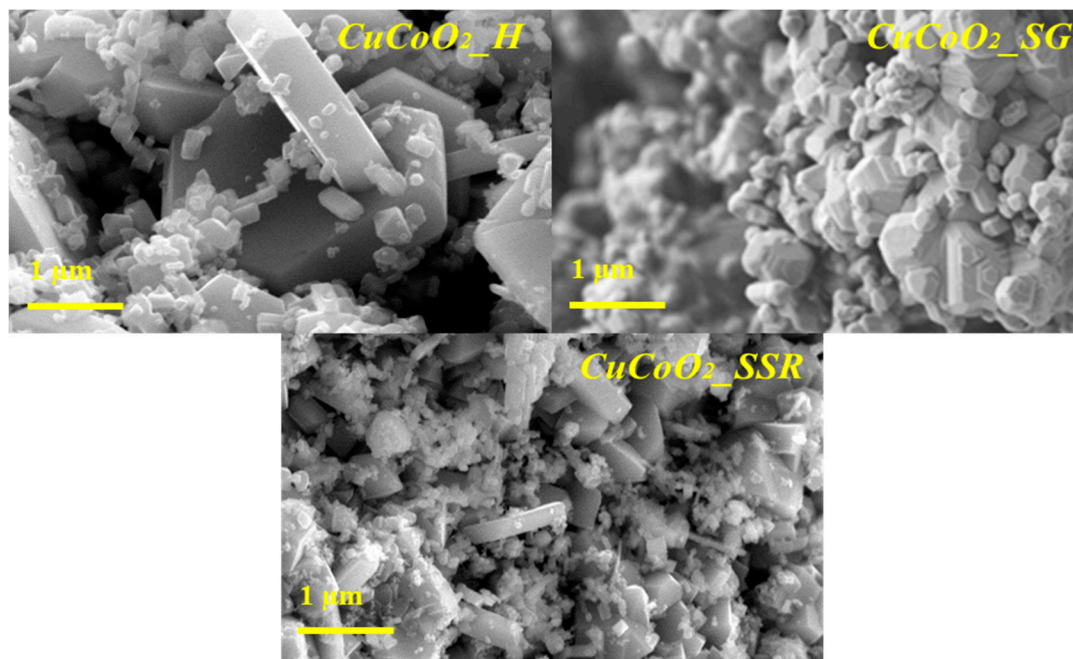
Figure 4. FE-SEM images for the three nano-crystalline powders.

FE-SEM images of the powders show the submicron CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR are in the form of a powder made up of crystals of different sizes, and they contain agglomerates of hexagonal particles and crystals like rhombohedral shapes. There have been no additional morphologies detected, which confirms the XRD results.

Figure 5 shows FE-SEM images of produced thin films comprising CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR at 1 μm. After deposition with the spray pyrolysis technique, all films had a uniform distribution of Nano crystalline particles.
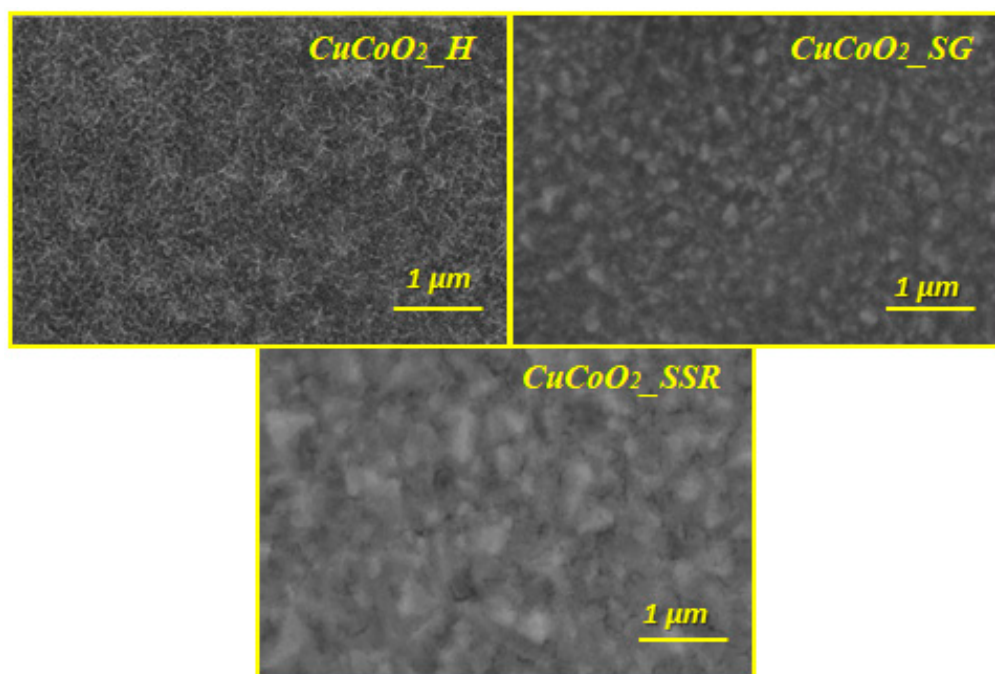
Figure 5. FE-SEM images for CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR films at 1μm.

3.3. HR-TEM Analysis

Figure 6 shows the characterization of the three samples with transmission electron microscope (TEM) characterization and the high resolution transmission electron microscopy (HRTEM) images of CuCoO₂ are illustrated in Figure 6. As a result, we could assume that the crystallinity of CuCoO2 retains a structure mostly constituted of nanocrystals smaller than 15 nm in diameter. HR-TEM d-spacings are likewise consistent with a mixture of rhombohedral and hexagonal CuCoO₂ chalcopyrite phases. The results from the FE-SEM agree with the HR-TEM images, which show clearly established small grains of some tens of nanometers.
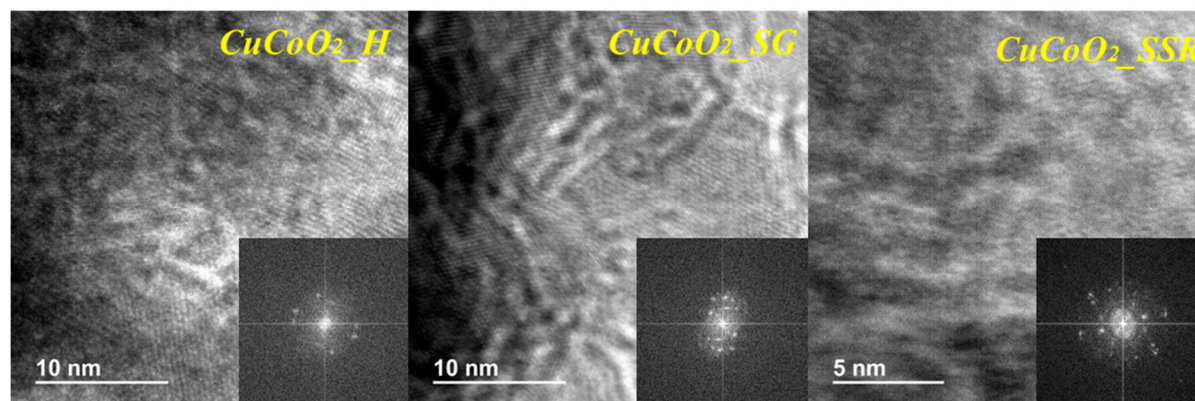
Figure 6. HRTEM images of CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR.

