## Supplementary material to

## Mechanisms of phase transformation and creating mechanical strength in a sustainable calcium carbonate cement

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## S.1 Growth and dissolution rates

### S.1.1 Growth or dissolution control

The crystal growth/dissolution rate, *r,* (dissolution if *r* is negative) can be expressed as:

where *k* is the growth rate constant (kg s-1 m-2), *A* is the reactive surface area of the growing phase, d*m*/d*t* is the change in solid mass of the phase per time, *Ea* is the apparent activation energy of the overall reaction, *R* is the gas constant, *T* is the absolute temperature and *f(ΔG)* introduces the dependence of the overall growth rate on the supersaturation state of the system expressed as a function, *f*, of the Gibbs free-energy change for the growth reaction, *ΔG*. The velocity of the growing/dissolving surface is

where *Vm* is the molar volume (vaterite: 3.8.10-5 m3.mol-1, calcite: 3.7.10-5 m3.mol-1).

The *ΔG* function may have a variety of forms. However, since the overall reaction here presented is an elementary reaction, *f(ΔG)* can be derived from transition state theory (TST) [1,2] in the following way:

The equation of the phase evolution rate when transport is much faster than growth and dissolution is

If the vaterite grains are of similar size there will be approximately a constant number of vaterite grains the whole time and therefore . If all calcite nuclei are formed initially and the number of calcite grains remain constant, then .

If the rate of calcite growth is much slower than the rate of vaterite dissolution, the aqueous CaCO3 concentration will equal the vaterite solubility and the driving force for the growth will be kept constant at the difference of the vaterite and calcite phase. Inversely, if vaterite dissolution is much slower than calcite growth, the aqueous CaCO3 concentration will equal the calcite solubility. Thus the same will be the constant driving force of vaterite dissolution. Since everything is constant but the mass and the area of the growing phases the rate equation simplifies to

for calcite growth control and

for vaterite dissolution control. The final time is inversely proportional to the rate: , Since the vaterite dissolution control model can be written:

Previous phase transformation studies in batch reactors with well stirred liquid solutions reported by Ogino et al. [3] and Rodriguez-Blanco et al. [4] found that and consequently concluded that the phase transformation ACC - vaterite - calcite was controlled by the rate of calcite growth.

### S.1.2 Nucleation control

Even though the calcite growth rate is 2-10 times larger than the vaterite dissolution rate (see below), a restricted number of calcite nuclei may render the phase transformation controlled by calcite growth. If *N0* new nuclei are formed initially and no new nuclei are formed after that, the mass of calcite will grow as , where (assuming constant supersaturation as above). This means that even though the rate constant of calcite growth is larger than that of vaterite dissolution the effective growth rate is proportional to the number of nuclei and can be much smaller than the vaterite dissolution rate. The observation by Ogino et al. [3] and Rodriguez-Blanco et al. [4] found that is thus probably due to the limited number of calcite nuclei in the well stirred liquid solutions.

### S.1.3 Diffusion control

If, on the other hand, diffusion is much slower than the growth and dissolution rates the transformation rate will be , where is the diffusion distance, *D* the diffusion coefficient, *V* a standard volume, and and are the solubilities of vaterite and calcite. Different models of how the diffusion distance evolves will yield different time dependence. We will assume that all calcite grains nucleate close to vaterite grains so that initially the diffusion distance is small. As the calcite grains grow the diffusion distance is proportional to the radius of the calcite grains, thus . This yields the rate equation:

This simple model is captures an essential part of the phase transformation of the cement. A numerical simulation is probably necessary to describe the full complexity.

### S.1.1 Driving forces for dissolution

#### S.1.1.1 Curvature:

where *r* is the radius of curvature of the surface J m-2 is the surface energy [5].

#### S.1.1.2 Phase change

ACC - Calcite: 15 kJ mol-1 [6]

Vaterite - Calcite: 6.2 kJ mol-1 [7]

### S.1.2 Rate constants

*Dissolution rate of vaterite* at 20°C:

*rd,v*=-, and mol.m-2s-1 from Cubillas et al [8] . Thus for a system controlled by the energy difference of vaterite and calcite, 6.2 kJ mol-1,

*rd,v=-*2.2.10-5 mol.m-2s-1.

