

Article

Inorganic Anion Regulates the Phase Transition in Two Organic Cation Salts Containing [(4-Nitroanilinium)(18-crown-6)]⁺ Supramolecules

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Abstract: Two novel inorganic–organic hybrid supramolecular assemblies, namely, (4-HNA)(18-crown-6)(HSO₄) (**1**) and (4-HNA)₂(18-crown-6)₂(PF₆)₂(CH₃OH) (**2**) (4-HNA = 4-nitroanilinium), were synthesized and characterized by infrared spectroscopy, single X-ray diffraction, differential scanning calorimetry (DSC), and temperature-dependent dielectric measurements. The two compounds underwent reversible phase transitions at about 255 K and 265 K, respectively. These phase transitions were revealed and confirmed by the thermal anomalies in DSC measurements and abrupt dielectric anomalies during heating. The phase transition may have originated from the [(4-HNA)(18-crown-6)]⁺ supramolecular cation. The inorganic anions tuned the crystal packings and thus influenced the phase-transition points and types. The variable-temperature X-ray diffraction data for crystal **1** revealed the occurrence of a phase transition in the high-temperature phase with the space group of *P*2₁/c and in the low-temperature phase with the space group of *P*2₁/n. Crystal **2** exhibited the same space group *P*2₁/c at different temperatures. The results indicated that crystals **1** and **2** both underwent an iso-structural phase transition.

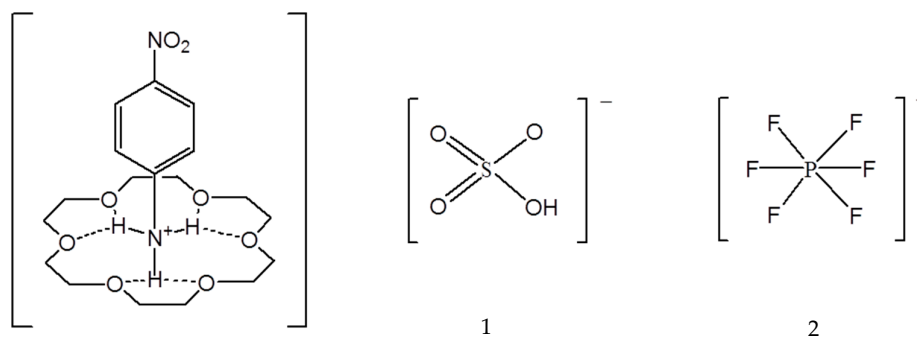
Keywords: Phase transition; Crystal structure; Dielectric anomaly; Supramolecular cation; Inorganic anion

1. Introduction

From both theoretical and application viewpoints, phase-transition compounds are popular and valuable in increasing the general understanding on structure–property relations and exploring functional materials with novel physical properties [1–5]. Phase-transition crystalline materials are usually accompanied by an abrupt change in some physical properties around the transition temperature. These materials show great applications in molecular sensors, switches, and data storage devices [6–7]. In particular, energy harvesting occurs during phase changes; thus, phase-transition materials may serve as energy-saving materials. Designing special inorganic–organic hybrid compounds with molecular dielectrics is an effective method to synthesize ideal phase-transition materials [8–9].

Phase-transition inorganic–organic hybrid compounds display novel crystal structures and interesting physical properties, including ferroelectric, dielectric, optical, and piezoelectric properties [10–12]. Among them, 15-crown-5 or 18-crown-6 are good candidates owing to their variable conformation, such as [(RNH₃)(18-crown-6)][A], where R is an alkyl or aryl group and A is an anion. The driving force of these phase-transition compounds can be ascribed to the motional

changes in the $R-NH_3^+$ guest cation (R = aryl group) or/and anionic units. The use of the R group as a molecular rotor or pendulum unit produces desirable properties [13–15]. As a result, the motion of organic ammonium cations changes their dynamic state. This alteration further leads to dielectric changes and ferroelectricity through the phase transitions between the disordered high-temperature phase (HTP) and ordered low-temperature phase (LTP) [16–17]. On the contrary, the asymmetric unit of the supramolecular adduct also contains counter anions. These anions are mainly the tetrahedral ions, such as BF_4^- , ClO_4^- , and IO_4^- . These anions easily change in position with varied temperatures and weak interactions because of their higher symmetry and relatively smaller volume [18–19]. In fact, inorganic anions (such as HSO_4^-) are rarely explored. Given these findings, we report herein the syntheses of (4-HNA)(18-crown-6)(HSO_4) (**1**) and (4-HNA)₂(18-crown-6)₂(PF_6)₂(CH_3OH) (**2**) to determine other suitable geometric anions that can regulate a potential phase transition. The structures, phase transitions, and dielectric properties of the two anions are revealed in Scheme 1.



