

Article

Optimal Surface Amino-Functionalization Following Thermo-Alkaline Treatment of Nanostructured Silica Adsorbents for Enhanced and Selective CO₂ Adsorption

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Abstract: Special preparation of SBA-15 materials has been carried out for creating adsorbents exhibiting an enhanced and selective adsorption toward CO₂. This creation starts from an adequate conditioning of the silica surface, via a thermo-alkaline treatment to increase the population of silanol species on it. CO₂ adsorption is only reasonably achieved when the SiO₂ surface becomes aminated after put in contact with a solution of an amino alkoxide compound in the right solvent. Unfunctionalized and amine-functionalized substrates were characterized through X-ray diffraction, N₂ sorption, Raman spectroscopy, electron microscopy, ²⁹Si solid-state NMR, and NH₃ thermal programmed desorption. These analyses proved that the thermo-alkaline procedure desilicates the substrate and eliminates the micropores (without affecting the SBA-15 capillaries), present in the original solid. NMR analysis confirms that the hydroxylated solid anchors more amino functionalizing molecules than the unhydroxylated material. The SBA-15 sample subjected to hydroxylation and amino-functionalization displays a high enthalpy of interaction, a reason why this solid is suitable for a strong deposition of CO₂ but with the possibility of observing a low-pressure hysteresis phenomenon. Contrastingly, CH₄ adsorption on amino-functionalized, hydroxylated SBA-15 substrates becomes almost 5 times lesser than the CO₂ one, thus giving proof of their selectivity toward CO₂.

Keywords: CO₂ enhanced capture; SBA-15 habilitation for CO₂ sorption; desilication; silanol functionalization; covalent coordinated CO₂ deposition

1. Introduction

1.1 The problem

The effective and abundant capture of CO₂ on the surface of nanoparticulate substrates is a fundamental experimental aim, since this gas plays a leading role in the upsurge of the greenhouse effect; this last phenomenon being nowadays one of the most important ecological problems to deal with. There already exist several strategies for achieving satisfactory CO₂ levels of capture; one implementation of particular and novel interest consists in the trapping and eventual in situ transformation of this gas into useful and environmentally friendly compounds, via chemical procedures. Hence, with the objective in mind of reaching an efficient carbon capture and storage (CCS) [1-3], three alternatives are available at present: (a) CCS pre-combustion of carbon sources before actual fuel combustion, (b) oxycombustion (i.e. carbon compound combustion under nearly

pure oxygen), and (c) post-combustion capture. Carbon capture after combustion is unavoidable due to the amount of industrial processes that depend on the overwhelming burning of fossil fuels; an example of this being cement production among many other processes involving the burning of huge amounts of carbon. Nowadays, the vast majority of CCS research is related to the occurrence of the following phenomena: (a) absorption, (b) cryogenic separation, (c) separation through membranes, and (d) adsorption. The materials employed to perform after combustion CCS operations require, besides of being selective, of some other crucial characteristics such as: (i) to be endowed with a high compound capture capacity, (ii) to be renewable, (iii) low cost, (iv) durable, and (v) practical.

1.2 CCS alternatives

In particular, several sorption experiments are very useful as CCS alternatives. According to the nature of the substrate employed, CO₂ sorption or capture occurs through two possible processes:

- Absorption processes, which involve the bubbling of the adsorptive in solutions containing species such as amines, carbonates, and ions. In this case, the solvent becomes very important, since in order to achieve a good capture of CO₂, solvents of a low polarity and relatively low vapour pressure are very much preferred. Nevertheless, absorption processes involve a great disadvantage that is concerned with the short lifetimes that the trapped species can attain as these molecules usually interact (react) with the solvent species.

- Adsorption processes implying the employment of structures such as zeolites, hydrotalcites, polymer fibres, activated carbons, metal oxide networks, functionalized nanoporous materials, and, recently, metal oxide frameworks (MOFs).

1.3 CO₂ Chemisorption and Physisorption

The substrates that are more usually chosen for CO₂ chemisorption capture are metal oxides (*e.g.* CaO and MgO); these solids can form carbonates in the presence of CO₂. However, a shortcoming arises with the employment of these substrates, in view of the fact that high temperatures are required for achieving their total regeneration, *i.e.* 287 °C and 800 °C for MgO and CaO, respectively [4]. Lithium salts, such as Li₂ZrO₃ and Li₄SiO₄ [5, 6], have been used to capture CO₂ (between 450 and 700 °C) thus apparently being good alternatives for reaching an efficient CO₂ adsorption, irrespectively of the high production cost involved with this kind of substrates.

Another fine alternative for effective CO₂ chemisorption is the use of hydrotalcites, ever since León *et al.* [7] demonstrated that the irreversible reaction between these minerals and CO₂ is due to the creation of unidentate adsorption sites. Likewise, the formation of bidentate sites, generate surface bicarbonates, which behave as weak adsorption sites, useful for highly reversible reactions. The concentration of the adsorbate phase is closely related to the number of irreversible adsorption sites.

Active carbon possesses attractive characteristics for the physisorption of diverse species such as high uptake capacity, hydrophobicity and low cost; however, this substrate shows a poor selectivity between CO₂ and N₂. Lu *et al.* [8] observed that carbon nanotubes depicted a higher CO₂ adsorption than zeolites when surface functionalized with amine groups (NH₂) through the attachment of 3-aminopropyltriethoxysilane (APTES). Due to the previous arguments, a convenient adsorbent for CO₂ capture depends on the temperature and pressure at which the adsorption

process is to take place. According to these requirements, surface-functionalized mesoporous materials stand apart from the rest of adsorbents because of their ability to physisorb or chemisorb diverse compounds, depending on their physicochemical interactions (weak or strong) with the adsorbent.

Recently, the implementation of MOFs has brought about great advantages: (i) a good structural stability, and (ii) a high porosity this characteristic being reflected in surface areas of about $2900 \text{ m}^2 \text{ g}^{-1}$ and large total pore volumes between $1\text{--}2 \text{ cm}^3 \text{ g}^{-1}$ [9] if compared to solids such as zeolites and molecular sieves. However, a great disadvantage surges when trying to employ MOFs for CO_2 adsorption, in view of the low selectivity of these solids towards CO_2 , especially when this adsorptive is mixed with other gases. Additionally, these materials are overwhelmed by high synthesis costs; furthermore, the capture capacity of MOF's decreases under humid conditions, even though this last event could be dealt with through surface functionalization via diverse organic groups [10–12].

2. Stages followed concerning the choice and implementation of proficient silica adsorbents for enhanced CO_2 capture

2.1 Adsorbent choice

The CO_2 molecule is a Lewis acid since this molecule can accept a pair of electrons from a base in order to establish a covalent coordinated bond. Therefore, the crucial steps for designing a proper CO_2 adsorbent include the selection of a substrate whose surface can be habilitated to possess a good number of basic sites; i.e. the surface has to be endowed with chemical species that can be combined with other matching molecules for attaining the desired characteristics.

Nanoporous silica supports [13–15] are provided, in general, with a high surface area and a considerable total pore volume. These characteristics make these substrates appropriate for adsorbing an assortment of chemical species. The employment of amphiphilic triblock copolymers as pore templating agents has allowed the creation of well-ordered mesoporous networks, which have been labeled as SBA-n (Santa Barbara) materials [16]. Among them the SBA-15 [17] substrate results to have an ideal surface for performing an efficient and enhanced CO_2 adsorption. On the one hand, this material is formed by a collection of cylindrical capillaries, all practically having the same diameters, in a hexagonal arrangement and lying parallel to each other along micrometric lengths. On the other hand, the capillaries open at both ends that are forming these mesoporous structures are endowed of such sizes as to allow the free transit of different molecules along the nanocylinder axis. This is a very important property since the adsorptive molecules can enter, while the adsorbate molecules can exit the structure easily.