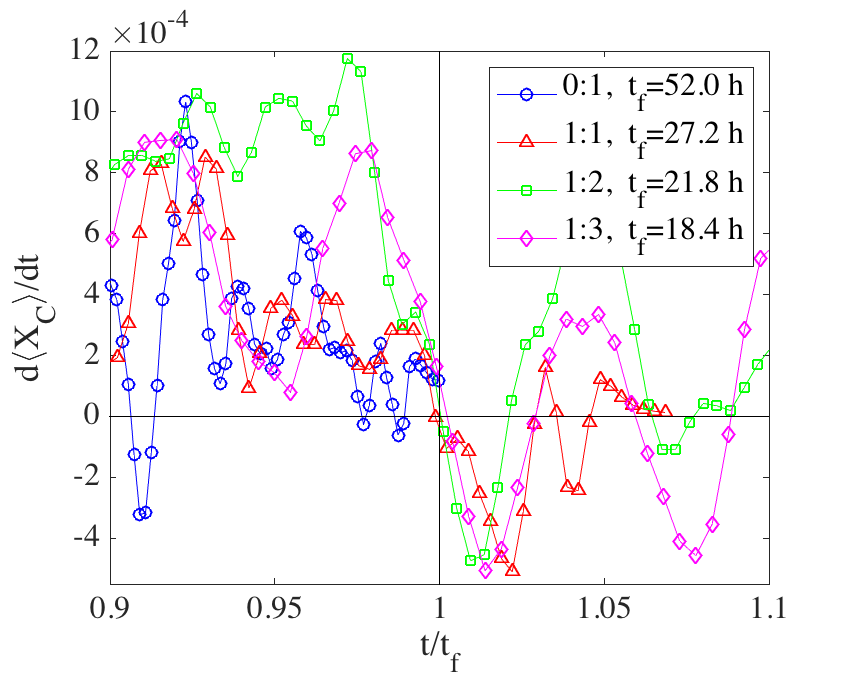
*Growth rate of calcite* at 20°C:

*rg,c=*, where nm h-1 = 2.37.10-5 mol.m-2s-1 from single crystal experiments [9] and 4.6.10-6 mol.m-2s-1 from batch reactor experiments [10]. Thus for a system controlled by the energy difference of vaterite and calcite, 6.2 kJ mol-1, *rd,v* is in the range from5.4.10-5 to 2.8.10-4 mol.m-2s-1.

*Mean diffusion coefficient of CaCO3 ions*: m2 s-1.

## S.2 Analysis of phase transformation data

The first step in our analysis is to estimate the final time, *tf*. From Figure S2 one observes that there is some vaterite left in the cement after the phase transformation has slowed down to a rate that is essentially zero at the resolution of our experiments. As a definition of *tf* we require that the local slope of the mass fraction curves is smaller than the fluctuation in the slope at *tf*. This is done using a local regression using weighted linear least squares and a 1st degree polynomial model with a window of 15 data points. Figure S1 shows the slopes of the calcite mass fractions around *t/tf*=1.



***Figure S1: Local slopes of calcite mass fractions at times approaching tf.***

Assuming dissolution, growth or diffusion to be the rate limiting step of the phase transformation yields models with different power law dependencies. There are two ways one may rescale the data to compare to the models. Either by rescaling the mass and by the final time, *tf* :

or by shifting the starting time, *ts*, and rescaling by the final time, *tf* :

In Figure S2 we have used both ways of rescaling and compared to the models that yield different exponents .