Figure 7 depicts the transmission electron microscopy (TEM-HR) mapping. The photos of the three $CuCoO_2$ powders reveal a good distribution of copper-cobalt oxide components, with the hydrothermal $CuCoO_2$ powder having a higher crystallinity.
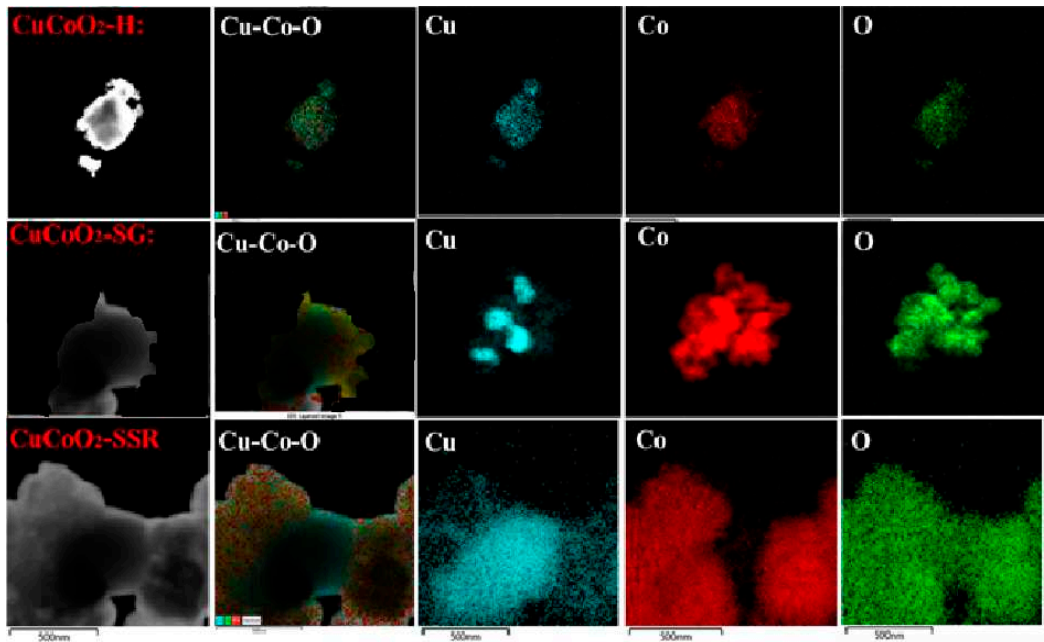
Figure 7. Element mapping analysis of $CuCoO_2$ with TEM.

4. EDX Analysis

We further investigate the chemical composition of our sample using the EDS technique. The figures and tables below show the results of the EDS analysis for the cracked surfaces of the $CuCoO_2\_H$, $CuCoO_2\_SG$, and $CuCoO_2\_SSR$ samples. From the table above the percentages of each Cu and Co in a position are approximately the same, but the percent oxygen is a bit high. This difference may be due to the oxidation of copper or cobalt.
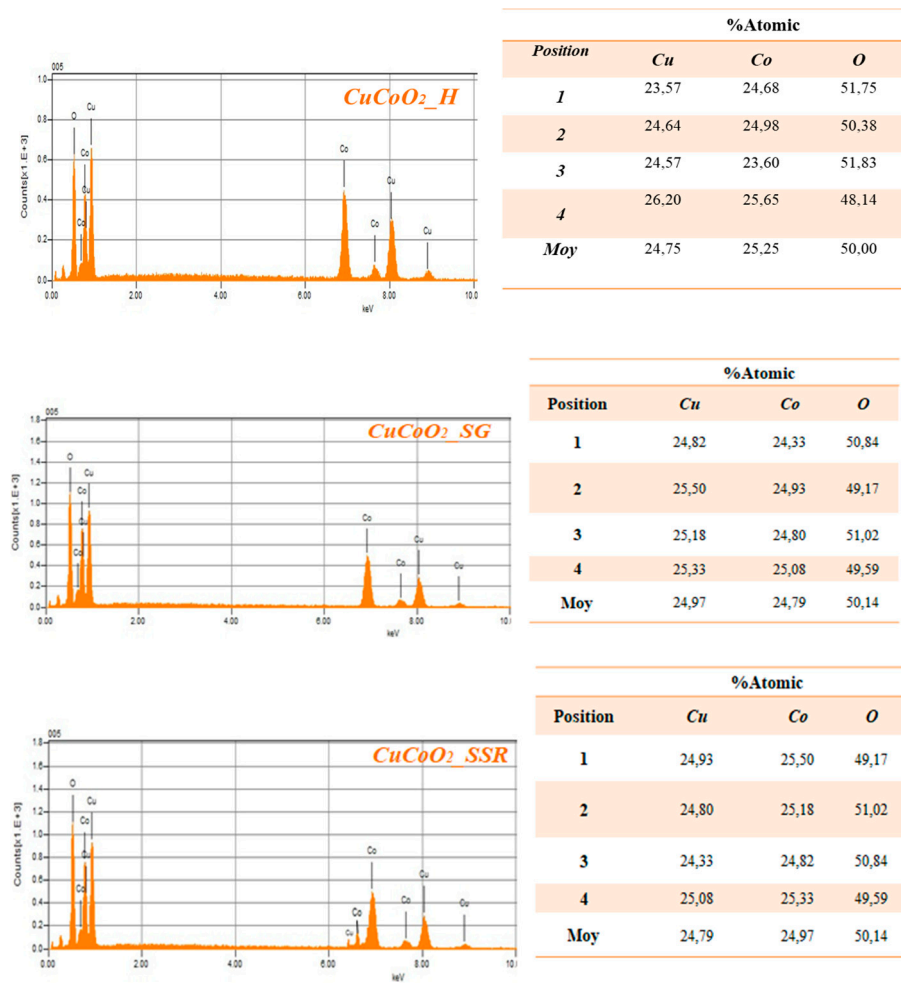
Figure 8. Cracked surface test results of three CuCoO$_2$ powders using EDS.