Besides, the functionalization of the surface via, as for instance, amino-alkoxides (i.e. the type of compounds that we will be employing for the amination of the SiO_2 surface) can also benefit from this size characteristic, given that the deposition of molecules usually occur after an adsorption from solution process. Another important property of amorphous silica is the presence of silanol (Si-OH) species throughout the surface. These species are determinant when the surface is to be functionalized with a given molecule. In the case of CO_2 , this molecule is in itself a weak Lewis acid, while the Si-OH groups are likewise. Then, in order for the silica surface to capture CO_2 , it has to be functionalized with pertinent species, e.g. amino-substituted alkoxides. The alkoxides groups can be attached to the surface after reacting with the silanol species; in turn, the CO_2 molecules become

coordinated to the NH_2 groups of the amino-functionalizing compound.

2.2 Habilitation of the SiO_2 surface to attain an efficient and sufficient amine-functionalization process.

Silica by itself manifests no significant affinity toward CO_2 adsorption, for this reason, it is necessary to perform the chemical modification of this surface. This transformation can rely on the silanol groups, since the character and concentration of these species are crucial for attaining a widespread amino-functionalization of the SiO_2 surface. Importantly, because of the thermal treatment required to eliminate remnant pore templating species, a relatively small population of silanol groups ($-\text{SiOH}$), subsists in the final network. Therefore, a hydroxylation process via treatment of the SiO_2 substrate with a lowly concentrated NaOH solution, set at the right temperature, has been found to be a suitable operation for achieving an efficient hydroxylation of the surface. The reason of the low NaOH concentration to be used in this process resides in preventing an excessive desilication (i.e. dissolution, leaching) of the SiO_2 matrix when put in contact with an alkaline solution [18]. Nonetheless, the previous chemical procedure can still result useful, not only for attaining more SiOH species on the surface, but to eliminate the micropores present in the parent silica material.

2.3 Supplementary aspects in relation to the choice of SBA-15 for enhanced CO_2 trapping

Other aspects related to the convenience of choosing SiO_2 as substrate for CO_2 adsorption are the possibility of readily regenerating the adsorbent surface by means of a mild thermal treatment in an oven, thus reducing the power requirements necessary to perform the surface regeneration process. Another striking characteristic is the selectivity that the silica surface shows toward the adsorption of CO_2 instead of adsorptives such as CH_4 .

Furthermore, the deposition of a functionalizing molecule on the silica surface can be made from either an organic or inorganic solution of the species in question. In the case of amine-functionalization it is better to perform this process employing an appropriate organic liquid for dissolving the amino alkoxide compound chosen for functionalizing the surface, e.g. toluene instead of ethanol. The reason of this is the chemical character of the silanol groups, which can interact synergistically or not with the solvent molecules with respect to an effective deposition of the functionalizing compound [19].

The objective of this work consists in showing evidence of a more abundant and efficient CO_2 capture on the basis of some appropriate textural changes that can be inflicted on model SBA-15 silica adsorbents via thermal and chemical surface functionalization procedures. A concomitant aim is to increase the number of silanol groups (i.e. $-\text{Si-OH}$) on the silica surface, since this is a key aspect for achieving an efficient functionalization of the surface via chemical anchorage methods. The ultimate goal is to accomplish an abundant capture of miscellaneous adsorptives (other than CO_2) for their storage and eventual transformation into other molecules of diverse practical interest.

This work is organized as follows. First, the experimental section detailing the synthesis of the silica substrates is presented; next, the thermo-alkaline habilitation of the surface of these solids is described; the following step consists in carrying out the functionalization of the previous materials with an especially selected amino-alkoxide compound. Subsequently, N_2 sorption isotherms at 76 K are shown together with a list of the most important structural parameters of the porous solids involved in this investigation. CO_2 adsorption isotherms, the center of this research, in the range of

263–313 K on the diverse SBA-15 materials are presented. Additionally, FTIR, XRD, TEM, TGA, solid NMR, and TPD characterization experiments are depicted. Finally, CH₄ adsorption isotherms at 298 K on some of the SBA-15 adsorbents are presented, in order to ascertain the selectivity of these substrates between this last adsorptive and CO₂.

3. Experimentation

3.1. Materials

Tetraethyl orthosilicate, TEOS (Aldrich, 98%), hydrochloric acid (J.T. Baker, 36.5–38.0% w/w, poly(ethylene glycol) Aldrich (for molecular biology), extra pure sodium hydroxide pellets (Riedel-deHaën, 99–100%), 3-Aminopropyltriethoxysilane (APTES, of Sigma-Aldrich, 99%), toluene (Merck) and denatured ethyl alcohol 95% v/v, Pluronic 123 (BASF).

3.2. Synthesis of precursory SBA-15 solids

The following molar ratio sequence was employed to prepare SBA-15 silica: TEOS: 197H₂O: 6.2HCl: 0.017P123, the synthesis was performed according to the method devised by Zhao et al. [13, 20]. The ensuing material was washed, filtered, put in contact with ethyl alcohol under stirring, and dried for 14 h at 373 K. Finally, the solid was calcined at 723 K for 6 h, while employing a heating ramp of 1 K/min.

3.3. Soft thermo-alkaline treatment performed over SBA-15 precursory materials

The mesoporous silica adsorbents of the SBA-15 kind were subjected to a thermo- (313 K) alkaline treatment through the employment of diluted (< 1M) NaOH solutions. Specifically, a 250 mg sample of each S15 substrate was subjected to a series of treatments under the previous conditions, each of these operations taking place every two hours. Subsequently, a desilication (i.e. hydroxylation) process was performed by adding a 0.025 M NaOH solution at 313 K (2 h). The resultant solid was washed with deionized water, filtered, and dried at 373 K for 2 h. The ensuing substrate was labeled as S15H.

3.4. Chemical anchoring of APTES on the surface of SiO₂ mesoporous solids

The chemical bonding of APTES on the surfaces of the S15 and S15H silica substrates was achieved once these solids were calcined in air at 723 K for 2 h. The advantage of this route is to guarantee the total removal of the surfactant molecules after the substrate has been annealed at 723 K.

Previously to the APTES surface deposition, the precursory SiO₂ solids were treated at 473 K under vacuum for 2 h in order to remove organic impurities and also to promote the activation of the silanol (SiOH) groups already residing on the surface. The molar ratio involved in this functionalization reaction corresponded to 3SiO₂: 2APTES and this procedure was carried out at 353 K under a N₂ atmosphere and continuous magnetic stirring of a reflux system during 20 h, while employing toluene as the APTES dispersion medium. The final solids were separated by filtration and dried at 373 K for 2 h. These materials were labeled as S15NH₂ and S15HNH₂, given that these aminated materials proceed from the respective S15 and S15H samples.

3.5. CO₂ sorption experiments at different temperatures

Prior to CO₂ sorption experiments, each sample was outgassed at 373 K under high vacuum for 6 h. The adsorption temperature range utilized for these experiments was chosen between 273 K and 298 K in step intervals of 10 K. Finally, the adsorption capacity of the amine-functionalized SiO₂ substrates was determined experimentally, in an automatic adsorption instrument (Quantachrome Autosorb 1LC). Additional textural and thermal (isosteric heat of adsorption) properties were calculated from the corresponding CO₂ adsorption isotherms evaluated at the different temperatures. CH₄ adsorption at 273 K, on some of the silica samples, was also carried out with the aim of visualizing the selectivity of the modified SBA-15 adsorbents towards CO₂.

