

Article

Thermodynamic Stability of Fenclorim and Clopyralid

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Abstract: The present work reports an experimental thermodynamic study of two nitrogen heterocyclic organic compounds, fenclorim and clopyralid, that have been used as herbicides. The sublimation vapor pressures of fenclorim (4,6-dichloro-2-phenylpyrimidine) and of clopyralid (3,6-dichloro-2-pyridinecarboxylic acid) were measured, at different temperatures, using a Knudsen mass-loss effusion technique. The vapor pressures of both crystalline and liquid (including supercooled liquid) phases of fenclorim were also determined using a static method based on capacitance diaphragm manometers. The experimental results enabled accurate determination of the standard molar enthalpies, entropies and Gibbs energies of sublimation for both compounds and of vaporization for fenclorim, allowing a phase diagram representation of the (p, T) results, in the neighborhood of the triple point of this compound. The temperatures and molar enthalpies of fusion of the two compounds studied were determined using differential scanning calorimetry. The standard isobaric molar heat capacities of the two crystalline compounds were determined at 298.15 K, using drop calorimetry. The gas phase thermodynamic properties of the two compounds were estimated through ab initio calculations, at the G3(MP2)//B3LYP level, and their thermodynamic stability was evaluated in the gaseous and crystalline phases, considering the calculated values of the standard Gibbs energies of formation, at 298.15 K.

Keywords: Fenclorim; Clopyralid; Vapor pressures; Phase transitions; Heat capacities; Thermodynamic stability

1. Introduction

Herbicides, also known as weedkillers, are a broad class of pesticides that are used to control or manipulate undesirable vegetation, such as grasses and weeds, that may compromise the growth and profitability of crops [1]. They are efficient and cost-effective means of controlling nuisance plants when compared to hoeing, mowing, cultivation, or hand pulling and play a key role in farmland management and in reducing labor intensity [2-4]. While these chemicals help improve and rise crop yields, they also pose risks to the crops themselves which can be sensitive to these agents. To overcome this issue herbicide safeners have been developed to increase crop selectivity [5,6]. These are synthetic substances with the ability to protect grass crops from herbicide injury by a physiological or molecular mechanism, without reducing herbicidal activity on target weed species [3]. The safener fenclorim (Fig. 1a) was designed to reduce the damage to rice (*Oryza sativa* L.) caused by herbicides belonging to the class of chloroacetanilides [7-9]. To ensure the safety of rice at an early stage, fenclorim is often formulated with pretilachlor [2-chloro-2',6'-diethyl-N-(2-propoxyethyl)-acetanilide], which is one of the most widely used herbicides in rice-producing countries but poses a phytotoxicity risk for this cereal. [7,9-11]. Fenclorim protects the rice from damage caused by pretilachlor mainly by speeding up the metabolism of this herbicide [7,9]. The pyridinecarboxylic acid herbicide Clopyralid (Fig. 1b) has been used effectively to control broadleaf weeds in pastures, turf, and in some agricultural crops such as barley, sugar beets, wheat, mint and oats [12]. It is a synthetic plant

Compound	CASNR	Source	Minimum initial purity	Purification Method	Final mass fraction purity	Analysis method ^a	Water content ^b (%)
Fenclorim	3740-92-9	TCI	0.999 ^c	Sublimation ^d	0.9993	GC	0.02 ± 0.01
Clopyralid	1702-17-6	TCI	> 0.98 ^e		0.9989		0.04 ± 0.01

^aGas-liquid chromatography with flame ionization detector (FID). ^bDetermined using Karl Fisher coulometric titration.

^cAnalysis certified by the manufacturer. ^dUnder reduced pressure. ^eMinimum purity degree announced by the supplier.

2.2 Thermal analysis

The Hitachi-DSC7020 heat flow calorimeter was used to verify the absence of possible phase transitions in the crystalline phase of fenclorim and clopyralid and to determine their onset temperatures and enthalpies of fusion. Four independent runs were carried out using fresh samples (not melted before) sealed in airtight aluminium crucibles for each compound. The samples were scanned at 2.0 K·min⁻¹ from $T = 298.15$ K to a temperature (20 to 25) K higher than their temperature of fusion under a controlled nitrogen flux that was used to avoid eventual contamination of the calorimeter. Calibration of the calorimeter was performed using the following high purity reference materials: benzoic acid (NIST SRM 39j), indium (Sigma-Aldrich, mass fraction >0.99999) and tin (Sigma-Aldrich, mass fraction >0.99999). The standard uncertainties of the calibration results are $u(T/K) = 0.34$ and $u[\Delta_{\text{cr}}^1 H_{\text{m}}^{\circ}(T_{\text{fus}})]/\text{kJ}\cdot\text{mol}^{-1} = 0.66$. The onset temperatures of fusion, T_{fus} , and molar enthalpies of fusion, $\Delta_{\text{cr}}^1 H_{\text{m}}^{\circ}(T_{\text{fus}})$, determined in each run are reported in Table S2, together with the derived values of the molar entropies of fusion, $\Delta_{\text{cr}}^1 S_{\text{m}}^{\circ}(T_{\text{fus}})$, and the available literature results.

2.3 Heat Capacity Drop Calorimetry

A high-precision heat capacity drop calorimeter [17-20] was used to measure the heat capacities of fenclorim and clopyralid, at $T = 298.15$ K. The calibration of the calorimeter was performed using sapphire (α -Al₂O₃ pellets, NIST-RM 720), $C_{p,m}^{\circ}(\alpha\text{-Al}_2\text{O}_3) = 79.03 \pm 0.08$ J·K⁻¹·mol⁻¹ [21], based on a single drop temperature step ($\Delta T = 10.00$ K) from $T_i = 303.15$ K to the final temperature $T_f = 293.15$ K. The accuracy of the calorimeter and methodology for the measurements of the heat capacities of crystalline and liquid compounds was evaluated based on the measurements of benzoic acid (Calorimetric Standard NIST 39j) and hexafluorobenzene as test substances [19-23], providing reliable high-quality heat capacity data for several compounds [24-27]. The standard isobaric molar heat capacity, $C_{p,m}^{\circ}$, at 298.15 K of each experiment for fenclorim and clopyralid, together with the mass of sample used in at least two independent series of drop experiments are presented in Table S3.