5. Optical Properties

The transmittance spectra have been used to examine optical qualities. In the visible area, all of the CuCoO$_2$ thin films displayed extensive absorption. The absorbance coefficients for thin film samples generated utilizing a spray pyrolysis method containing Nano crystals were determined. In the transmittance spectrum of materials, the absorption coefficient is connected with the optical energy gap or in the strong absorption zone, which may be estimated using Tauc's equation[89]

$$\alpha = \frac{A(h\vartheta - Eg)^n}{h\vartheta} \tag{2}$$

A is a constant, *h* is the Planck constant, υ is frequency, and n is an indicator of the optical absorption process. It equals 2 for directly permitted transitions and 0.5 for indirectly permitted transitions. Figure 9 shows (a) the transmission values and (b) the Tauc plot for CuCoO$_2$ thin films. The arrangement of Tauc's figure suggests that the CuCoO$_2$ thin film under deposit has a straight band gap. Eg may be estimated by extrapolating by projecting a horizontal line to the point of zero absorption coefficient ($\alpha = 0$). The band gaps were calculated by graphing (αhυ)^2 vs. energy in eV and extrapolating the linear portion of the spectrum (hυ). According to the transmission graph (Figure 9a), the powder CuCoO$_2$_SG has the highest transmission value compared to the two other powders CuCoO$_2$_H and CuCoO$_2$_SSR. The order of the delafossite powders transmission is obviously confirmed with the band gap graphs (Figure 9b).
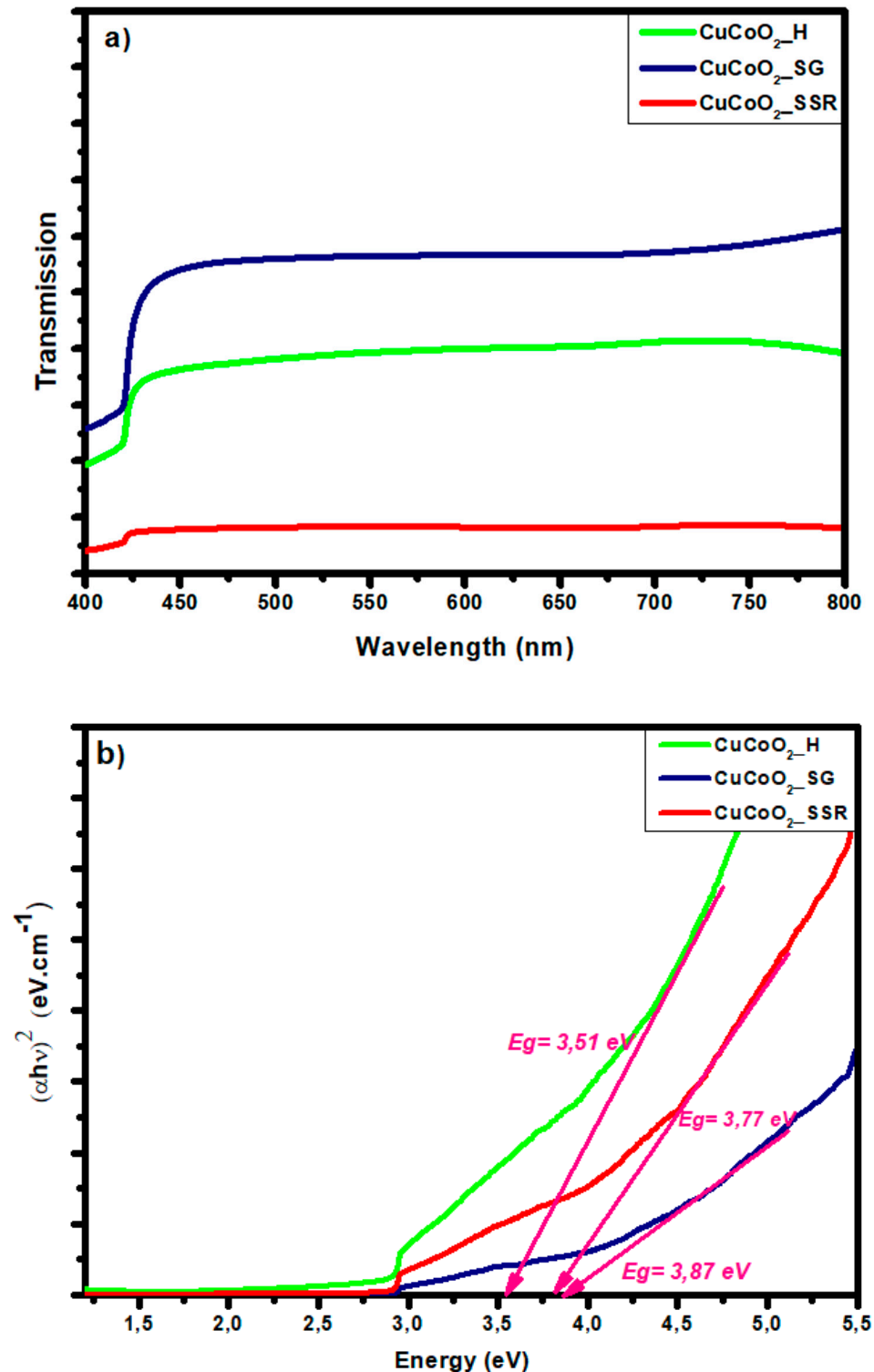
Figure 9. a) Transmission and b) Tauc's plot for CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR precursor powders.