4. Characterization Techniques

- **FTIR** This technique will be helpful for realizing the presence of silanol and siloxane species on the surface of the adsorbents. FTIR signals were obtained from a Perkin Elmer Paragon 1000 (FT-IR) instrument equipped with an Attenuated Total Reflectance (ATR) tool.

- **TGA** This analysis was performed on a Perkin Elmer Diamond instrument on a N₂ flow of 50 mL min⁻¹, subjected to a temperature ramp of 5°C min⁻¹ along a temperature interval of 323–823 K.

- **Raman**. These spectra are useful for providing an overall view of the structural consolidation of the adsorbents prepared in this work; in addition to the structure of the substrate itself, in particular, of silica bonds as well as free or vicinal silanol groups that lie on the silica surface. Through this technique it could be possible, in principle, of realizing the disposition of adsorbate molecules throughout the adsorbent surface. Raman Microscope (Thermo Scientific) with 636 nm laser delivering a maximum power of 10 mW.

- **TEM** images were obtained from a HRTEM Jeol 2100F microscope operating at an acceleration voltage of 200 keV, in order to observe the structural arrangement of the samples and to realize if the preparation procedures have affected (or not) the shape and structure of the SiO₂ substrates.

- **XRD** was not only useful for checking the ordered hexagonal arrangement of the nanotubes making the SBA-15 hexagonal arrangement but also for determining the pore width and the thickness of their pore walls. The XRD parameters of the network were measured from a Bruker D8 Advance instrument employing a Cu-K α radiation wavelength of 1.54 Å.

- **N₂ Sorption**. The textural properties of the silica adsorbents were determined from an ASAP 2020 automatic instrument, employing N₂ at its boiling point (76 K at Mexico City's altitude). The materials were previously outgassed at 373 K for 12 h. under a turbomolecular vacuum of 10⁻⁶ mbar. The surface area was determined from either the BET or *t*-methods, while the pore size distribution was obtained through the Non-Localized Functional Theory (NLDFT) approach. The kernel employed for this calculation was that corresponding to the filling of cylindrical pores along the boundary N₂ adsorption curve at 77 K.

- **NH₃ TPD**. The total acidity (Brönsted + Lewis) was obtained from a Micromeritics TPD/TPR 2900 NH₃ programmed desorption (TPD) device provided with a TCD detector. The experimental procedure consisted in outgassing the sample under a mixture of N₂ and air from room temperature up to 773 K, the system was then kept at this temperature for 15 min and the heating was continued under NH₃ flux to the final temperature of 1073 K. Subsequently, the system is left cooling down to 303 K for about one hour. The definitive TPD experiment was carried out under a NH₃ flow

following to a heating ramp of $10^{\circ}\text{C min}^{-1}$ up to 1073 K; while registering the amount of NH_3 desorbed at each temperature.

- **Solid NMR.** Solid State Nuclear Magnetic Resonance experiments were run on a Bruker Advance II300 spectrometer, operating at 59.62 MHz for ^{29}Si . The magic angle spinning technique was altogether employed with the NMR results as a reference of tetramethylsilane (TMS). This technique is very important in order to ascertain the number of bonds established between the surface silanol groups and the ethoxy groups of the APTES functionalizing species.

- **CO_2 and CH_4 adsorption and selectivity.** Finally, CO_2 and CH_4 adsorption studies at different temperatures were performed on a Quantachrome Autosorb-1LC instrument; for this task, the samples were previously outgassed at 373 K for 6 h.

5. Results and Discussion

5.1. Middle-infrared spectroscopy studies

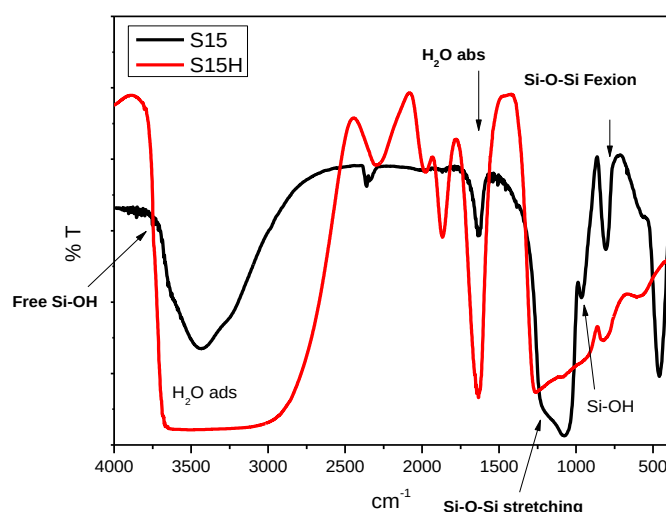


Figure 1. FTIR spectra of: (a) SBA-15 precursor silica (S15), and (b) SBA-15 sample after being treated with 0.025 M NaOH at a temperature of 313 K (S15H).

The FTIR spectra in **Fig. 1** evidence vibrational changes of the molecules; therein, it is possible to note a prominent signal widening corresponding to physisorbed water in the S15H samples subjected to the hydroxylation treatment, if compared to the spectrum of the S15 precursor specimen. This difference surges even if all samples were previously annealed under vacuum at 473 K for 2h. On the other hand, the intramolecular Si-OH (963 cm^{-1}) and Si-O-Si flexion (802 cm^{-1}) signals become sharper. However, for hydroxylated samples, the band assigned to free silanol groups (Si-OH) appearing at 3769 cm^{-1} [21-22] results masked by the widening of the vibrational band of physisorbed water. Nevertheless, a larger amount of free SiOH groups sites interact with water through hydrogen bonds and promotes a rapidly hydration, which can be reflected in a larger amount of physisorbed water, again with respect to the same quantity in the S15 precursor sample.

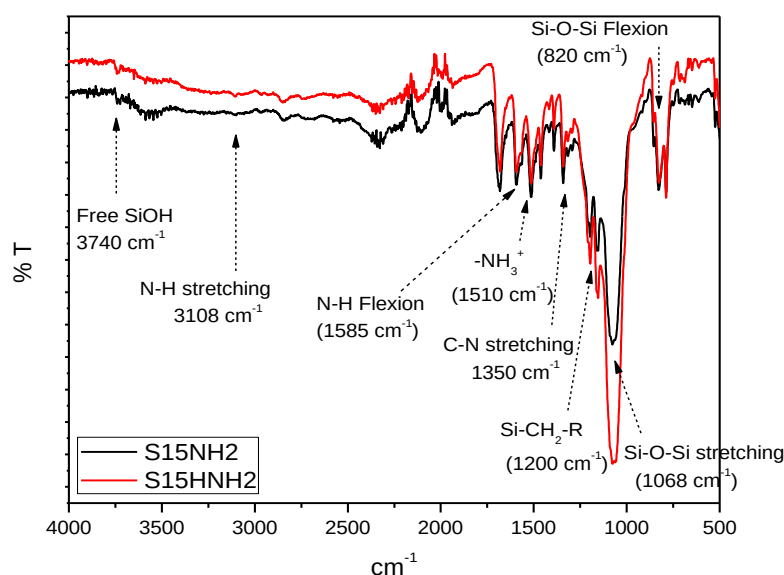


Figure 2. FTIR spectra of SBA-15 materials: (a) S15precursory and (b) S15H SBA-15 + NaOH 0.025 M samples, both characterized after functionalization with the same amount of APTES.