2.4. Vapor pressure measurements

2.4.1. Knudsen mass-loss effusion method

The sublimation vapor pressures of fenclorim and clopyralid were measured at different temperatures using the Knudsen mass-loss effusion method. The apparatus used in this work allows the simultaneous operation of nine effusion cells contained in cylindrical holes inside three aluminium blocks, controlled at different temperatures. Three cells with different effusion orifice sizes are inserted in the holes of each block. The effusion orifices made by Goodfellow™ on platinum foil of (0.0125 ± 0.001) mm thickness, have the following areas: $A_o(A1) = A_o(A2) = A_o(A3) = (0.636 \pm 0.004)$ mm², $A_o(B1) = A_o(B2) = A_o(B3) = (0.785 \pm 0.004)$ mm², $A_o(C1) = A_o(C2) = A_o(C3) = (0.985 \pm 0.004)$ mm², where the uncertainties were calculated by the Root Sum Square (RSS) method. The Clausing factors of the effusion orifices were calculated as $w_o = 1/\{1 + (l/2r)\}$, where l is the thickness of the platinum foil and r is the radius of the orifices, yielding the results 0.986, 0.988 and 0.989

for w_0 of the orifices of the series A, B and C, respectively. Further details of this set-up, procedure and testing have been described before in detail [28].

In each effusion experiment, the mass of the sublimed sample, Δm , was determined by weighing each effusion cell (± 0.01 mg), before and after a convenient effusion time, t , in a system evacuated to a pressure near 1×10^{-4} Pa. At each temperature T , the vapor pressure p of the crystalline sample is calculated using Eq. (1),

$$p = \frac{\Delta m}{A_0 w_0 t} \left(\frac{2\pi RT}{M} \right)^{0.5} \quad (1)$$

where M is the molar mass of the effusing vapor and R is the molar gas constant ($8.314462618 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ [29]). The standard uncertainties of the temperatures and vapor pressure measurements were estimated as $u(T/\text{K}) = 0.01$ and $u(p/\text{Pa}) = 0.02$.

2.4.2. Static method based on capacitance diaphragm manometers

A static method was used to determine the vapor pressures of both crystalline and liquid phases (including supercooled liquid) of fenclorim (the limited amount of purified sample of clopyralid did not allow measurements of vapor pressure using this technique). This static apparatus is based on capacitance diaphragm gauges, previously tested and depicted in detail [30-32]. The capacitance diaphragm absolute gauge used in this work was obtained from MKS Instruments, Inc. and operate at self-controlled constant temperature: Baratron 631A11TBFP ($T_{\text{gauge}} = 473 \text{ K}$) capable of measuring pressures over the range (0.5 to 2.6×10^2) Pa and temperature of the condensed sample from (253 to 463) K. The uncertainty of the temperature measurements is estimated to be $u(T/\text{K}) = \pm 0.01$ and the expanded uncertainties (0.95 confidence level, $k = 2$) of the pressure measurements are adequately described by the expression $U(p/\text{Pa}) = 0.01 + 0.0050 (p/\text{Pa})$. Before starting the vapor pressure measurements, the samples are conveniently outgassed until repeated measurements at a selected temperature deliver consistent pressure results after escaping eventual traces of volatile impurities, including water.

2.5. Computational Chemistry Calculations

Standard ab initio molecular orbital calculations of fenclorim, clopyralid and all the auxiliary molecules considered in this work were performed with Gaussian 09 software package [33], using the G3(MP2)//B3LYP composite method [34]. This methodology is a variation of the Gaussian-3 (G3) theory [35] that uses the B3LYP density functional method for geometry optimization and zero-point energies. Information about that method is detailed described in the literature [34].

3. Results and Discussion

3.1. Thermodynamic properties of phase transitions

The vapor pressures of the two compounds studied, measured using the effusion or the static methods, are listed in Table 2, where the effusion pressures are the mean of the results determined through the different effusion orifices at each temperature, presented in detail in Tables S4 and S5 (Supporting Material). The vapor pressures of the crystalline phase of fenclorim and clopyralid were measured, respectively, in the temperature ranges $T = (311.1$ to $333.2) \text{ K}$ and $T = (334.1$ to $356.4) \text{ K}$, using the Knudsen mass-loss effusion technique. Moreover, the sublimation vapor pressures of fenclorim were measured in the temperature interval (326.0 to 365.5) K using a static method based on capacitance diaphragm manometers. This technique was also used to determine the vaporization vapor pressures of this compound in the temperature range $T = (338.8$ to $399.9) \text{ K}$. The truncated form of Clarke and Glew equation [36], Eq. (2), was used to fit the experimental (p, T) data.

$$R \ln \left(\frac{p}{p^\circ} \right) = -\frac{\Delta_{\text{cd}}^{\text{g}} G_{\text{m}}^{\circ}(\theta)}{\theta} + \Delta_{\text{cd}}^{\text{g}} H_{\text{m}}^{\circ}(\theta) \left(\frac{1}{\theta} - \frac{1}{T} \right) + \Delta_{\text{cd}}^{\text{g}} C_{p,\text{m}}^{\circ}(\theta) \left[\left(\frac{\theta}{T} \right) - 1 + \ln \left(\frac{T}{\theta} \right) \right] \quad (2)$$

In this equation, p° is a selected reference pressure ($p^\circ = 10^5$ Pa in this work), p is the vapor pressure at the temperature T , θ is a reference temperature (in this work, unless stated otherwise, $\theta = 298.15$ K), R is the molar gas constant and $\Delta_{\text{cd}}^{\text{g}} G_{\text{m}}^{\circ}$, $\Delta_{\text{cd}}^{\text{g}} H_{\text{m}}^{\circ}$ and $\Delta_{\text{cd}}^{\text{g}} C_{p,\text{m}}^{\circ}$ are thermodynamic properties of sublimation or vaporization (respectively, the standard Gibbs energy, the enthalpy and the isobaric heat capacity).