6. In-plane conductivity measurements

The measurement of conductivity in powders and polymeric thin membranes is a complex task that is influenced by various factors, including sample casting preparations, thermal/hydrothermal treatments, relative humidity, and the cell configuration used for film resistance measurements, as well as the pressure applied between the probe electrodes[90]. The four-probe method is commonly

used to measure sample sheet resistance and estimate in-plane conductivity, but it may result in inaccuracies, particularly for materials with morphological anisotropy[91-92]. According to this method, the sheet conductivity can be calculated, from the inverse of the sheet resistivity, as

$$\sigma = \frac{1}{\rho} = \frac{1}{R_s \cdot t} \qquad (3)$$

where ρ is the sheet resistivity, R_s the in plane resistance and t is the sheet sample thickness. In our samples the thickness values are 140nm, 136nm and 142 nm for CuCoO₂_H, CuCoO₂_SG, and CuCoO₂_SSR, respectively. Figure 10 depicts the physical setup for evaluating in-plane conductivity in a typical four-probe arrangement. The resistivity may be calculated through determining the potential difference between the RE and S electrodes, that is caused by current passing via an easily identifiable connection between the WE and CE electrodes (Figure 10). The in-plane conductivity of a sample is measured throughout its surface using this approach. Nevertheless in reality, ionic conduction perpendicular to the membrane is now recognized as a strong electrochemical approach for measuring through-plane conductivity in membranes[93]. In powders used for solar cells, the four point probe method is the most often used method for assessing the electrical characteristics of conducting films[94]. This approach has been utilized to evaluate the in-plane conductivity of CuCoO₂ films on top of non-conductive substrates (in our instance, glass), which are typically created by spray pyrolysis technology the dispersions employed in this work. The four point probe method makes use of relatively basic equipment in which four probes are arranged equidistantly in a straight line and pushed against the film such is shown in Figure 10.
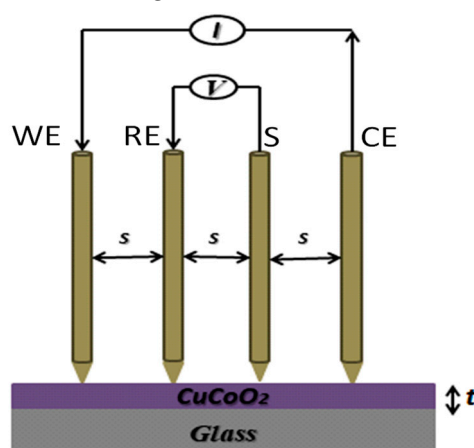
Figure 10. Schematic diagram of the four-point probe method.

The experimental procedure used is the following. The four point probe is attached to a source meter that may supply a certain current. A source meter's current (I) flows through the two outer probes, and a voltammeter can measure the voltage (V) across the two inner probes. Plotting the voltage measured for each current intensity permit us determine the sheet resistance, Rs, such as is shown in Figure 11.
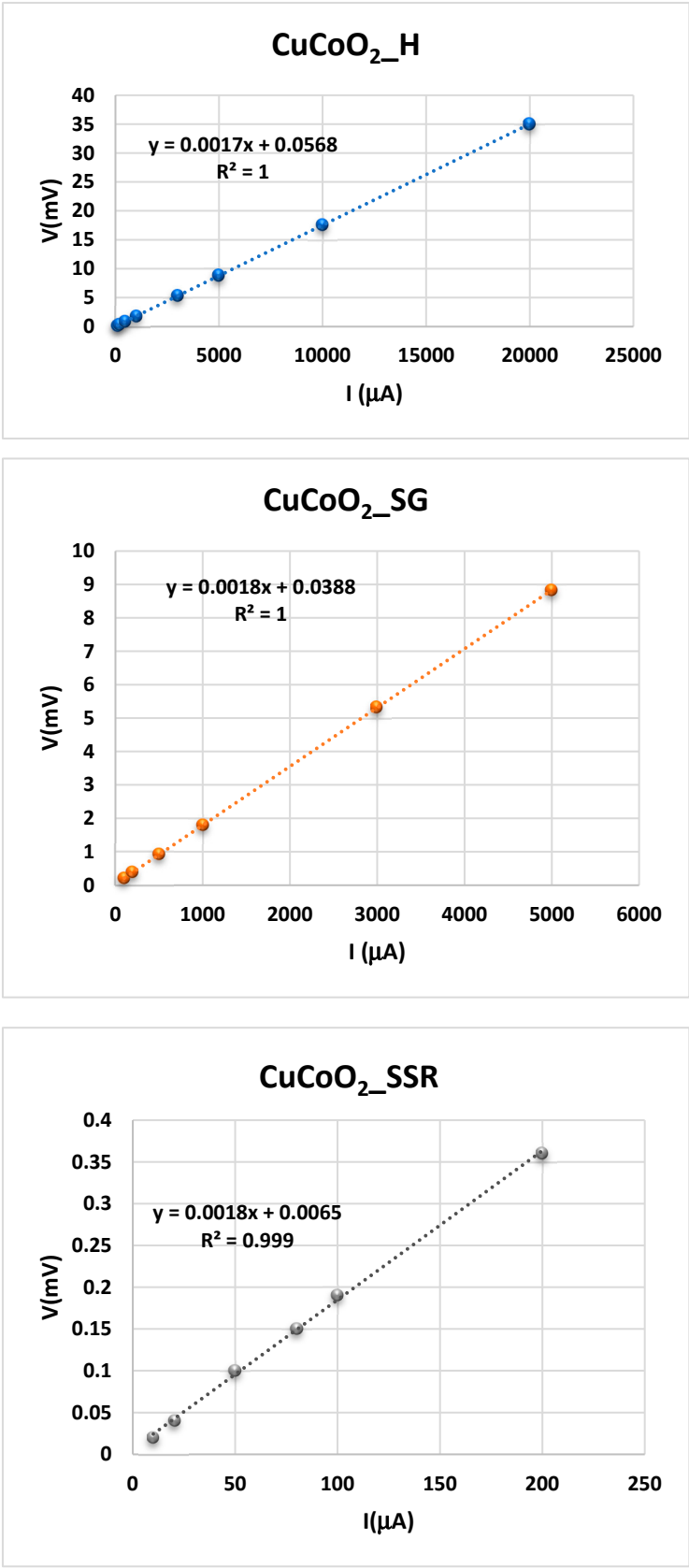
Figure 11. Variation of the potential between RE and S electrodes when a current intensity is given between WE and CE electrodes for the samples $CuCoO_2$ thin films. (●), $CuCoO_2\_H$, (●) $CuCoO_2\_SG$, and (●) $CuCoO_2\_SSR$, at ambient temperature:.

A close inspection of these figures permit us observe that sample resistance (Rs) of the CuCoO$_2$ films is constant and its values can be obtained from the slope of the experimental fit determined from the plot of the voltage versus intensity, where a clear linearity is observed for all the samples. According to the KIT used to measure the resistance of the film by means of the four points method, the value of R_s is given by

$$R_S = 4{,}532 \times \frac{V}{I} \qquad\qquad (4)$$

Being the constant 4,532 the value given for said KIT as a consequence of the geometry used in the measurement. Therefore, the in-plane conductivity, given in Eq. (3), is determined from the sheet resistance from Eq.(4). The results obtained, measured at ambient temperature, of three produced CuCoO$_2$ samples, using the four point probe technique, are presented in Table 1. The conductivity values change depending on the chelating agent and the aging duration. These findings imply that aging period and thickness change have an effect on electrical conductivities.