The samples under scrutiny were now functionalized through a chemical bonding method, i.e. a process employing APTES and that was already described in the experimental section. In this case, a covalent bond is established between the surface of the substrate and the bridging amino species (i.e. APTES), see **Fig. 2**. In this FTIR spectrum, it is possible to note silica vibrational bands as well as typical primary amine bands ($-\text{NH}_2$), as in the case of APTES [23–24]. Notice that the bands corresponding to S15H, substrates previously subjected to a treatment with 0.025 M NaOH are the most intense of all, especially if these signals are compared to those obtained in the parent S15 specimen that is subjected to no alkaline exposure.

5.2 Raman Spectra of precursory and thermo alkaline treated silica samples

Raman spectra are useful in order to provide additional structural characterization, besides of that corresponding to the substrate itself, in particular to evidence the presence of free or vicinal silanol groups that are lying on the silica surface.

The traditional Raman spectra of vitreous SiO_2 are characterized by bands appearing at around 410, 800, 1065, and 1200 cm^{-1} and that are associated to vibrations of the amorphous silica network [25]. In our case, however, the bands that exist at around 410, 500, and 610 cm^{-1} can be linked to the vibrations of siloxane groups (Si-O-Si), which are assigned to rings of diverse sizes [26]. The signal at around 410 cm^{-1} is associated to the bending mode vibration of oxygen atoms inside a ring made of more than 4 silicon atoms ($n > 4$); this peak is usually known as the R band. The signal at around 490 cm^{-1} is associated to the breathing mode vibration of rings depicting 4 silicon atoms; this signal usually being known as the D_1 band; the peak appearing at around 605 cm^{-1} is associated to the breathing mode vibration of rings formed by 3 silicon atoms, and branded as D_2 bands [27]. Additionally, the peak at around 980 cm^{-1} is associated with the vibrational mode of (OH)-groups with respect to Si.

In the Raman spectra of the pristine S15 sample, it can be observed shallow R, D₁, and D₂ bands which are associated to rings of diverse sizes existing in the silica network (**Fig. 3**). Additionally, the two bands appearing at around 850 and 1060 cm⁻¹ can be linked [28c] to the emissions located at around 850, 900, 950-1000, and 1050-1100 cm⁻¹ and that can be due to tetrahedral silicon atoms having four, three, two or only one bridging oxygen atoms, respectively [28]. Non-bridging oxygen atoms could correspond to hydroxyl groups or to different metals, such as polysilicate alkaline cations [28c]. In the Raman spectrum of the alkaline treated sample (S15H), an intense D1 band appears at around 490 cm⁻¹. Besides, an intense and broad band appears from 600 to 1100 cm⁻¹. This signal pathway suggests the existence of a slightly high population of silica rings formed by four silicon atoms. Furthermore, the intense band located at higher frequencies suggests the presence of a higher amount of tetrahedral silicon atoms having no bridging oxygens as consequence of the alkaline treatment to which the samples were subjected.

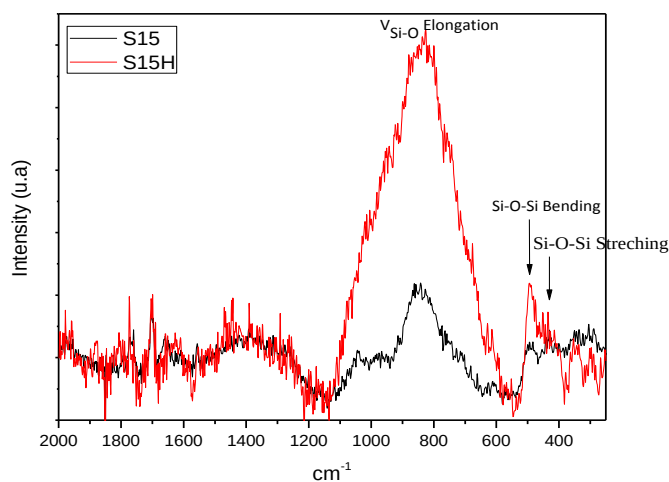


Figure 3. Raman spectra of SBA-15 materials: /a) precursory SBA-15 substrate (S15) and (b) silica after being treated with NaOH 0.025 M (S15H).

5.3 TGA analysis of precursory and functionalized silica samples

The S15 precursory sample was subjected to TGA analysis (**Fig. 4**). The weight loss of 6.0 wt. % observed at around 373 K can be due to water desorption. Afterwards, a slightly decreasing slope is developed as the temperature is raised and in such a way that when the analysis is fulfilled at 823 K an additional mass loss of 2.0 wt % is detected. A slight mass loss of about 2.0 wt.% is possibly due to the partial dehydroxylation of the SiO₂ matrix, or in other words, the Si-OH groups condense to generate water molecules and Si-O-Si siloxane species, consequently the total mass loss is about 8.0 wt.%. Nonetheless, for the S15H sample subjected to hydroxylation, a mass loss of 11.0 wt. % arises thus confirming that this sample was more hydrated than the precursory solid. This, in principle, may be due to the fact that the hydroxylation treatment causes the rupture of siloxane groups to generate vicinal Si-OH groups, which can be easily rehydrated and, as the temperature is raised during the TGA analysis, become dehydrated promoting, at the same time, the condensation of them. Since as it is well known, silica hydration-condensation processes are reversible phenomena [29].

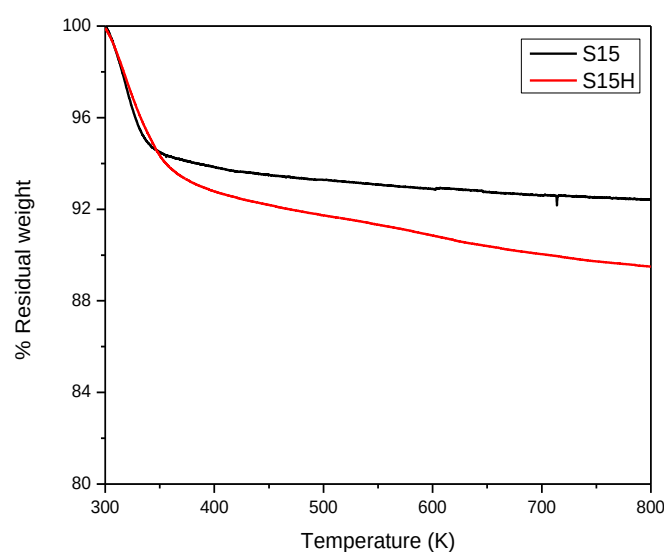


Figure 4. Thermogravimetric analysis under a N₂ atmosphere of the S15 parent sample and of the S15H solid proceeding from the corresponding alkaline hydroxylation of the previous substrate.