Table 2. Vapor pressure results.^a

T/K	p/Pa	$100\Delta p/p^{\text{b}}$	T/K	p/Pa	$100\Delta p/p^{\text{b}}$	T/K	p/Pa	$100\Delta p/p^{\text{b}}$
Fencloirim								
<i>Crystalline phase</i> (Knudsen effusion method) ^c								
311.11	0.106	0.2	319.31	0.279	0.0	327.25	0.677	-0.5
313.29	0.136	-1.2	321.27	0.355	1.6	329.11	0.833	0.0
315.23	0.172	-0.8	323.10	0.430	0.1	331.27	1.051	0.0
317.10	0.220	1.8	325.28	0.538	-1.8	333.23	1.301	0.6
<i>Crystalline phase</i> (Static method)								
326.01	0.58	-1.1	339.81	2.54	0.8	353.58	9.67	0.8
327.95	0.72	-0.7	341.81	3.09	0.3	355.57	11.51	-0.2
329.90	0.89	-0.7	343.77	3.75	0.2	357.55	13.79	-0.3
331.91	1.11	-0.1	345.72	4.57	0.8	359.49	16.35	-0.8
333.91	1.37	-0.1	347.70	5.52	0.5	361.50	19.64	-0.4
335.88	1.69	0.3	349.70	6.69	0.5	363.49	23.32	-0.9
337.86	2.09	1.1	351.65	8.02	0.2	365.53	28.05	0.2
<i>Liquid phase</i> (Static method)								
338.80	4.40 ^d	-0.1	360.41	20.97 ^d	0.8	382.19	80.91	-0.4
340.84	5.13 ^d	-0.3	362.38	23.91 ^d	0.8	384.18	90.85	-0.4
342.76	5.95 ^d	0.0	364.43	27.33 ^d	0.7	386.23	102.8	0.2
344.71	6.92 ^d	0.5	366.43	30.85 ^d	-0.1	388.06	114.7	0.8
346.72	8.00 ^d	0.2	368.43	35.37	0.7	390.07	126.6	-0.6
348.67	9.21 ^d	0.1	370.36	39.54	-0.3	392.06	143.2	0.7
350.60	10.55 ^d	-0.1	372.36	45.07	0.3	393.98	157.7	-0.1
352.60	12.18 ^d	0.1	374.42	50.47	-1.1	396.07	177.0	0.1
354.43	13.72 ^d	-0.8	376.30	57.21	0.0	397.94	194.6	-0.4
356.55	15.91 ^d	-0.7	378.37	64.37	-0.7	399.88	217.0	0.2
358.53	18.29 ^d	-0.2	380.22	72.49	0.2			
Clopyralid								
<i>Crystalline phase</i> (Knudsen effusion method) ^c								
334.11	0.107	1.9	342.36	0.266	-1.3	350.40	0.640	-0.4
336.38	0.137	0.1	344.36	0.330	-1.6	352.13	0.783	1.4
338.46	0.172	-1.0	346.12	0.410	0.8	354.41	0.985	0.7
340.12	0.211	0.5	348.34	0.508	-1.6	356.38	1.203	0.4

^aThe standard uncertainty of the temperature is $u(T/\text{K}) = 0.01$ and the expanded uncertainties (0.95 confidence level, $k = 2$) of the vapor pressures are $U(p/\text{Pa}) = 0.01 + 0.0050 (p/\text{Pa})$ for static pressures and $u(p/\text{Pa}) = 0.02$ for the effusion pressures. ^b $\Delta p = p - p_{\text{calc}}$, where p_{calc} is calculated from the Clarke and Glew, Eq. (2), with parameters given in Table 3. ^cThe reported effusion pressures

are the mean of the values obtained using the small, medium, and large effusion orifices. ^dVapor pressures of the supercooled liquid.

Table 3 reports, for each compound, results of those properties and of $\Delta_{\text{cd}}^{\text{g}}S_{\text{m}}^{\circ}$ (calculated using Eq. (3)) and their related uncertainties for three different temperatures ($\theta = 298.15$ K, $\theta =$ mean temperature of the experiments and $\theta =$ temperature of the triple point, for fenclorim). The vapor pressure calculated from Eq.(2) for the three different temperatures are also reported in this table.

$$\Delta_{\text{cd}}^{\text{g}}S_{\text{m}}^{\circ}(\theta) = \frac{\Delta_{\text{cd}}^{\text{g}}H_{\text{m}}^{\circ}(\theta) - \Delta_{\text{cd}}^{\text{g}}G_{\text{m}}^{\circ}(\theta)}{\theta} \quad (3)$$

Table 3. Standard ($p^{\circ} = 10^5$ Pa) thermodynamic properties of sublimation and of vaporization of the compounds studied.