Scheme 1. Structural formulae of compounds **1** and **2**

2. Results and discussion

2.1 Crystal structure of **1**

The crystal structure of **1** was characterized at different temperatures by X-ray diffraction to confirm whether the phase transition was associated with structural changes. Compound **1** crystallized in the monoclinic space group $P2_1/c$ with $a = 10.450(6)$ Å, $b = 18.453(11)$ Å, $c = 13.466(8)$ Å, $\beta = 112.672(7)^\circ$, and $V = 2396(2)$ Å³ in the HTP at 296 K. At the LTP (100 K), crystal **1** crystallized in the monoclinic space group $P2_1/n$ with small changes in cell parameters, i.e., $a = 10.3715(18)$ Å, $b = 18.235(3)$ Å, $c = 13.384(3)$ Å, $\beta = 112.736(2)^\circ$, and $V = 2334.4(8)$ Å³ (Table 1).

Table 1. Crystallographic data and structural refinement details for **1** and **2**

	1-LTP	1-HTP	2-LTP	2-HTP
Chemical formula	C ₁₈ H ₃₂ N ₂ O ₁₂ S	C ₁₈ H ₃₂ N ₂ O ₁₂ S	C ₃₇ H ₆₈ F ₁₂ N ₄ O ₁₇ P ₂	C ₃₇ H ₆₈ F ₁₂ N ₄ O ₁₇ P ₂
Formula weight	500.52	500.52	1130.88	1130.88
Temperature / K	100	296	100	296
Crystal size (mm ³)	0.21×0.20×0.19	0.21×0.20×0.19	0.316×0.223×0.139	0.316×0.223×0.139
Crystal system	monoclinic	monoclinic	monoclinic	monoclinic
Space group	<i>P</i> 2 ₁ /n	<i>P</i> 2 ₁ /c	<i>P</i> 2 ₁ /c	<i>P</i> 2 ₁ /c
a (Å)	10.3715(18)	10.450(6)	10.9491(9)	10.9429(9)
b (Å)	18.235(3)	18.453(11)	23.701(2)	23.687(2)
c (Å)	13.384(3)	13.466(8)	21.9038(16)	21.8926(16)
α (°)	90.00	90.00	90.00	90.00
β (°)	112.736(2)	112.672(7)	113.961(4)	113.962(4)
γ (°)	90.00	90.00	90.00	90.00
V (Å ³)	2334.4(8)	2396(2)	5194.3(7)	5185.6(7)
Z	4	4	2	2
D _{calc} (g·cm ^{−3})	1.424	1.388	1.351	1.167
F(000)	1064	1064	2222	1946
m (mm ^{−1})	0.204	0.198	0.168	0.112
Measured 2θ range (°)	0.999–25.01	0.999–25.01	0.992–24.50	0.981–25.01
R _{int}	0.0372	0.0511	0.0992	0.0998
R (I > 2 σ(I)) [a]	0.0442	0.0738	0.1820	0.1871
WR (all data) [b]	0.1066	0.1608	0.3184	0.2917
GOF	1.039	1.015	1.020	0.988