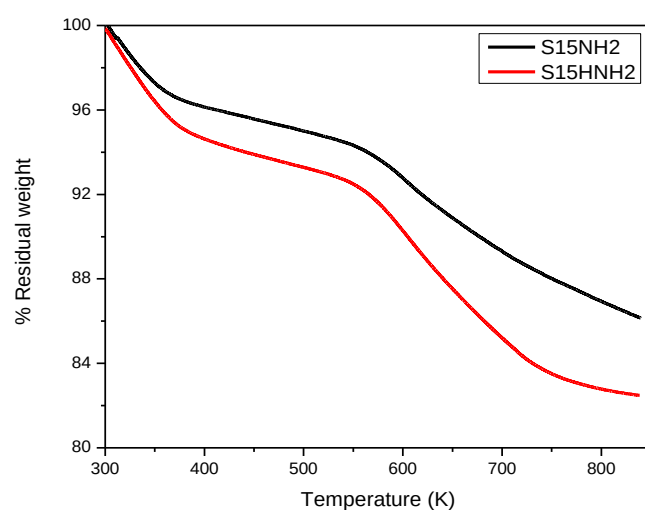


Figure 5. Thermogravimetric analysis, under N₂ flow, of the S15NH₂ and S15HNNH₂APTES-functionalized silica specimens.

The TGA study corresponding to the samples functionalized with APTES are shown in Fig. 5; namely, for the S15NH₂ and S15HNNH₂ substrates. The S15HNNH₂ hydroxylated simple suffers the highest mass loss of all. The first TGA region is linked to desorption of the physisorbed water at around 373 K; i.e. 3.0 wt. % for the S15NH₂ specimen, and 5.0 wt. % for the S15HNNH₂ solid. Nevertheless, the significant mass loss depicted from 543 K onwards, is caused by APTES thermal degradation. Specifically, the S15NH₂ material presents a mass loss of 14.1 wt. % while the S15HNNH₂ solid corresponds to a mass loss of 17.5 wt. %; this implies a difference between these specimens of 19.4 wt. %, which evidences that a major weight loss occurs for the hydroxylated and subsequently APTES functionalized simple.

5.4 XRD analysis at low angle

The XRD diffractograms obtained for both precursory and functionalized silica materials are shown in **Fig. 6**. Therein, it is observed for all cases, the structural parameters of a pore network consisting of cylindrical pore in an hexagonal packing with relatively sharp peaks appearing at (100), (110), and (200) planes, which are proof of this assertion.

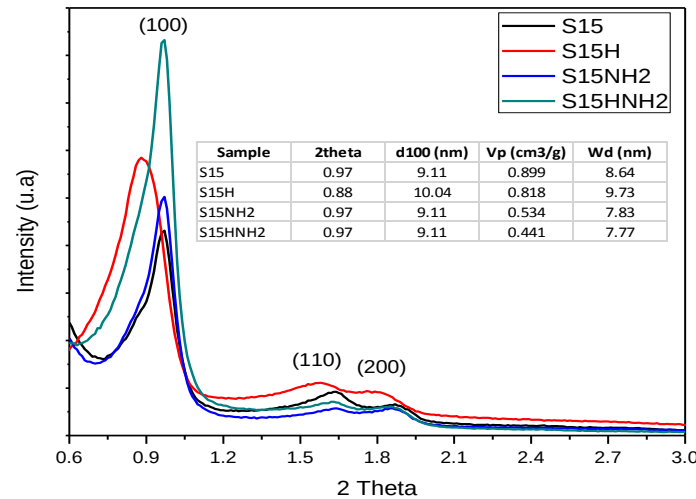


Figure 6. Low angle XRD diffractograms of: S15, S15H, S15NH2, and S15HNNH2 solids

As can be seen in **Fig. 6**, there is no change in the interplanar 100 distances depicted by both S15 precursory and S15NH2 functionalized samples; therefore, it is clear that the presence of APTES in the interior of the S15NH2 pore array. However, when comparing the W_d pore sizes obtained from Eqn. 1 [30] for the S15 and S15NH2 specimens, it is readily realized an evident pore size diminution provoked by the APTES presence inside the pores of the SBA-15 substrate.

$$W_d = 1.213 \cdot d_{100} \left(\frac{V_p}{\frac{1}{\delta} + V_p + V_{mic}} \right)^{\frac{1}{2}} \quad (1)$$

Where V_p and V_{mic} are the mesopore and micropore volumes, respectively, δ is the silica substrate density, and d_{100} is the 100 interplanar distance obtained from the XRD pattern.

In the case of the samples treated with alkaline solution, it can be noted a slight increase in the interplanar distance; this effect confirms that the hydroxylation treatment increases the pores size of the precursory solid without significantly altering the array and shape of cylindrical pores that constitute void volume of the substrate. Contrastingly, those samples functionalized with APTES suffered partial structural disarray, thus concluding that the S15HNNH2 sample accept a highest amount of chemisorbed APTES in comparison to the S15NH2 sample, even if the two samples are put in contact with the same 3SiO₂:2APTES molar ratio. This suggests that the hydroxylation process favoured a large generation of free Si-OH sites. It is thus adequate to stress that the abundance and distribution of OH groups on the surface of silica adsorbents determine the occurrence of subsequent surface reactions.

5.5 Transmission electron microscopy (TEM)

The TEM micrographs shown in **Fig. 7** corroborate that the S15 precursory specimen consists of an ordered array of cylindrical capillaries disposed in a hexagonal packing, something that is evident in these images.

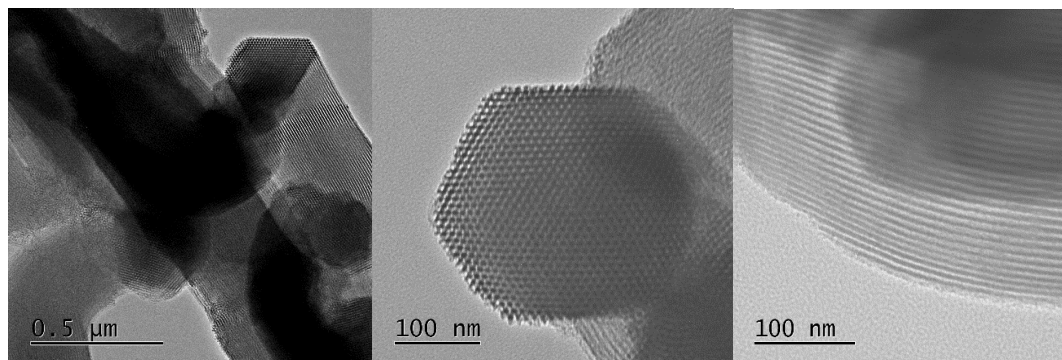


Figure 7. TEM images of the S15 precursory silica specimen.

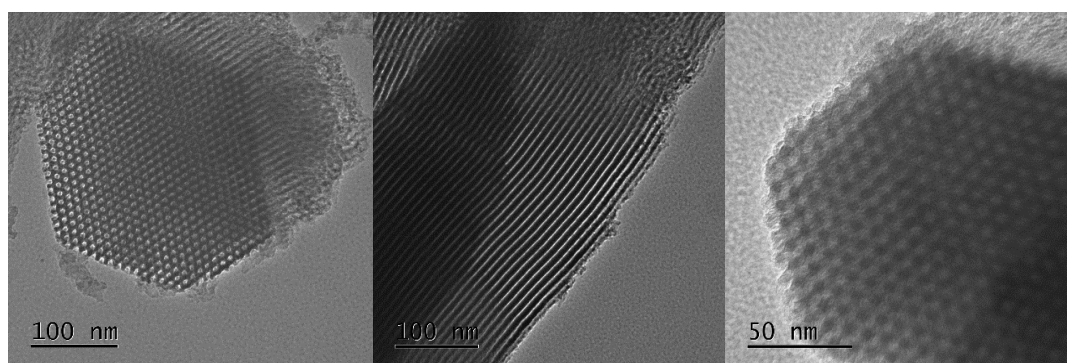


Figure 8. TEM images of the SBA-15 original sample, previously to the NaOH mild hydroxylation treatment, and of the S15H hydroxylated specimen.