ΔT	θ	$\Delta_{\text{cd}}^{\text{g}}G_{\text{m}}^{\circ}(\theta)^{\text{a}}$	p^{b}	$\Delta_{\text{cd}}^{\text{g}}H_{\text{m}}^{\circ}(\theta)^{\text{a}}$	$\Delta_{\text{cd}}^{\text{g}}S_{\text{m}}^{\circ}(\theta)^{\text{c}}$	R^2	$-\Delta_{\text{cd}}^{\text{g}}C_{p,\text{m}}^{\circ}(\theta)^{\text{a}}$	$\sigma_{\text{r}}^{\text{d}}$
K	K	kJ·mol ⁻¹	Pa	kJ·mol ⁻¹	J·K ⁻¹ ·mol ⁻¹		J·K ⁻¹ ·mol ⁻¹	
Fenclorim								
<i>Crystalline phase</i> (Knudsen effusion method)								
311.1 to 333.2	298.15	38.19 ± 0.06	2.04×10 ⁻²	98.1 ± 0.8	200.9 ± 2.7	0.9998	23.5 ± 5.7 ^e	0.0110
	322.17 ^f	33.38 ± 0.02	3.87×10 ⁻¹	97.6 ± 0.8				
<i>Crystalline phase</i> (static method)								
326.0 to 365.5	298.15	38.22 ± 0.03	2.01×10 ⁻²	98.1 ± 0.2	200.8 ± 0.7	1.0000	23.5 ± 5.7 ^e	0.0064
	345.77 ^f	28.74 ± 0.02	4.55	97.0 ± 0.2				
<i>Crystalline phase</i> (Knudsen effusion + static methods)								
311.1 to 365.5	298.15	38.18 ± 0.06	2.05×10 ⁻²	98.0 ± 1.1	200.6 ± 3.7	1.0000	26.4 ± 13.3 ^g	0.0092
	338.32 ^f	30.20 ± 0.01	2.17	96.9 ± 0.2				
	367.39 ^h	24.50 ± 0.03	32.9	96.1 ± 0.2				
<i>Liquid phase</i> (static method) ⁱ								
338.8 to 399.9	298.15	33.85 ± 0.10	1.17×10 ⁻¹	76.1 ± 1.0	141.7 ± 3.4	1.0000	61.5 ± 7.1 ^g	0.0053
	369.34 ^f	24.25 ± 0.01	37.2	71.7 ± 0.1				
	367.39 ^h	24.50 ± 0.01	32.9	71.9 ± 0.1				
Clopyralid								
<i>Crystalline phase</i> (Knudsen effusion method)								
334.1 to 356.4	298.15	45.82 ± 0.14	9.39×10 ⁻⁴	109.1 ± 1.0	212.2 ± 3.4	0.9998	21.5 ± 4.6 ^e	0.0124
	345.24 ^f	35.91 ± 0.02	3.69×10 ⁻¹	108.1 ± 1.0				

^aUncertainties are expressed as the expanded uncertainty (0.95 level of confidence, $k = 2$). ^bCalculated from Eq. (2) for three different temperatures ($\theta = 298.15$ K, $\theta =$ mean temperature of the experiments and $\theta =$ temperature of the triple point). ^cCalculated using Eq. (3); uncertainties calculated through the RSS method. ^d σ_{r} is the relative standard deviation

of the fit, defined as $\sigma_{\text{r}} = \left[\sum_{i=1}^n (\ln p - \ln p_{\text{calc}})_i^2 / (n-m) \right]^{1/2}$. ^eCalculated as $\Delta_{\text{cr}}^{\text{g}}C_{p,\text{m}}^{\circ}(\theta) = C_{p,\text{m}}^{\circ}(\text{g}) - C_{p,\text{m}}^{\circ}(\text{cr})_{\text{exp}}$. ^fMean temperature.

^gAdjustable parameter derived from the fittings of Eq. (2) to the (p, T) data. Uncertainties are standard deviations of the least-squares regressions. ^hTemperature of the triple point. ⁱIncluding supercooled liquid.

If experimental sublimation or vaporization vapor pressures are determined over a wide temperature range (~50 K), the fit of Eq. (2) to the experimental (p, T) data frequently yields accurate values of $\Delta_{\text{cd}}^{\text{g}}C_{p,\text{m}}^{\circ}(\theta)$. In this work, the values of $\Delta_{\text{cr}}^{\text{g}}C_{p,\text{m}}^{\circ}(\theta)$ and $\Delta_{\text{l}}^{\text{g}}C_{p,\text{m}}^{\circ}(\theta)$ were

derived directly from the linear regression of Eq. (2) to the crystalline and liquid experimental results of fenclorim, respectively. The sublimation results, determined through Knudsen effusion and static methods, were fit together using Eq. (2), enabling to derive a reliable value of $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(\theta)$. Since the sublimation temperature range considered in this work for clopyralid (using Knudsen effusion method) was not large enough, the value of $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(\theta)$ could not be derived for this compound by such procedure. As an alternative, it was calculated as $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(\theta) = C_{p,m}^{\circ}(\text{g}) - C_{p,m}^{\circ}(\text{cr})$, where $C_{p,m}^{\circ}(\text{g})$ and $C_{p,m}^{\circ}(\text{cr})$ are, respectively, the gas and crystalline isobaric molar heat capacities. The values of $C_{p,m}^{\circ}(\text{cr})$, determined through heat capacity drop calorimetry, and the derived standard isobaric specific heat capacities, c_p° , are reported in Table 4, together with the specific densities and the volumetric heat capacities, $C_{p,m}^{\circ}/V_m$, determined at $T = 298.15$ K for each compound. The experimental results of $C_{p,m}^{\circ}(\text{cr})$ are in agreement with the ones estimated using the group contribution values proposed by Acree Jr. and Chickos [37], also listed in this table.

Table 4. Experimental and estimated crystalline standard molar heat capacities $C_{p,m}^{\circ}$, specific heat capacities, c_p° , densities and volumetric heat capacities, $C_{p,m}^{\circ}/V_m$, at $T = 298.15$ K of fenclorim and clopyralid.

Compound	$C_{p,m}^{\circ}$ ^a JK ⁻¹ ·mol ⁻¹	Molar Mass g·mol ⁻¹	c_p° J·K ⁻¹ ·g ⁻¹	Density g·cm ⁻³	$C_{p,m}^{\circ}/V_m$ ^b J·K ⁻¹ ·cm ⁻³	$C_{p,m}^{\circ}(\text{Estimated})$ ^c J·K ⁻¹ ·mol ⁻¹
Fenclorim	214.3 ± 0.6	225.074	0.952 ± 0.003	1.541 ^d	1.467 ± 0.005	221.8 ± 17.0
Clopyralid	175.3 ± 0.6	191.999	0.913 ± 0.003	1.64 ± 0.05 ^e	1.497 ± 0.050	181.4 ± 17.0

^aThe reported experimental uncertainties are twice the standard deviation of the mean and includes the calibration uncertainty. ^bCalculated considering the specific heat capacities, c_p° , and the experimental density values. ^cEstimated using the group contribution values proposed by Acree Jr. and Chickos [37]. ^dRef. [38]. ^eAverage of three measurements of the volume and mass of three pellets.