Table 1. The electrical properties of the CuCoO$_2$ thin films determined at ambient temperature.

Samples	*Thickness* *(nm)*	*Sheet* *Resistivity(ρ_s)* $\times 10^{-3}$ **(Ω.cm)**	*Sheet resistance* *(R_s)* **(Ω)**	*Sheet* *Conductivity* $\times 10^3$ **(S.cm^{-1})**
CuCoO$_2$_H	140±5	0.114 ±0.002	8.16±0.05	8.8±0.2
CuCoO$_2$_SG	136±5	0.105±0.002	7.70±0.04	9.5±0.2
CuCoO$_2$_SSR	142±5	0.116±0.002	8.16±0.04	8.6±0.2

7. Dielectric Spectra Analysis

The electrical impedance spectroscopy (EIS) measurements were performed on all samples to determine the conductivity measured in transversal direction, named direct current conductivity (σ_{dc}). Such measurements were carried out over a temperature interval of 20°C to 120°C in two steps to ensure reproducibility within the temperature interval. The experimental data obtained for the samples from the Novocontrol, were examined of the complex dielectric permittivity function, denoted as $\varepsilon^*(\omega,T)$, and the complex conductivity function, denoted as $\sigma^*(\omega,T)$, where j is the imaginary unit (j^2 = -1), ε_0 is the vacuum permittivity, and ω is the angular frequency of the electric field that was applied (ω =2πf). Different methods have been used to determine the dc-conductivity of a sample from dielectric spectroscopy data analysis[95-105]. In this work we have used the Bode diagram obtained from the complex dielectric spectra, where the complex conductivity is given by $\sigma^*(\omega,T)$=j ε_0 ω $\varepsilon^*(\omega,T)$, which can be expressed in terms the real and imaginary part, '(ω,T) and " (ω,T) , respectively, and the direct current conductivity σ_{dc} was calculate [104-107]. This technique was used in this study to examine data for the real component of conductivity in dry conditions by graphing conductivity (in S cm^{-1}) vs frequency (in Hz) using appropriate Bode diagrams for all temperature ranges.

The Bode diagrams for CuCoO$_2$_H, CuCoO$_2$_SG, and CuCoO$_2$_SSR delafossite materials were investigated at temperatures ranging from -20ºC to 120ºC, with increments of 20ºC, as shown in Figure 12. Additional graphs demonstrating the fluctuation of phase angle (ϕ) vs frequency at

identical temperatures are included in the **Supplementary Information**. Upon closer examination of the figures, it can be observed that the conductivity tends to a constant value (plateau) when the phase angle (ϕ) approaches zero or reaches a maximum, indicating the direct-current conductivity (σ_{dc}) of the sample. Furthermore, a decrease in conductivity with decreasing frequency was observed in the high-frequency region, along with a transition zone where the cut-off frequency ranges from 10^5 to 10^6 Hz for CuCoO$_2$_SG and CuCoO$_2$_SSR samples, and starts increasing with frequency. In the case of the hydrothermal sample **CuCoO$_2$_H**, the real part of conductivity remains constant at low frequencies until a cut-off frequency between 10^3 Hz and 10^6 Hz, after which it starts increasing with frequency. The initial process is connected to the sample's resistance/stability, but this second process is connected to the dispersion (charge transfer) caused by the charge's mobility, as the sample behaves like a capacitor. The conductivity values presented were derived using the peak frequency when the phase angle approaches 0.
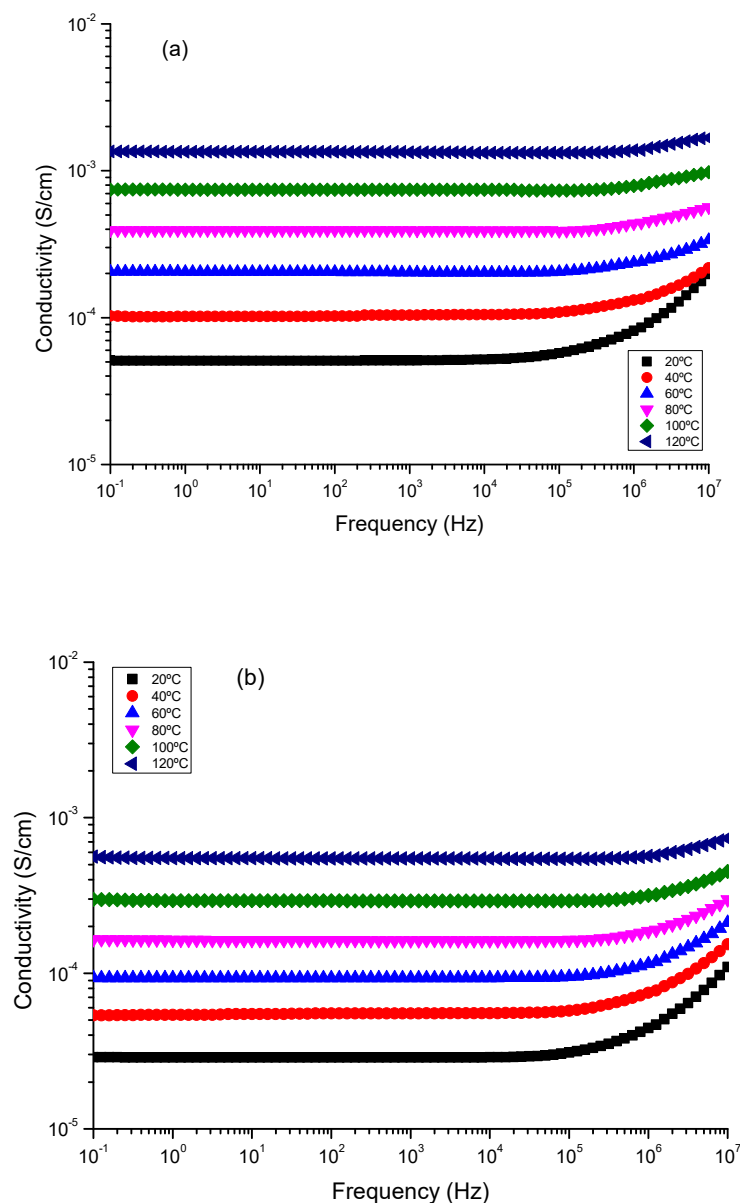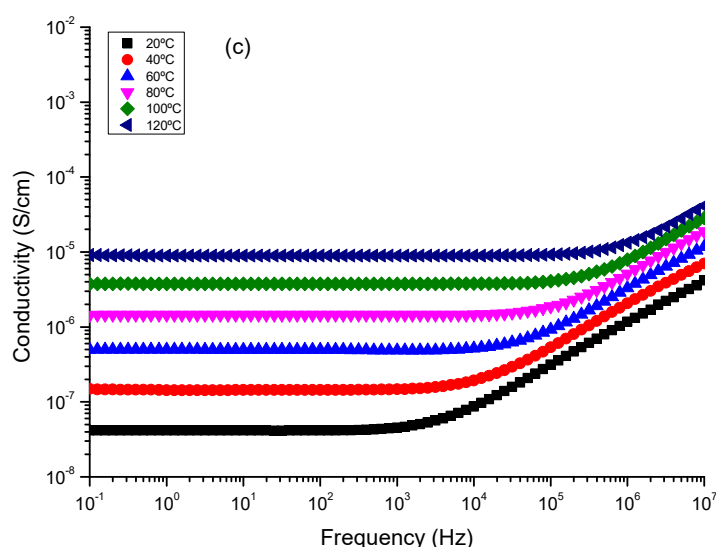
Figure 12. Bode diagrams of the conductivity for a) CuCoO$_2$_SG, b) CuCoO$_2$_SSR, and c) CuCoO$_2$_H at different temperatures.