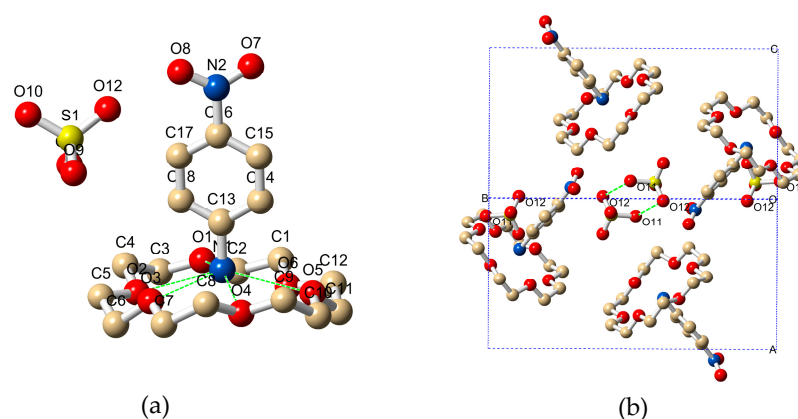


Fig. 1 Crystal structure of crystal **1** at 100 K. (a) Asymmetric unit for crystal **1**, dashed lines indicate hydrogen bonds. (b) Unit cell of crystal **1** viewed along the $a+c$ axis. Most hydrogen atoms on carbon atoms are omitted for clarity.

The crystal structures of **1** are similar between the LTP and HTP structures. This result reveals that the asymmetric unit was composed of one cationic [(4-HNA)(18-crown-6)]⁺ moiety and one anionic HSO₄[−] (Fig. 1a). The 4-HNA cations were connected to the 18-crown-6 ring to form a supramolecular rotator–stator structure through N–H⋯O interactions between the –NH₃⁺ group and the six O (O₁, O₂, O₃, O₄, O₅, and O₆) atoms of 18-crown-6. The average hydrogen-bonding N–O distances of 2.874 and 2.888 Å at 100 K and 296 K, respectively, were almost similar to those of the standard NH₃⁺⋯O distance for the crown ether molecular-based system (Table S1). The π -plane of the 4-HNA cation was nearly perpendicular to the mean plane of the oxygen atoms. The N₁ atom of the 4-HNA cations was located higher than the best plane of the oxygen atoms of the crown ring rather than in the nesting position (Fig. 1a). The dihedral angles between the aromatic ring and the crown ether ring were 92.48° (100 K) and 93.47° (296 K). In Fig. 1b, the packing diagram of complex **1** is located along the $a+c$ axis. The packing diagram for the LTP (100 K) indicates that the dimer structure of the two HSO₄[−] anions is linked by O–H⋯O hydrogen bonds that fill in the space to form four [(4-HNA)(18-crown-6)]⁺ supramolecular cations. In addition, the O–H⋯O hydrogen-bonding interaction becomes stronger than the N–H⋯O hydrogen bonds with bond distances of 2.647 Å at 100 K and 2.657 Å at 296 K, respectively. This finding indicates that the movement of H protons was more difficult between the donor (O₁₁) and acceptor atoms (O₁₀). The most significant differences between the structures at 296 K and 100 K were the distances of the two crown ether rings. Compared with the distances at 100 K, those at 296 K changed from 4.208 Å to 4.063 Å. C–H⋯ π interactions were noted in the [(4-HNA)(18-crown-6)]⁺ complex cations. The aromatic rings formed two C–H⋯ π interactions with distances of 3.195 Å at 100 K and 3.267 Å at 296 K (Fig. 2).

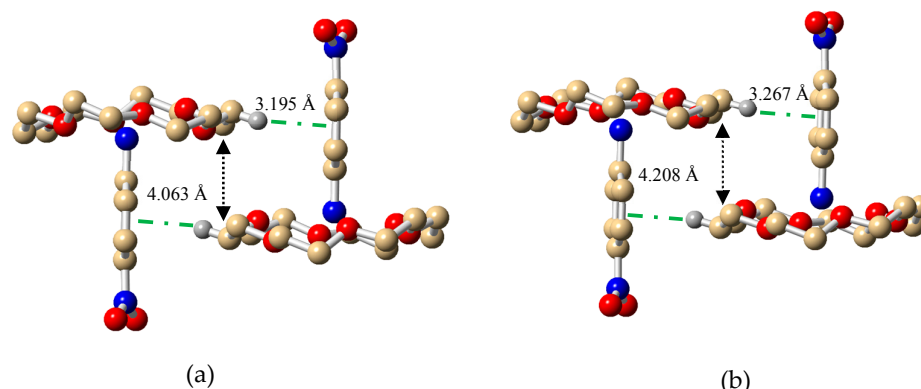


Fig. 2 Arrangement of the supramolecular cations in **1** at 100 K (a) and 296 K (b) viewed in the $a+c$ plane; the dashed lines show the C–H \cdots π interactions.