The values of $C_{p,m}^{\circ}(\text{g})$, determined at the temperature 298.15 K for the two compounds, were derived from statistical thermodynamics, calculated by means of the Gaussian 09 software package [33], using the vibrational frequencies from G3(MP2)//B3LYP calculations (scaled by a factor of (0.960 ± 0.022) [39]). The results of $C_{p,m}^{\circ}(\text{g}, 298.15 \text{ K})/\text{J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ for fenclorim and clopyralid were, respectively, (190.8 ± 5.7) and (153.8 ± 4.6) . The standard uncertainties in $C_{p,m}^{\circ}(\text{g})$ were estimated as $u[C_{p,m}^{\circ}(\text{g})] = 0.03 \cdot C_{p,m}^{\circ}(\text{g})$ [40]. As the values of $C_{p,m}^{\circ}(\text{cr})$ were determined at 298.15 K it was assumed that $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(\theta)$ is approximately constant inside the assigned uncertainties. For fenclorim, the value $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(298.15 \text{ K}) = -(26.4 \pm 13.3) \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$, derived from the fitting of Eq. (2) to the sublimation (p, T) data, is in agreement with the one calculated through the theoretical value of $C_{p,m}^{\circ}(\text{g})$ and the experimental result of $C_{p,m}^{\circ}(\text{cr})$, $\Delta_{\text{cr}}^{\text{g}} C_{p,m}^{\circ}(298.15 \text{ K}) = -(23.5 \pm 5.7) \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$.

The average of the onset of fusion temperatures, T_{fus} , molar enthalpies, $\Delta_{\text{cr}}^{\text{l}} H_m^{\circ}(T_{\text{fus}})$, and entropies, $\Delta_{\text{cr}}^{\text{l}} S_m^{\circ}(T_{\text{fus}})$, of the two compounds studied are reported in Table 5, together with the fusion properties of fenclorim derived indirectly through static vapor pressure measurements.

Table 5. Fusion properties: temperature, molar enthalpy and entropy of the compounds studied.

T_{tp}/K	T_{fus}/K^a	$\Delta_{cr}^1 H_m^o(T)^b / kJ \cdot mol^{-1}$	$\Delta_{cr}^1 S_m^o(T)^{b,c} / J \cdot K^{-1} \cdot mol^{-1}$	Method/Ref.
Fenclorim				
	368.61 ± 0.35	23.08 ± 0.69^a	62.6 ± 1.9	DSC / this work
367.39		24.2 ± 0.2		VP / this work
Clopyralid				
	422.63 ± 0.38	27.59 ± 0.67^a	65.3 ± 1.6	DSC / this work

^aStandard uncertainty calculated through the RSS method combining the expanded uncertainties of the four experimental runs (0.95 level of confidence, $k = 3.18$) with the standard uncertainties of the DSC calibration. ^b T represents the temperature of fusion or the temperature of the triple point (T_{tp}).

^cUncertainties calculated through the RSS method.

Figure 2 presents the phase diagram of fenclorim in the neighborhood of the triple point and figure 3 shows the plot of vapor pressures against the reciprocal temperatures of clopyralid. To the best of our knowledge no (p, T) study was reported before for the two compounds studied in this work.

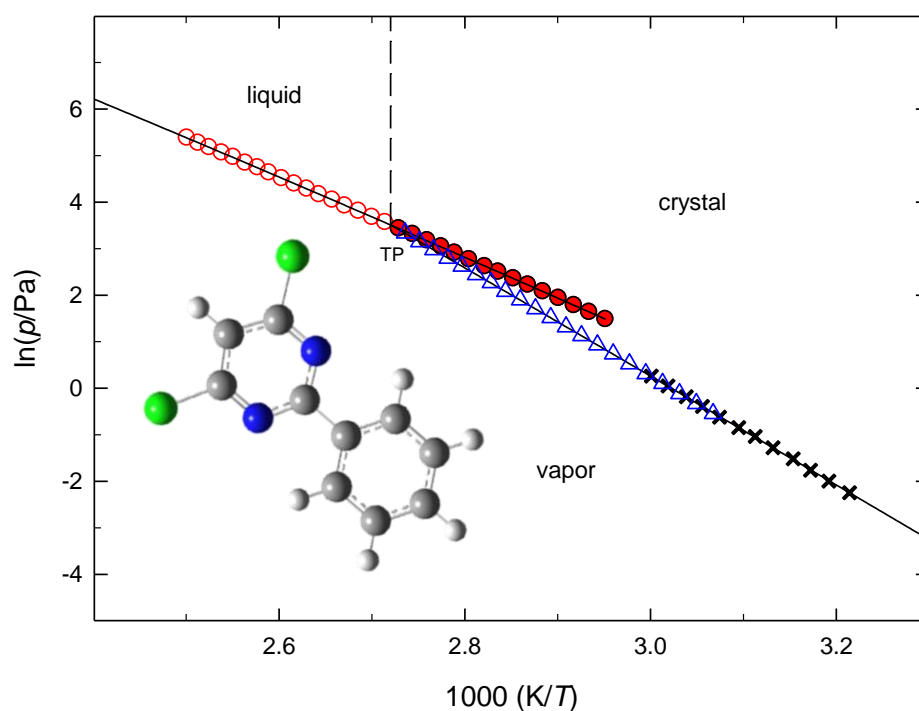


Figure 2. Phase diagram of fenclorim. \circ , vaporization; \bullet , vaporization (supercooled liquid); Δ , sublimation (static method); \times , sublimation (mean values of Knudsen effusion vapor pressures). Triple point data determined in this work: $T = 367.4$ K; $p = 32.9$ Pa.