Upon careful examination of Figure 12, it can be observed that samples **CuCoO$_2$_SG** and **CuCoO$_2$_SSR** exhibit a nearly constant conductivity across a wide range of frequencies and temperatures, that is standard behavior for a conductive material. Identical behavior has been reported in previous studies on nanocomposites of multilayer graphene in polypropylene[106]. This phenomenon is due to Debye relaxation, which occurs as a result of the mobility and redirection of dipoles and localized charges at high frequencies in response to an applied electric field and dominates direct-current conductivity[107-109]. The change in dc conductivity of the samples at various temperatures may be determined from the plateau where the phase angle approaches zero. Moreover, with rising temperature the frequency at which the point of equilibrium occurs moves to greater frequencies, suggesting thermal activation of ionic transport. The constant value of conductivity suggests that the sample solely shows resistive contribution, and the quantity measured represents the sample's electrical conductivity. For example, Figure 12 shows, at 20 °C, the trough plane conductivity values followed the trend: σ_{CuCoO2_SG} (5.2 × 10^{-5} S cm^{-1}) > σ_{CuCoO2_SSR} (2.9 × 10^{-5} S cm^{-1}) > σ_{CuCoO2_H} (4.53 × 10^{-8} S cm^{-1}). Similar trends can be observed for the other temperatures studied (for example, 40ºC, 60ºC, 80ºC, 100ºC and 120ºC). For 120ºC the conductivity values obtained, follow the trend, σ_{CuCoO2_SG} (1.4 × 10^{-3} S cm^{-1}) > σ_{CuCoO2_SSR} (5.7 × 10^{-4} S cm^{-1}) > σ_{CuCoO2_H} (1.0 × 10^{-5} S cm^{-1}), respectively. Among the three samples, the greatest proton conductivity of about 10^{-3} S cm^{-1} at 120 °C was found for the **CuCoO$_2$_SG** one order of magnitude higher than **CuCoO$_2$_SSR,** where also excellent ionic conductivities of about 10^{-4} S cm^{-1}. These results shown that the preparation method used is very relevant to get excellent results in the measured transversal conductivity, in our case around of one order of magnitude than the sample **CuCoO$_2$_H** was reached. All these values have better conductivities than CIGS:Cr crystalline Nano powders and CuInGaSe$_2$ (CIGS) chalcopyrite thin films doped with Cr varying its concentrations[110].

From the plot showed in Figure 13, we observe that dc-conductivity of all mixtures increases with temperature increase and follow an Arrhenius behavior for the thermal activation energy. The measurements of the activation energy calculated from the slopes follow the trend E$_{act}$ (**CuCoO$_2$_SSR**) = 27.4 kJ/mol< E$_{act}$ (**CuCoO$_2$_SG**) = 30.8 kJ/mol < E$_{act}$ (**CuCoO$_2$_H**) =52.3 kJ/mol, respectively.
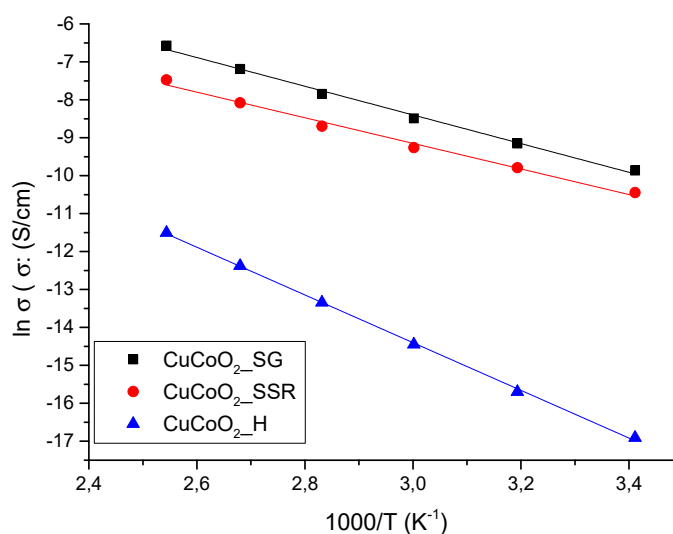
Figure 13. Temperature dependence of conductivity determined using Bode graphs for all samples examined.

These results indicate that the thin films prepared from hydrothermal synthesis have higher activation energy and lower conductivities than the samples prepared from: a) solid state reaction where stoichiometric amounts of the above-mentioned powders have been grinded with ethanol solution and after had calcined at temperature 800°C for 5h. b) Sol gel; copper (II) nitrate and cobalt (II) nitrate were mixed in ethylene glycol and after the gelation occurred on a magnetic-stirring hot plate till a purple color appeared become darker. The amorphous powder was heated until 800 °C, by the step of 50 °C.