On the other hand, when the hydroxylation treatment is performed S15H, there exist no obvious changes of the ordered silica pore arrangement with respect to the precursory S15 solid (**Fig. 8**). This indicates, in a way similar to that suggested by the DRX analysis, that the sample subjected to the mild hydroxylation process, previously described, preserves the arrangement of its cylindrical voids and that only a slight lixiviation phenomenon occurs in a relatively uniform and soft manner.

5.6 N_2 adsorption at 76 K

In **Fig. 9**, N_2 sorption isotherms at 76 K on precursory, S15, and hydroxylated, S15H, specimens and on their respective APTES functionalized solids (samples S15NH₂ and S15H₂NH₂) are displayed. The isotherms correspond to an IUPAC Type IV shape with an associated H1 hysteresis loop [31–32]. The figure inset in **Fig. 9**, shows the pore-size distributions (PSD) of the same samples obtained by means of the Non Local Density Functional Theory (NLDFT) treatment for N_2 desorption taking place along the boundary descending curve and according to a cylindrical pore model. Even if the dihydroxylation process is somewhat intense, because of lixiviation, the concentration of the caustic compound that was used to accomplish this end was sufficiently low as to keep the pore structure mostly unaffected, with practically no significant deformation of the pore shape. This can be also assessed by the close appearance of the XRD peak analysis. It is also pertinent

to mention that the alkaline treatment, in the case of the S15 substrate, caused a pore width increment of 0.29 nm (see **Table 1**).

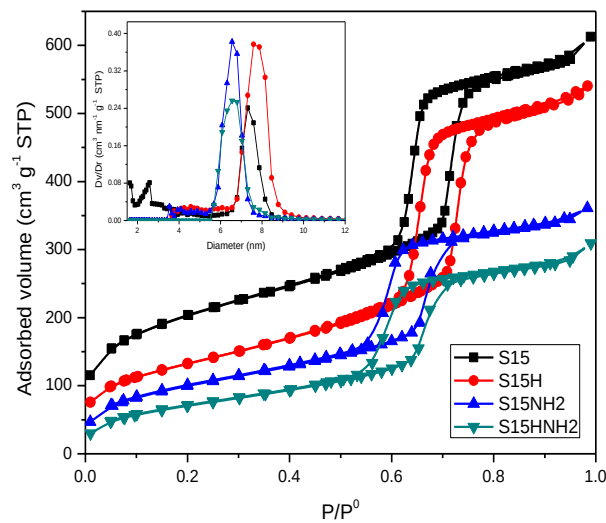


Figure 9. N₂ sorption isotherms at 76 K and PSD curves (Fig. inset) of SBA-15 substrates: S15, S15H, S15NH₂, and S15HNNH₂.

The main textural parameters of the different substrates were also calculated through the same N₂ sorption analysis. The changes of the structural properties underwent by the S15 precursory sample can be due to both hydroxylation with NaOH and subsequent functionalization with APTES; these parameters are listed in **Table 1**). The temperature of 76 K corresponds to the approximate boiling point of N₂ at Mexico City’s altitude conditions.

Table 1. Textural properties obtained from N ₂ sorption at 76 K					
Sample	A _{BET} (m ² g ⁻¹)	V _t (cm ³ g ⁻¹)	A _{ext} (m ² g ⁻¹)	A _{mic} (m ² g ⁻¹)	D _{mode} (nm)
S15	696.4	0.899	456.4	240.0	7.30
S15H	466.9	0.818	439.7	27.2	7.59
S15NH ₂	360.9	0.534	335.8	25.1	6.56
S15HNNH ₂	260.9	0.441	260.9	0.0	6.56

A_{BET} is the BET Surface area, V_t is the total adsorbed volume, A_{ext} is the external surface area, A_{mic} is the virtual micropore surface area and D_{mode} is the modal pore diameter arising from the NLDFT approach employing a kernel that employs the boundary desorption isotherm relative to N₂ at 76 K in cylindrical pores.

The above results confirm the beneficial effect of subjecting the surface of the substrates to a previous hydroxylation treatment that produces the elimination of micropores (i.e. the loss of A_{mic}), and concomitantly, an increase of the mean pore size favoured by the deposition of APTES on the surface of mesopores, while involving no pore mouth blocking. This happens since when SiO₂ substrates are exposed to a hydroxylation process, some silica lixiviation occurs, the intensity of which depends on the concentration of the caustic solution; this process being known as alkaline desilication. It is also important to remember that the hydroxylation operation was performed at 313 K.

Now, when comparing the textural parameters listed in Table 1 and associated to the pioneering materials and their corresponding solids obtained after APTES functionalization of the surface, the next important effects are as follows:

- 1) The presence of APTES led (sample S15HNH2), with respect to the precursory unhydroxylated substrate (S15), to the reduction of the mean pore width by 10.1%; likewise and a concomitant decrease of 40.6% of the total pore volume. In turn, the S15H adsorbent showed a pore size diminution of 13.2% and a reduction of 44.1% of the total pore volume. These facts being evidence that more APTES molecules are anchored on the hydroxylated specimen.
- 2) The APTES molecules fixed on the surface of the APTES seem to be uniformly dispersed on the mesopore surface. However, the disappearance of micropores from the S15 solid indicates that the access to these voids is somewhat precluded by the APTES species, perhaps from the mere pore entrances. The surface area was decreased by 10.4% with respect to their original magnitude. On the other hand, the S15H material depicted no micropores from the start, thus suggesting that practically all APTES molecules reside in the interior of mesopores.

From the above statements, it is evident that the S15H adsorbent becomes the most efficient toward APTES capture if compared to the S15 parent specimen. The origin of this phenomenon may reside on the generation of higher number of free silanol groups (Si-OH) on the silica surface by action of the NaOH alkaline treatment.

5.7 NH₃ Thermal Programmed Desorption from silica specimens (TPD-NH₃)

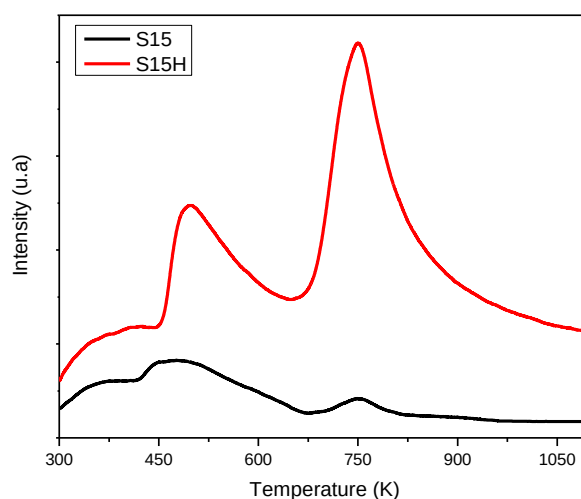


Figure 10. TPD-NH₃ analysis of the S15 precursory solid as well as for the S15H hydroxylated material following thermo-alkaline treatment.

The amount of acid sites on the surface of the different adsorbents was determined by the amount of NH₃ desorbed from the surface in the temperature range of 298–923 K. The results are schematically shown in **Fig. 10**. In this figure, the silica solids depict two types of sites: (i) weak sites are evident by the band appearing from 443 and 673 K, and (ii) sites considered strong appear

between 743 and 923 K. This implies the presence of Si-OH centers on the surface of the silica adsorbents, which are responsible of NH₃ adsorption. The tendency found with respect to the abundance of acid centers, in our samples, is as follows: S15H >> S15. This result confirms that the hydroxylation treatment employed induces the formation of a greater number of acid sites, a fact that is reflected by the area displayed under each of the TPD peaks.