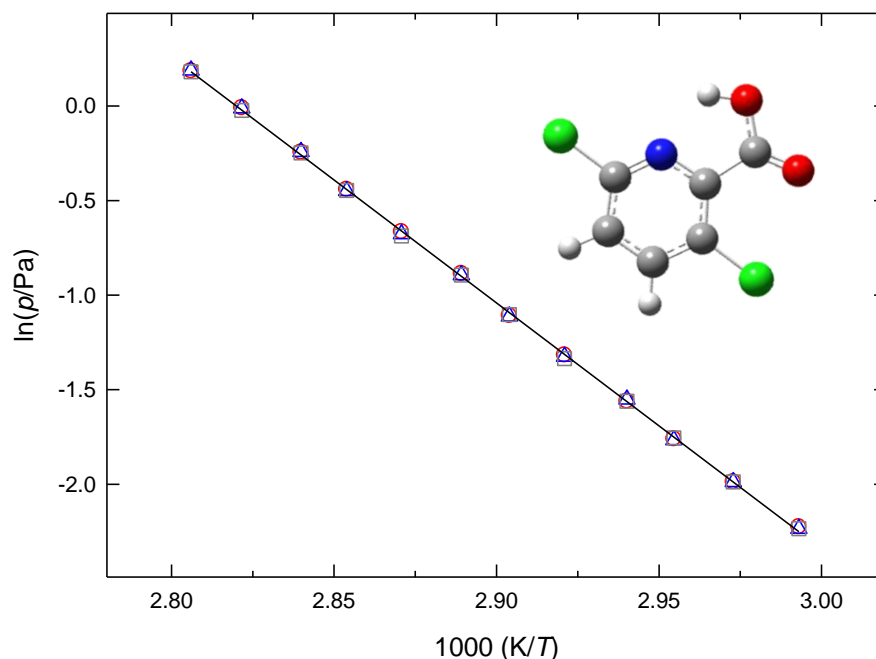


Figure 3. Plot of $\ln p$ against $1/T$ for clopyralid. \circ , small effusion orifices; Δ , medium effusion orifices and \square , large effusion orifices.

3.2. Thermodynamic stability of fenclorim and clopyralid

The standard molar Gibbs energy of formation, $\Delta_f G_m^\circ$, may be used to evaluate the thermodynamic stability of a compound at standard conditions. The values of this property are quite important to obtain equilibrium constants of reactions where they act as reagents or products. Therefore, the values of $\Delta_f G_m^\circ$ of fenclorim and clopyralid were calculated in this work, at $T = 298.15$ K, in crystalline and gaseous phases. The results of the gas phase were calculated using Eq. (4) and those of the crystalline phase were obtained by subtracting from $\Delta_f G_m^\circ(\text{g}, 298.15 \text{ K})$ the results of $\Delta_{\text{cr}}^\circ G_m^\circ(298.15 \text{ K})$, derived through static and/or Knudsen effusion measurements.

$$\Delta_f G_m^\circ(298.15 \text{ K}, \text{g}) = \Delta_f H_m^\circ(298.15 \text{ K}, \text{g}) - 298.15 \cdot \Delta_f S_m^\circ(298.15 \text{ K}, \text{g}) \quad (4)$$

3.2.1. Thermodynamic properties of formation in gaseous phase

Theoretical gas phase standard molar enthalpies of formation of fenclorim and clopyralid, $\Delta_f H_m^\circ(\text{g})$, were calculated from a set of hypothetical working reactions that involve maximum bonding pattern similarity between the products and the reactants. They also satisfy the requirement that accurate experimental data for all the auxiliary molecules used are reported in the literature. These reactions consider the enthalpies of reaction, $\Delta_r H^\circ(298.15 \text{ K})$, obtained through computational absolute standard enthalpies, $H_{298.15 \text{ K}}^\circ$, and the experimental gas phase enthalpies of formation of all included molecules, available in the literature. The G3(MP2)/B3LYP absolute enthalpies and the gas phase enthalpies of formation, at $T = 298.15$ K, for all the molecules considered in this work are provided as Supporting Material, in Table S6. The gas phase hypothetical reactions, the calculated values for the enthalpy variations as well as the resulting estimates of the enthalpies of formation in the gas phase, at $T = 298.15$ K, are reported in the Tables S7 and S8, for fenclorim and clopyralid, respectively.

The gas phase standard molar entropies of formation, $\Delta_f S_m^\circ(\text{g})$, were calculated from the gas phase standard absolute entropies, $S_m^\circ(\text{g}, 298.15 \text{ K})$, obtained by G3(MP2)/B3LYP composite method, and from the following reference standard entropy values [41]: $S_m^\circ[\text{C}(\text{graphite})] = 5.740 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$, $S_m^\circ(\text{H}_2, \text{g}) = 130.680 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$, $S_m^\circ(\text{Cl}_2, \text{g}) = 223.079 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$, $S_m^\circ(\text{N}_2, \text{g}) = 191.609 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ and $S_m^\circ(\text{O}_2, \text{g}) = 205.147 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ (for clopyralid).

Considering the estimated results of $\Delta_f H_m^\circ(\text{g}, 298.15 \text{ K})$ and $\Delta_f S_m^\circ(\text{g}, 298.15 \text{ K})$, the values of $\Delta_f G_m^\circ(\text{g}, 298.15 \text{ K})$ of fenclorim and clopyralid reported in Table 6 were determined using Eq. (4).