Figure 14, shows the relationship between the relaxation time obtained from the cut-off frequency where the conductivity changes from a constant value in the bode diagram (plateau) to an increasing with the frequency when the frequency increase for all temperatures. In such circumstances, such as is observed, more focused, in case of $CuCoO_2$_H (see figure 12-C) for all the temperatures, we can say according the power low model, that real part of the conductivity $\sigma'(\omega,T)$ can be expressed in terms of dc-conductivity σ_{dc} and hopping diffusion rate of protons $\omega_H \approx 1/\tau$ (in this case) as[111]

$$\sigma'(\omega, T) = \sigma_{dc}\left[1 + \left(\frac{\omega}{\omega_H}\right)^n\right] \tag{5}$$

where n is an exponent compress between 0 and 1, related to interactions between mobile ions (H^+ in our case) and dimensionally of conduction pathway[112], such is happen in polymer electrolytes of P[VBTC][Cl]$_{80}$-ran-PMMA$_{20}$ at different temperatures (303K to 363K), and P[VBTC][TFSIl]$_{80}$-ran-PMMA$_{20}$ at (308K to 378K), respectively[113]. From the fit of the real part of the conductivity showed in the Bode diagrams showed in Figure12 we have obtained the values of sample relaxation time that we are plotted in Figure 14 for each temperature.
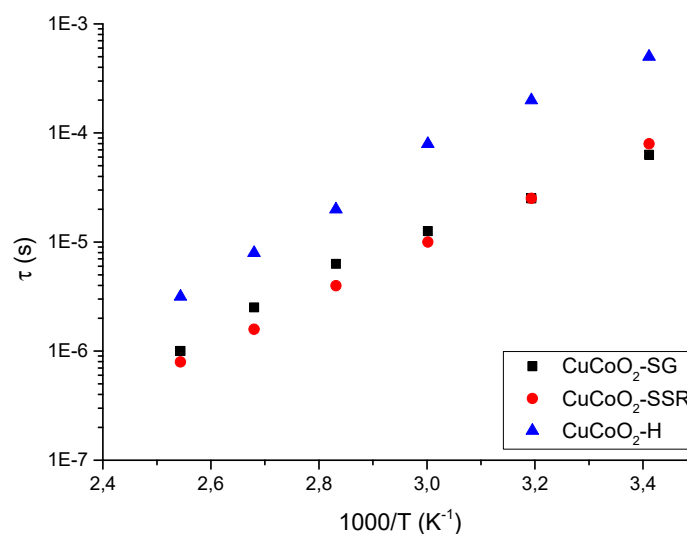
Figure 14. Variation of the sample relaxation time against the reciprocal of temperature for the delafossites investigated.

From Figure 14 we can see that the relaxation time follow an Arrhenius behavior in all the powders studied, but is interesting to observe that samples prepared using method 2 ($CuCoO_2$_SSR and method 3 ($CuCoO_2$_SG), have a relaxation time around one order of magnitude smaller than sample prepared using the method 1 ($CuCoO_2$_H). That's mean that the method to produce powders to build thin films is very important if we want enhanced optical and electrical properties. In these results we observe that delafossites are extremely sensitive on a wide variety of parameters, in particular the method used in its preparation, and the results indicate that these materials have the potential destined for use in thin-film solar cells.

8. Theoretical insight

Theoretical simulations are critical in the understanding of the properties of systems at the atomic level. Therefore, first-principles calculations were performed to complement the experimental results with an insight on the electronic structure and properties of. For this purpose the atomic positions were optimized and the electronic properties were computed within the Density Functional Theory (DFT) approach, based on the framework of the generalized Kohn–Sham scheme[114-117] in combination with the projector augmented-wave (PAW) method[118] and the Heyd–Scuseria–Ernzerhof hybrid functional with the modified fraction of screened short-range Hartree–Fock exchange (HSE06)[119-121] as implemented in the Vienna ab initio simulation package (VASP)[122-123]. Hybrid functionals allow a more accurate description of electronic properties of some systems compared to simple DFT calculations with a generalized gradient approximation for the exchange and correlation term in the Kohn–Sham scheme. The more accuracy, however, is reached at the cost of a higher computational time needed to achieve convergence.

To model the $CuCoO_2$ in the tetragonal phase, a unit cell with 12 atoms ($Co_3Cu_3O_6$) was used, while for the hexagonal cell a smaller cell with 8 atoms ($Co_2Cu_2O_4$) was needed. The electronic wave functions were expanded in a plane wave basis setup to a kinetic energy cutoff of 400 eV, and the atomic positions were optimized using the conjugate gradient method up until the forces on each atom were less than 0.01 eV A^{-1} and the energy convergence less than 10^{-8} eV for the optimization, and less than 10^{-5} eV for the calculations with the hybrid functional. For the Brillouin zone integration, a 12 × 12 × 6 Monkhorst–Pack scheme k-point mesh was used[124,125] both for the optimization and electronic structure calculation.

Figure 15 shows the atom positions and geometrical structure of the converged unit cells. The delafossite structure of this ternary oxide can be appreciated fairly. In the tetragonal structure an alternate stacking of O-Cu-O dumbbells parallel to the z axis, and a layer of Co-centered octahedral in the xy planes. The stacking follows an ABCABC pattern, forming a trigonal system with lattice parameters a = 2.86 Å and c = 16.98 Å corresponding to a rhombohedral Bravais lattice of volume 120.1 Å^3. In the hexagonal cells, the same layers of O-Cu-O dumbbells and Co-centered octahedral are observed, but with an ABAB stacking pattern. The lattice constant a = 2.83 Å is practically identical, but c = 11.30 Å is significantly lower, as well as the unit cell volume of 78.47 Å^3. However, these differences are due to the fact that a smaller distance is enough to represent the structure, because of the stacking sequence. Hence, important distances within the unit cell, such as the distance between O atoms in the O-Cu-O dumbbells (3.71 Å for the tetragonal 3.69 Å in the hexagonal) or the distance from cornered O to centered Co in the octahedral (1.92 Å for the tetragonal and 1.91 Å in the hexagonal) are practically the same.
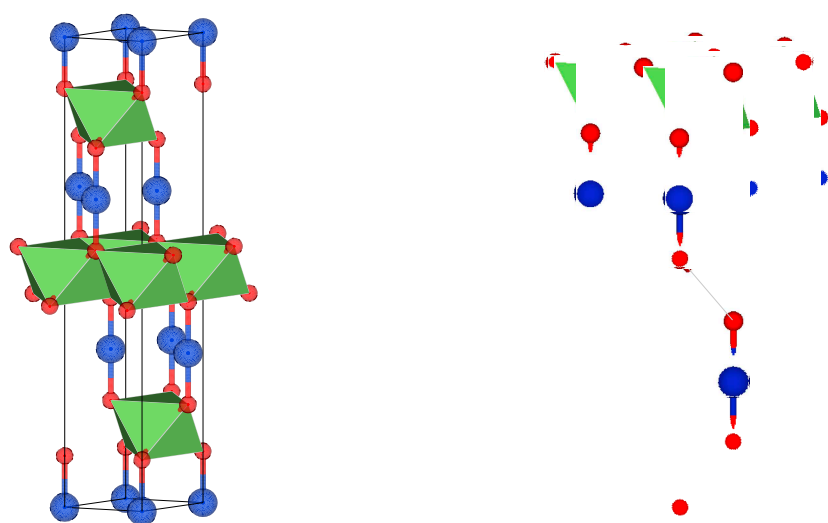
Figure 15. Trigonal and hexagonal structures of $\boldsymbol{CuCoO_2}$. The atoms of Cu are represented in blue; O atoms are presented in red, and the metal atoms lies inside the green octahedral.