5.8 Nuclear magnetic resonance (NMR) studies

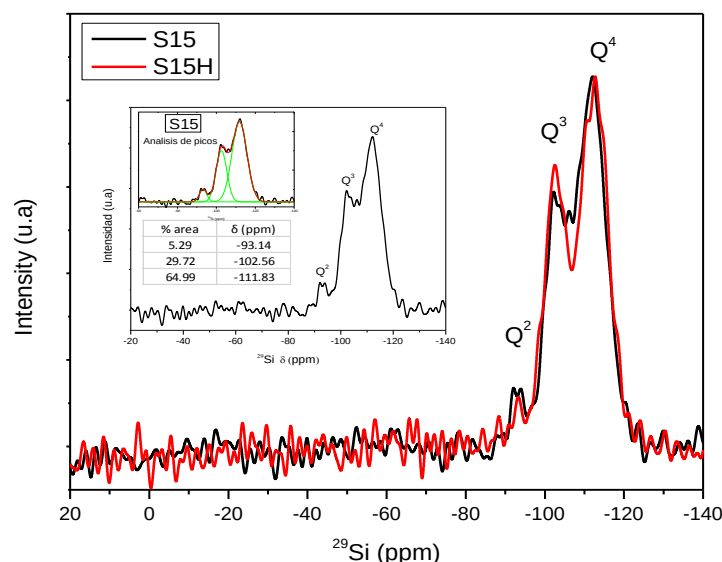


Figure 11. ²⁹Si NMR spectra (HPDEC) for the S15 and S15H samples. The figure inset exemplifies a peak analysis performed after deconvolution of the set of different signals obtained through ²⁹Si NMR analysis on the S15 solid.

Fig. 11 shows an assessment of the High Power Decoupling (HPDEC) NMR spectra corresponding to the precursory, still unfunctionalized SBA-15 silica materials. On the other hand, **Fig. 12** depicts a comparison between the RMN spectra of the S15H hydroxylated substrate before and after being chemically functionalized with APTES. Finally, the NMR spectra displayed in **Fig. 13** shows a comparison of the signals emitted by the S15 (unhydroxylated) and S15H (hydroxylated) solids after being modified on their surface by the anchoring of APTES. All these NMR spectra sets of signals appearing at chemical displacements of -112, -103, and -93 ppm and are due to the silica structure, and denoted as Q⁴, Q³, and Q², respectively. The Q², Q³, and Q⁴ bands are referred to silicon atoms bonded to two, three or four *siloxane* chains (–O–Si–) and the respective *silanol* groups (Si–OH) in order to complete four unions, respectively [33, 34].

The ²⁹Si NMR spectra (HPDEC) of the S15 and S15H samples, **Fig. 11** reveals similar population of Q⁴ and Q² type silicon sites and an increased one of the Q³ type for the hydroxylated sample. This situation suggests that the pristine SBA-15 network maintains most of its silicon atoms totally integrated (condensed) into the network (Q⁴) while still possessing silicon atoms having one or two hydroxyl groups, but in lower amounts (Q³ and Q²); Furthermore, the above evidence suggests that hydroxylation promotes the formation of silicon atoms depicting a solitary hydroxyl group (Q³), as reveals the spectrum of sample S15H of **Fig. 11**.

The signals centered at around 69 and -59 ppm observed in the spectra of **Figs. 12** and **13** are associated to silicon atoms bonded to organic chains (Si-C) and labeled as T³ and T² signals, which are equally distributed [33, 34] and merge as consequence of the establishing of the chemical bond established between the free silanol groups and the APTES functionalizing molecule. These bands can be attributed to the existence of silicon atoms bonded to carbon atoms and either two (T₂) or three (T₃) siloxane chains (Si-O-Si); additional surface hydroxyl groups that are required for completing the four unions of each silicon atom [33, 34]. Therefore, therein it is evident the presence of T³ and T² signals for the S15HNH2 solid, thus confirming the efficient anchoring of APTES on the surface of the precursory S15H material.

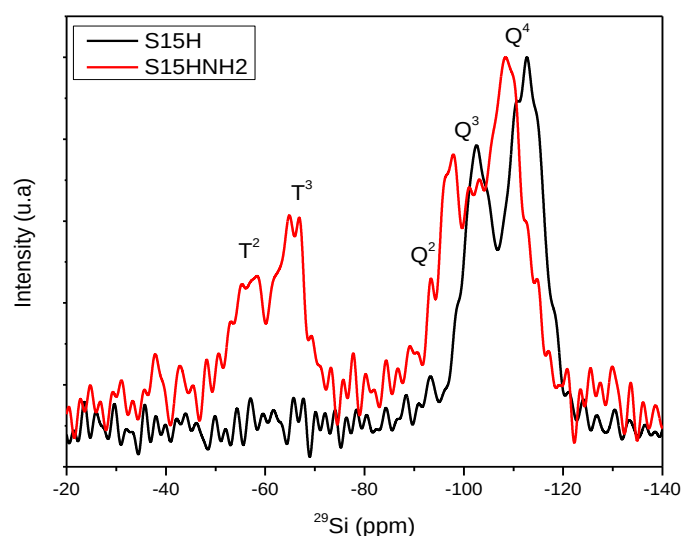


Figure 12. ²⁹Si NMR spectra (HPDEC) for the S15H precursory and APTES (S15HNH2) modified samples.

Yet, in the case of the ²⁹Si NMR spectra in Fig. 12 there can be also observed a slight chemical shift of the Q⁴ and Q³ bands for the case of the APTES functionalized sample (S15HNH2), with respect to the signals of the S15H hydroxylated solid (S15H). This could be due to a kind of screening effect because of the presence of APTES in the functionalized material, something that modifies the chemical ambient around the silicon atoms causing some shrinkage of the starting silica matrix. The emergence of a signal between the Q² and Q³ bands in the spectrum of the S15HNH2 can be interpreted as consequence of the surface attaching of APTES, preferably trough one or two unions established over the silicon atoms having one hydroxyl group (Q³). This assumption is confirmed by the apparition the T² and T³ groups. In other words, the attachment occurs by the substitution of one or two of the ethoxy groups (C₂H₅O-) of the APTES molecule with similar number of surface hydroxyl group (Si-OH) of the mesoporous substrate. Then, the slight increase of the Q² population can be due to the integration of new silicon atoms of the APTES molecule with two remnant hydroxyl groups as consequence of the hydrolysis of their ethoxy groups. In order to avoid confusion, we have to remember that it is a well-known fact that the alkoxide groups of APTES or of any silicon alkoxide are readily substituted by the hydroxyl groups arising from water. However, in the absence of this substance, the only way for reaching condensation consists in the substitution of

alkoxide groups by oxygen atoms proceeding from hydroxyl groups with the subsequent formation of the respective alcohol.