3.2.2. Thermodynamic properties of formation in crystalline phase

Considering the values of $\Delta_f G_m^\circ(\text{cr}, 298.15 \text{ K})$ and of $\Delta_f S_m^\circ(\text{cr}, 298.15 \text{ K})$, the standard molar enthalpy of formation in crystalline phase, $\Delta_f H_m^\circ(\text{cr}, 298.15 \text{ K})$, was also estimated for both compounds. The gas phase standard molar entropies of formation, $\Delta_f S_m^\circ(\text{cr})$ were determined from the values of $S_m^\circ(\text{cr}, 298.15 \text{ K})$ — calculated considering the $\Delta_{\text{cr}}^{\text{g}} S_m^\circ(298.15 \text{ K})$ results — and from the reference standard entropy values presented above.

Table 6 reports the gas phase standard entropies of the two nitrogen heterocycles and the values of their standard molar enthalpies, entropies and Gibbs energies of formation, in crystalline and gaseous phases, as well as the results of the sublimation properties derived through the vapor pressure measurements.

The results reported in Table 6 and shown in Figure 5 indicate that clopyralid is thermodynamically more stable — lower values of $\Delta_f G_m^\circ$ — than fenclorim in both crystalline and gaseous phases. It is easily noticeable that the enthalpic values of clopyralid have a major contribution to the smaller values of $\Delta_f G_m^\circ$ of this compound in both phases.

The dominant enthalpic influence in the crystalline phase of clopyralid seems to be a consequence of the eventual O-H \cdots O and N-H \cdots N hydrogen bonds that are probably formed between clopyralid molecules. Although the crystal structure of this compound was not reported in the literature, we guess, as a credible hypothesis, that those intermolecular hydrogen bonds are formed in the crystalline phase of clopyralid similarly to the intermolecular bonds noticeable in the crystalline pattern of 2-pyridinecarboxylic acid [42], schematically represented in figure 4. In the crystal structure of this compound, a zigzag chain is formed by the hydrogen bonds, N—H \cdots N and O—H \cdots O, where the protons are displaced around the twofold axis, or the center of symmetry and their site occupation factors were assumed to be 50% [42]. The authors of the article [42] concluded that the intermolecular N—H \cdots N hydrogen bond is asymmetric and hydrogen atom is trapped at one of the two nitrogen atoms. Those bonds are much stronger than the CH \cdots N hydrogen bonds and the C-Cl \cdots π and π - π interactions present in the crystalline structure of fenclorim, reported in the literature [38].

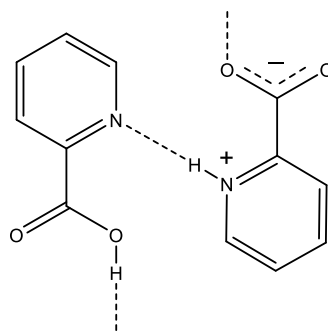


Figure 4. Schematic representation of the crystalline packing arrangement of 2-pyridinecarboxylic acid (adapted from ref 41).

The stability of clopyralid in the gaseous phase hints the possible occurrence of an intramolecular hydrogen bond O-H...N in the isolated molecules of this compound. This bond should occur due to the proximity and orientation of the hydrogen of the hydroxyl group toward the nitrogen of the pyridine ring, with a distance of 1.95×10^{-10} m. This kind of hydrogen bond can be estimated as the difference in energy between the most stable conformation and the corresponding conformation optimized in a trans orientation relative to the proton acceptor fragment [43,44]. In this way, we estimated the value 11.5 kJmol^{-1} for the O-H...N bond strength in this compound. This result is in close agreement with the ones reported in the literature for other substituted pyridine carboxylic acids: 14.0 kJmol^{-1} , 11.8 kJmol^{-1} , 12.4 kJmol^{-1} , and 13.9 kJmol^{-1} , respectively, for 2-pyridinecarboxylic acid, pyridine-2,4-dicarboxylic acid, pyridine-2,5-dicarboxylic acid, and pyridine-2,6-dicarboxylic acid [43].

Table 6. Standard ($p^0 = 10^5 \text{ Pa}$) molar absolute entropies and standard molar enthalpies, entropies and Gibbs energies of formation and sublimation, at $T = 298.15 \text{ K}$.

		Fencloirim	Clopyralid
Gas phase	$\Delta_f H_m^{o a} / \text{kJ.mol}^{-1}$	201.4 ± 2.9	-280.3 ± 1.7
	$S_m^{o b} / \text{J.K}^{-1}.\text{mol}^{-1}$	447.6	407.7
	$\Delta_f S_m^o / \text{J.K}^{-1}.\text{mol}^{-1}$	-416.5	-346.8
	$T\Delta_f S_m^o / \text{kJ.mol}^{-1}$	-124.2	-103.4
	$\Delta_f G_m^{o c} / \text{kJ.mol}^{-1}$	325.6 ± 2.9	-176.9 ± 1.7
<i>Thermodynamic Properties of Sublimation^d</i>	$\Delta_{cr}^g H_m^o / \text{kJ.mol}^{-1}$	98.0 ± 1.1	109.1 ± 1.0
	$\Delta_{cr}^g S_m^o / \text{J.K}^{-1}.\text{mol}^{-1}$	200.6 ± 3.7	212.2 ± 3.4
	$\Delta_{cr}^g G_m^o / \text{kJ.mol}^{-1}$	38.2 ± 0.1	45.8 ± 0.1
Crystalline Phase	$\Delta_f H_m^o / \text{kJ.mol}^{-1}$	103.4 ± 3.1	-389.4 ± 2.0
	$S_m^{o f} / \text{J.K}^{-1}.\text{mol}^{-1}$	247.0 ± 3.7	195.5 ± 3.4
	$\Delta_f S_m^o / \text{J.K}^{-1}.\text{mol}^{-1}$	-617.1 ± 3.7	-559.0 ± 3.4
	$T\Delta_f S_m^o / \text{kJ.mol}^{-1}$	-184.0 ± 1.1	-166.7 ± 1.0
	$\Delta_f G_m^{o e} / \text{kJ.mol}^{-1}$	287.4 ± 2.9	-222.7 ± 1.7

^aThe uncertainty assigned correspond to the expanded uncertainty determined from the estimated standard deviation of the mean (0.95 level of confidence) for the working reactions reported in SI (Tables S7 and S8). ^bObtained from G3(MP2)//B3LYP method for a frequency factor scale of 1.0029 [45]. ^cCalculated using Eq.(4). ^dDerived through vapor pressure measurements. ^eCalculated from $\Delta_f G_m^o(298.15 \text{ K, cr}) = \Delta_f G_m^o(298.15 \text{ K, g}) - \Delta_{cr}^g G_m^o(298.15 \text{ K})$. ^fCalculated from $S_m^o(298.15 \text{ K, cr}) = S_m^o(298.15 \text{ K, g}) - \Delta_{cr}^g S_m^o(298.15 \text{ K})$.