The bands structure and density of states, for the tetragonal phase can be appreciated in Figure 16. Tough the inclusion of hybrid functional allows the possibility of avoiding the typical underestimation of gaps in pure DFT approaches; the bandgap values obtained seems to remain underestimated: 2.31eV for the tetragonal structure and 2.34 for the hexagonal one. However, the structure of the bands and density of states form of the curves is usually well described in this type of calculations, and despite the small difference in bandgap values for both structures, the form of the density of states is practically the same, which is why only one of them is plotted.

The obtained gap is an indirect one, which can be appreciated by simple inspection of the bands structure in the left part of Figure 16. The right part shows the density of states in which it is appreciable that the edge of the covalent band is apparently dominated by Cu-3d states, and O states show a wide dispersion through the band, but the presence of Co-3d states is high in this region, and except for the peak at the edge of the band, is comparable with the contribution from Cu states. Also, the first peak in the conduction band is mainly due to Co-3d states followed by a second peak with a structure similar to the one that can be found in $CuAlO_2$ and $CuGaO_2$ as it has been previously reported[109,110].
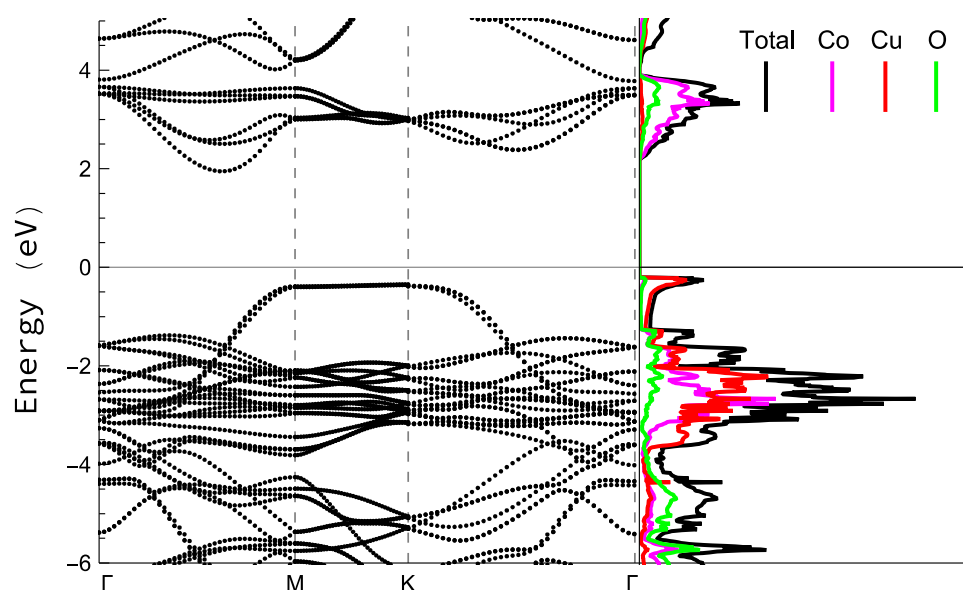
Figure 16. Bands structure (left) and states' density (right) acquired by HSE simulations for the tetragonal structure. Energies are expressed with respect to the Fermi Energy. Black line accounts for total DOS in the right part, and coloured lines accounts for the contribution of every species to the density of states: Co atoms are represented by pink, Cu atoms by red, and O atoms by green.

9. Conclusions

In summary, the preparation of delafossite $CuCoO_2$ using three cost-effective techniques yields a well-ordered crystalline mixture comprising two structures: rhombohedral and hexagonal, exhibiting similar properties and favorable band gap values. To ensure the development of an effective coating solution, careful treatment of the resultant nano-sized precursor powder is necessary. The desired thin films, namely $CuCoO_2$_H (produced via hydrothermal method), $CuCoO_2$_SG (produced via sol-gel method), and $CuCoO_2$_SSR (produced via solid-state reaction method), were successfully obtained through the spray pyrolysis process. Electrochemical impedance spectroscopy measurements reveal that the Sol-Gel method yields films with superior conductivity compared to the other preparation methods. As a result, $CuCoO_2$ thin films hold significant potential for solar cell applications. This is supported by the examination of the electronic properties of the rhombohedral $CuCoO_2$ through theoretical simulations. Furthermore, electrochemical impedance spectroscopy measurements demonstrate that the copper cobalt delafossites prepared using different synthesis methods have the potential to serve as semiconductor materials. The conductivity values, measured through plane, increase with temperature as expected: $\sigma_{CuCoO2\_SG} > \sigma_{CuCoO2\_SSR} > \sigma_{CuCoO2\_H}$. Among the three samples, $CuCoO_2$_SG exhibits the highest conductivity of approximately 10^{-3} S cm^{-1} at 120 °C, which is one order of magnitude greater than $CuCoO_2$_SSR, while still maintaining good ionic conductivities of approximately 10^{-4} S cm^{-1}. This value is approximately two orders of magnitude higher than that of $CuCoO_2$_H. These conductivity measurements, obtained through impedance spectroscopy, are in agreement with the values determined using the four-probe method, where resistivity follows the trend of $\rho_{CuCoO2\_SG} < \rho_{CuCoO2\_SSR} \approx \rho_{CuCoO2\_H}$. Lastly, our study demonstrates that samples prepared using method 2 ($CuCoO_2$_SSR) and method 3 ($CuCoO_2$_SG) have relaxation times approximately one order of magnitude smaller than samples prepared using method 1 ($CuCoO_2$_H). This result emphasizes the significant impact that the synthesis method and sample preparation can have on producing powders for building thin films to enhance optical and electrical properties, particularly for solar cell applications.

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