2.2 Crystal structure of 2

The crystal structure of **2** was analyzed in the HTP (296 K) and LTP (100 K) forms. When the HSO_4^- anion was replaced with the PF_6^- anion, compound **2** at both 296 K and 100 K crystallized in the monoclinic system with the same space group $P2_1/c$. Although the temperature changed, the space group of crystal **2** remained unchanged. Hence, no structural symmetry breaking occurred in this temperature range. Crystallographic data and details on the collection and refinement for 296 K and 100 K are listed in Table 1.

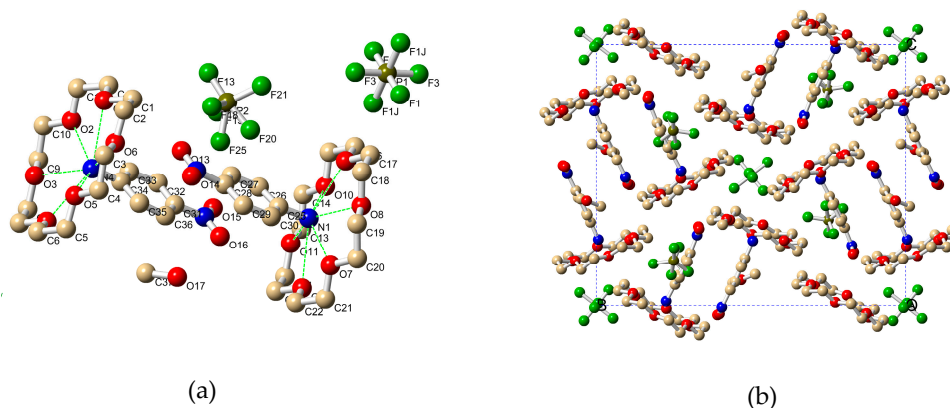


Fig. 3 Asymmetric unit (a) and packing diagram (b) of crystal **2** viewing along the a -axis at 100 K and 296 K. Most of the hydrogen atoms on the carbon atoms are omitted for clarity.

The asymmetric unit of compound **2** consisted of two independent $[(4\text{-HNA})(18\text{-crown-6})]^+$ supramolecular cations, two PF_6^- anions, and one methanol molecule at 296 K and 100 K (Fig. 3a). In contrast to the ordered HSO_4^- anions in **1**, the PF_6^- anions in **2** were disordered at both 100 K and 296 K. The molecular structure of **2** with atomic labeling is shown in Fig. 3a. The disordered PF_6^- anion and methanol molecules filled the structure between neighboring supramolecular cations (Fig. 3b). The most striking structural features in the LTP and HTP forms were the distances of the two crown ether rings. The distances of crystal **2** from the neighboring $[(4\text{-HNA})(18\text{-crown-6})]^+$ supramolecular cations were 8.877 Å (100 K) and 9.082 Å (296 K) (Fig. 4). The $-\text{NH}_3^+$ group resided in the perching position and attained a configuration similar to that of crystal **1**; the group was linked by the oxygen

atoms of the crown ethers through the six N–H⋯O hydrogen bonds. Apparent hydrogen bonding interactions occurred between the nitrogen and oxygen atoms with bond lengths of 2.821–2.956 Å and 2.816–2.966 Å at 100 K and 296 K, respectively (Table S2). In crystal **2**, the distance of two adjacent crown ether rings was nearly twice longer than that in crystal **1** at 100 K and 296 K (Fig. 4). In crystal **1**, the dimer structure of two HSO₄[−] anions occupied a larger space volume, resulting in a relatively closer packing pattern for the [(4-HNA)(18-crown-6)]⁺ supramolecular cation. Furthermore, no C–H⋯π interactions existed in the supramolecular cations. However, weak π⋯π interactions were found between the aromatic rings of the [(4-HNA)(18-crown-6)]⁺ supramolecular cations. These interactions stabilized the crystal packing and formed the alternated inorganic–organic hybrid structure in the *bc* plane (Fig. S5).