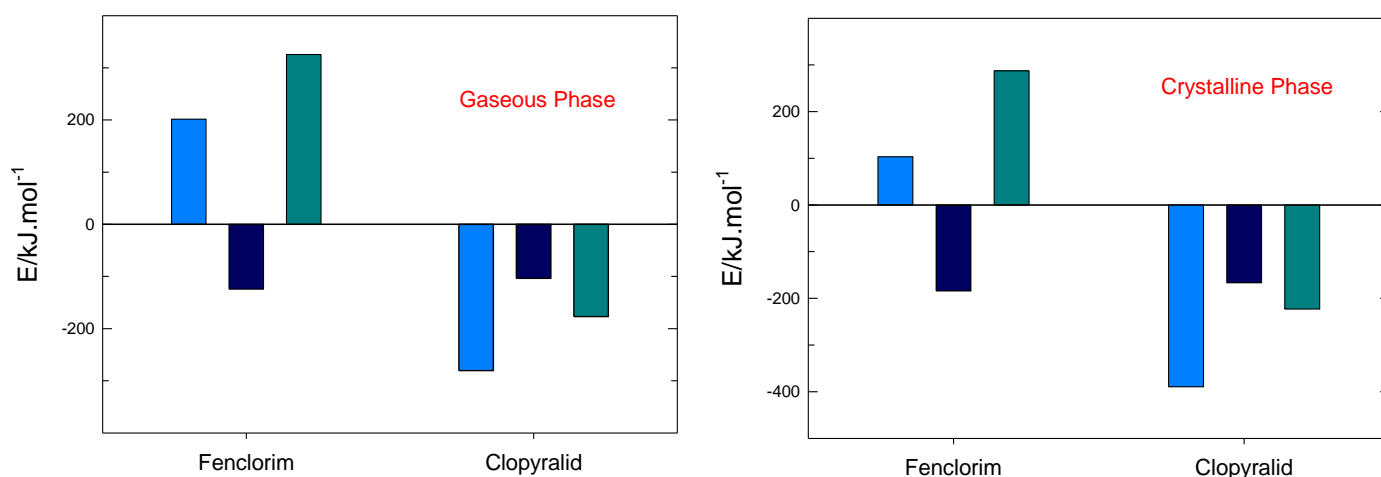


Figure 5. Relation between $\Delta_f H_m^\circ$, $T\Delta_f S_m^\circ$ and $\Delta_f G_m^\circ$ for fenclorim and clopyralid in both gaseous and crystalline phases.

• $E = \Delta_f H_m^\circ$; • $E = T\Delta_f S_m^\circ$; • $E = \Delta_f G_m^\circ$.

4. Conclusions

The relevant conclusions of the present work are the following:

- The temperatures and molar enthalpies of fusion of the herbicide safener fenclorim and of the herbicide clopyralid were determined using DSC and their crystalline isobaric molar heat capacities were measured, at 298.15 K, using heat capacity drop calorimetry. From these results the standard isobaric specific heat capacities and the volumetric heat capacities were obtained.
- The enthalpies, entropies and Gibbs energies of sublimation of both compounds and of vaporization of fenclorim were derived through vapor pressure measurements and the phase diagram representation of the (p, T) results, including triple point coordinates of this compound, is reported.
- Computational calculations at the G3(MP2)//B3LYP level were carried out, and the estimated gas-phase enthalpies of formation and absolute entropies of the two compounds were used to determine their Gibbs energies of formation in this phase. These results together with the sublimation ones enabled the calculation of the Gibbs energies of formation in crystalline phase.
- Clopyralid is thermodynamically more stable than fenclorim, in crystalline and gaseous phases, mainly due to the enthalpic contribution.
- The standard molar enthalpy of formation in crystalline phase, $\Delta_f H_m^\circ(\text{cr}, 2981.5 \text{ K})$ was estimated for both compounds.
- It is probable that intramolecular OH...N hydrogen bonds occur in the gaseous phase of clopyralid and that intermolecular O-H...O and N-H...N hydrogen bonds exist between molecules of this crystalline compound, just like in the crystal structure of 2-pyridinecarboxylic acid.

Supplementary Materials: The following are available online at www.mdpi.com/xxx/s1, Table S1. Specific densities of fenclorim and clopyralid. Table S2. DSC results: temperatures, molar enthalpies and entropies of fusion of fenclorim and clopyralid. Table S3. Standard molar heat capacity results, at $T = 298.15$ K, for fenclorim and clopyralid. Table S4. Effusion vapor pressure results for crystalline fenclorim. Table S5. Effusion vapor pressure results for crystalline clopyralid. Table S6. Calculated absolute enthalpies, H , (in Hartree, Eh) of all considered molecules estimated using G3(MP2)/B3LYP. Literature values of $\Delta_f H_m^0$ (g). Table S7. Working reactions and computed enthalpies of reaction, $\Delta_r H_m^0$, and formation, $\Delta_f H_m^0$, of fenclorim in the gaseous state, at $T = 298.15$ K. Table S8. Working reactions and computed enthalpies of reaction, $\Delta_r H_m^0$, and formation, $\Delta_f H_m^0$, of clopyralid in the gaseous state, at $T = 298.15$ K.

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Conflicts of Interest: The authors declare no conflict of interest

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GRAPHICAL ABSTRACT