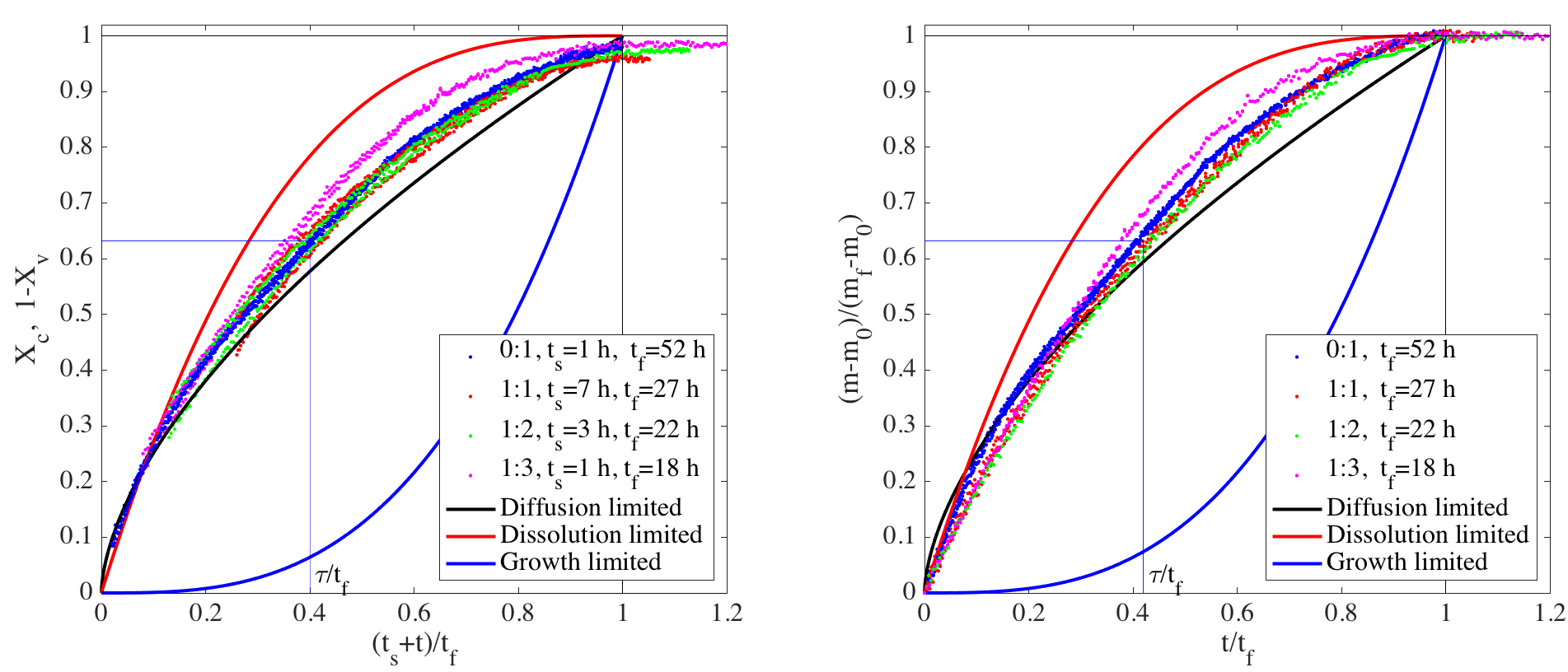


Figure S2. Normalized mass evolution with normalized time for four different CaCO3 cement compositions (0:1, 1:1, 1:2 and 1:3 wt.% ACC:V). Using a standard definition of the relaxation time : ln[(m0-m)/(mf-m0)] = t /, the relaxation times are  ~ 0.4 tf, thus  = 21, 11, 9 and 7 hours, respectively. The solid lines demonstrate the three different phase transformation models.

The time constants, *tf,* are 3-10 times longer than those found for stirred liquid experiments with ACC:V=0:1, *tf,liquid* =5 h at 20°C [4]. One observes that the two ways of rescaling the data is roughly equivalent within the accuracy of the data. Using a standard definition of the relaxation time : ln[(m0-m)/(mf-m0)] = ln(X) = t /, the relaxation times , are τACC:V 1:1 = 11 h, τACC:V 1:2 = 9 h, τACC:V 1:3 = 7 h and τACC:V 0:1 = 21 h.

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