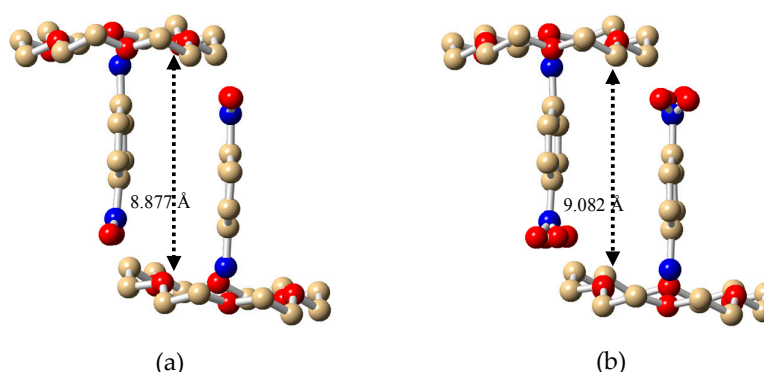


Fig. 4 Arrangement of the supramolecular cations in **2** at 100 K (a) and 296 K (b) viewed along the *a*-axis; the dashed lines show the distance of the adjacent two crown ether rings.

2.3 Differential scanning calorimetry

Differential scanning calorimetry (DSC) is commonly used to detect whether a compound displays a phase transition triggered by temperature. This approach is also used to confirm the existence of a heat anomaly during heating and cooling. When a compound undergoes a structural phase transition with a thermal entropy change, reversible heat anomalies are detected during heating and cooling. In the obtained DSC spectrum of crystal **1**, a main endothermic peak and a main exothermic peak were observed at heating of 257 K and cooling of 252 K, respectively, with a 5 K hysteresis width (Fig. 5a). These exothermic and endothermic peaks clearly reveal the occurrence of a reversible phase transition. The entropy change (ΔS) of the phase transition was exceedingly low to be estimated from the DSC. The wide heat hysteresis and peak shapes reflect the characteristics of a first-order phase transition. The driving force of the phase transition was confirmed by evaluating the crystal structure across varying temperatures. For example, apparent differences were observed for the distances of two crown ether rings, with values of 4.063 Å at 100 K and 4.208 Å at 296 K. Additionally, the dihedral angle between the aromatic ring and the crown ring slightly changed from 92.48° in the LTP form to 93.47° in the HTP form.

Compared with the DSC measurements for crystal **1**, those for crystal **2** were conducted within the temperature range 240–320 K and revealed two anomalies at 266 K (heating) and 261 K (cooling) (Fig. 5b). The thermal hysteresis was about 5 K, and the entropy change (ΔS) of the phase transition was also too low to be estimated from the DSC. The almost non-existent hysteresis and small heat

anomalies clearly reflect the continuous characteristic of the phase transition and thus effectively show features of a second-order phase transition. The distances of neighboring [(4-HNA)(18-crown-6)]⁺ supramolecular cations were 8.877 Å (100 K) and 9.082 Å (296 K) for crystal **2**. For these reasons, crystals **1** and **2** both revealed reversible phase transitions despite containing different anions. This finding implies that the phase transition may have originated from the [(4-HNA)(18-crown-6)]⁺ supramolecular cations, a finding observed in similar compounds.

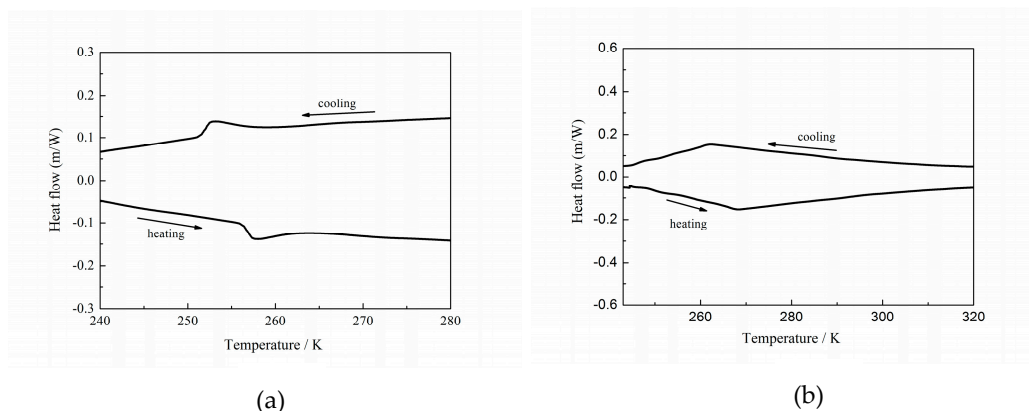


Fig. 5 DSC curves of crystal **1** (a) and crystal **2** (b) in a heating-cooling cycle.

2.4 Dielectric property

The variable-temperature dielectric response is another common method to study phase transitions, especially at relatively high frequency ranges. This method is useful for identifying phase transitions. The temperature-dependent dielectric constant of the powder samples of crystals **1** and **2** were presented on four selected frequencies, namely, 5 KHz, 10 KHz, 100 KHz, and 1 MHz. Strong and significant dielectric anomalies were observed around the T_c . For crystal **1**, the dielectric constant slowly increased with rising temperature below 250 K (Fig. 6a). Interestingly, when the temperature neared 255 K, the dielectric constant sharply increased to a maximum value of 22.5 at 5 KHz. Then, the dielectric constant abruptly decreased and achieved a minimum value of 11.91 at about 265 K. However, there is no distinct dielectric anomaly with the decrease in temperature. The sharp peak-like dielectric anomaly further confirms the phase transition in **1** and is consistent with the DSC result.

For crystal **2**, the dielectric constant slowly increased with an abrupt slope at around 265 K during heating (Fig. 6b). The maximum dielectric constant value was about 8.14 at 5 KHz; this value corresponds to a high dielectric state. The dielectric constant of **2** achieved a sudden rapid increase at room temperature. Compound **2** attained small DSC and dielectric anomalies at about 265 K; this pattern is a feature of a second-order phase transition. The relatively weak dielectric anomaly of crystal **2** relative to that of **1** is presented in the temperature ranges. This finding is probably due to the unchanged electric polarizations of ions and molecules in the crystal lattice.

The variable-temperature crystal structures of **1** and **2** reveal that the phase transition and dielectric anomaly in the same temperature range may be closed by the structural interactions and different crystal packing patterns in the LTP and HTP forms. Crystals **1** and **2** achieved significant differences at 296 K and 100 K, i.e., the distances between the two adjacent crown rings of **2** were 8.877 Å (100 K) and 9.082 Å (296 K). These distances were larger than the values of 4.063 Å (100 K)

and 4.208 Å (296 K) in **1**. This finding suggests that the inorganic anion tunes the crystal and hence affects the phase-transition points and types.

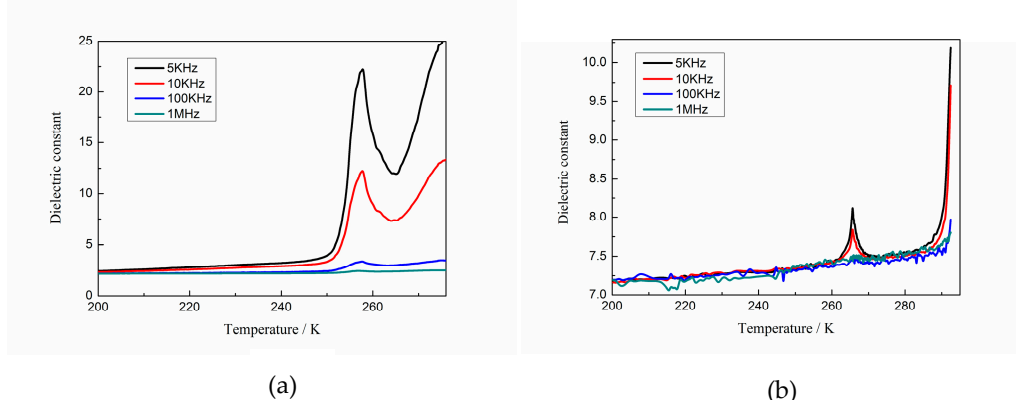


Fig. 6 Dielectric properties for crystal **1** (a) and **2** (b) measured as a function of temperature under a frequency of 5 KHz to 1 MHz.

3. Materials and Methods

3.1 Materials and instrument

The chemicals and solvents employed in this work were commercially obtained as chemically pure and used without any further purification. Infrared (IR) spectra were obtained using an Affinity-1 spectrophotometer and KBr pellets in the 4000–400 cm^{-1} region. Thermogravimetric analyses (TGA) were performed with a SHIMADZU DTG-60 thermal analyzer in a nitrogen atmosphere from room temperature to 800 K with a heating rate of 10 K/min using aluminum crucibles. Elemental analyses were conducted using a Vario EL Elemental Analysensysteme GmbH at the TRW Research Collaboration Center, YanZhou, Shandong. DSC measurements were performed by heating and cooling the samples (16 mg) within the temperature range 210–280 K on a TA Q2000 DSC instrument under nitrogen. Meanwhile, the crystal dielectric constants were measured with a TH2828 Precision LCR meter within the frequency range of 500 Hz–1 MHz, applied voltage of 1.0 V, and temperature sweeping rate of approximately 2 K/min.

X-ray single crystal diffraction: X-ray diffraction experiments were conducted on crystals **1** and **2** using a Bruker CCD diffractometer with Moka radiation ($\lambda=0.71073$ Å) at 100 K and 296 K. The structures of **1** and **2** were elucidated by direct methods and refined by the full-matrix method based on F^2 using the SHELXL 97 software package. All non-hydrogen atoms were anisotropically refined, and the positions of all hydrogen atoms were geometrically generated. CCDC:1529205 (100 K), CCDC:1529207 (296 K) for **1** and CCDC:1529231 (100 K), CCDC:1529232 (296 K) for **2** contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Center via www.ccdc.cam.ac.uk/data_request/cif

3.2 Preparation of compound (4-HNA)(18-crown-6)(H₂SO₄) (**1**)

Yellow crystals **1** were obtained by evaporating an alcohol solution containing 4-nitroaniline (20 mg), H₂SO₄ (50 mg), and 18-crown-6 (200 mg) at room temperature for 5 days to a 52% yield (based on 4-nitroaniline). The calculated percentage compositions of C, H, and N for C₁₈H₃₂N₂O₁₂S were 43.19% C, 6.44% H, and 5.60% N, respectively, whereas the measured percentage compositions were

43.11%, 6.42% H, and 5.76% N, respectively. The IR spectrum of the single crystal **1** is given in Fig. S1.

3.3 Preparation of compound (4-HNA)₂(18-crown-6)₂(PF₆)₂(CH₃OH) (**2**)

Single crystals of **2** were prepared by slowly evaporating a mixture of HPF₆ (50 mg), 4-nitroaniline (20 mg), and 18-crown-6 (200 mg) in methanol solution (50 mL). The methanol solution was allowed to stand for approximately 5 days under room temperature. The single crystals of salt **2** were colorless transparent crystals obtained at 58% yield. The calculated percentage compositions of C, H, and N for C₃₇H₆₆F₁₂N₄O₁₇P₂ were 39.37% C, 5.89% H, and 4.96% N, respectively, whereas the measured percentage compositions were 39.35% C, 5.87% H, and 4.89% N, respectively. The IR spectrum (KBr) of the single crystal **2** is presented in Fig. S2.

4. Conclusion

Two inorganic–organic hybrid crystals based on the [(4-HNA)(18-crown-6)]⁺ supramolecular cation were reported in this study. Variable-temperature crystal structure analyses and thermal measurements (DSC) showed that crystals **1** and **2** exhibited similar crystal packings and underwent reversible phase transitions at 255 K and 265 K, respectively. The dielectric anomalies of crystals **1** and **2** in the temperature–frequency ranges further confirmed the existence of such phase transitions. These results indicated that the phase transitions may be caused by the [(4-HNA)(18-crown-6)]⁺ supramolecular cations. The inorganic anions (PF₆[−] and HSO₄[−]) played an important role in the crystal packing and regulated the phase-transition points and types.

Supplementary Materials: IR spectra and TG data of crystals **1** and **2** (Figs. S1–S4); Arrangement of the supramolecular cations of crystal **2** (Fig. S5); Hydrogen bond geometry of crystals **1** and **2** (Tables S1–S2).

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Conflicts of Interest: The authors declare no conflict of interest.

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