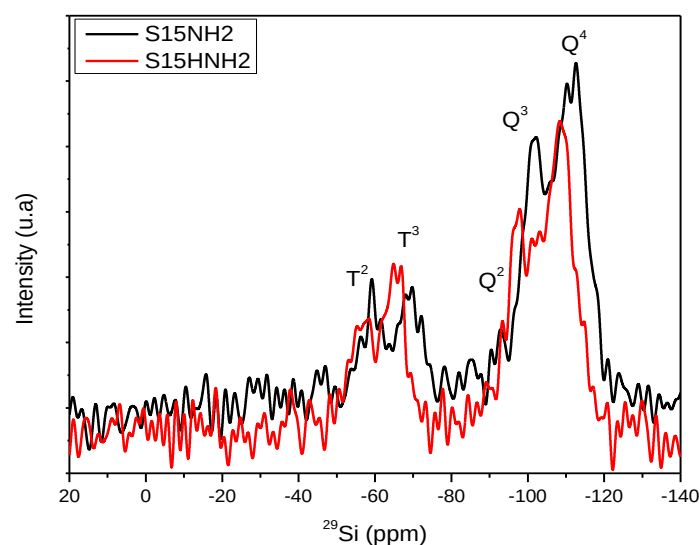


Figure 13. ^{29}Si NMR (HPDEC) spectra of solids modified on the surface via APTES: S15NH₂ and S15H₂NH₂ substrates.

Fig. 13 also depicts the same shifting effect, i.e. for the S15H₂NH₂ substrate a chemical shift of about 4 ppm is evident with respect to the signals emitted by the S15NH₂ solid [33, 34]. Therefore, the S15H₂NH₂ silica matrix provides emissions toward higher chemical displacements; this can be due to the change in the chemical environment around the silicon atoms or a contraction of the porous matrix caused by the APTES species existing on the pore walls of the previously hydroxylated specimen; this can be confirmed after N₂ adsorption characterization.

This event is not surprising given that chemical displacements are very sensible to the chemical environment, and these conditions can be different even for the same molecule on the same support, depending on the amount of these chemisorbed species. In the same fashion, it is also evident in Fig. 13 that the S15H₂NH₂ solid possesses a larger amount of T groups, if compared to those present in the S15NH₂ specimen. This effect is also reflected on the relative decrease of the amount of available Q³ groups and due to Si-OH groups that have been employed to bind the APTES molecules to the surface. However, in this S15NH₂ sample the T² band displays a higher intensity than the T³ signal.

The intensity differences found between the S15NH₂ and S15H₂NH₂ samples confirm that a condensation reaction occurred between the hydroxyl groups (Q² and Q³) on the respective surface of the substrate (S15 or S15H) and two or one ethoxy groups of APTES. For this reason, on the S15NH₂ sample the most intense T² band (having one hydroxyl group) and the increased Q³ band can only be associated to the presence of the silicon atoms of APTES. Additionally the higher intensity of the T³ band in the S15H₂NH₂ sample suggest a higher integration or condensation of the silicon atom of APTES as consequence of higher amount of vicinal silanol groups over the surface of the S15H substrate used.

Finally, in Table 2 the areal percentages associated to the Q and T NMR peaks are listed, once an appropriate deconvolution of these signal patterns has been performed (just as it was exemplified in Fig. 11 for the case of the S15 sample).

Table 2. Areal percent of signals proceeding from NMR spectra obtained from the deconvolution of the respective signals of precursory and amine-functionalized substrates.

Material	% Q ² Area	% Q ³ Area	% Q ⁴ Area	% T ² Area	% T ³ Area
S15	5.29	29.72	64.99	-	-
S15H	2.87	35.28	61.85	-	-
S15NH ₂	0.16	32.84	48.75	7.91	10.34
S15HNH ₂	5.30	21.25	45.00	13.54	14.90

The results listed in Table 2 confirm the statements induced from the aspects of Figs. 11–13. In this way, the argument advanced from the NMR characteristics of the S15 and S15H materials corroborates that these materials are appropriate for the anchoring of different amounts of APTES on their surface. Even if every synthesis is carried out, under the same conditions and employing the same amount of APTES, the S15H specimen adsorbs a larger amount of this compound. The net effect of hydrolyzing (S15H) is reflected in a larger proportion of Q³ species, i.e. free or adjacent silanol groups. There exists a difference of 15.6 % in the amount of Q³ species between the S15 and S15H substrates. Besides, it can also be observed that the hydroxylated matrix appears to be less condensed than the matrix that was not subjected to this treatment; at this respect, just check the Q⁴ ratios against each other for the two materials.

In the same way, when comparing the functionalized solids, the largest amounts of T² and T³ signals were found in the S15HNH₂ sample rather than in the S15NH₂ solid, something that can only be associated to the attached APTES species. Once the respective areal ratios are observed and considering the molar ratios chosen for carrying out the chemical anchoring on silica, as described in the Experimental Section, it is found (in a semiquantitative way) that 1.85 mmol and 3.3 mmol of APTES have been anchored on the surface of the S15NH₂ and S15HNH₂ samples, respectively. This difference strongly confirms that the soft alkaline hydroxylation improves as much as 43.8% the anchoring efficiency of APTES towards the SiO₂ surface (Table 2).

In brief, the hydroxylation alkaline process generates a higher efficiency toward the anchoring of APTES molecules than the unhydroxylated pioneering specimen does. This is mostly due to the larger amount of silanol groups on the pore walls of the hydroxylated samples. Under the presence of these species, the anchoring of the functionalizing amine species occurs more efficiently under molar ratios of 2:1 and 1:1. In contrast the unfunctionalized S15 precursory specimen is only able to anchor APTES molecules under a 2:1 Si-OH: APTES molar ratio.

5.9 CO₂ adsorption studies and determination of the isosteric enthalpy of adsorption

CO₂ adsorption isotherms on the assorted four silica substrates were determined at 273 K, 283 K, 293 or 298 K for each one of them (Figs. 14–16). When the isotherms of the unfunctionalized substrates (i.e. S15 and S15H) were analysed, it was observed that the hydroxylation treatment endowed these adsorbents with a higher CO₂ adsorption capacity (Fig. 14). Nevertheless, the

values of the isosteric enthalpy of adsorption evidence a relatively weak interaction (since these data are rather lower than the enthalpy of sublimation of CO₂) between the adsorbent surface and the CO₂ molecules. The possible reason of this behaviour can be due to the intrinsic acidic nature of the silica surface [35]; there is no a drastic difference in the isosteric enthalpies depicted by the S15 and S15H silica samples.

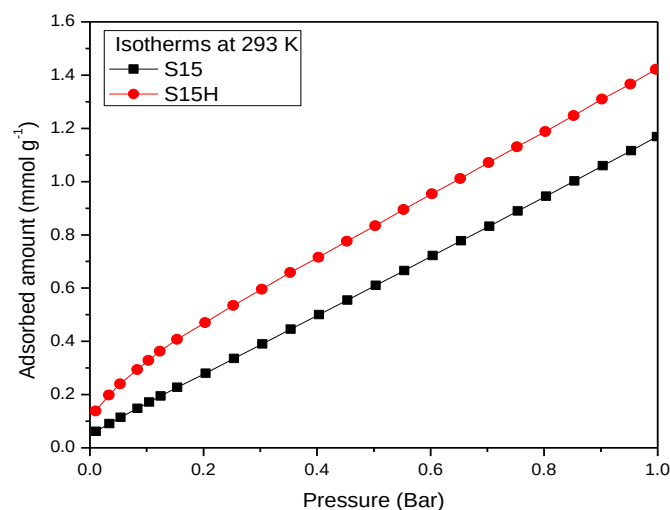


Figure 14. CO₂ adsorption isotherms at 293 K, on SBA-15 precursory samples and the hydroxylated (S15 and S15H).

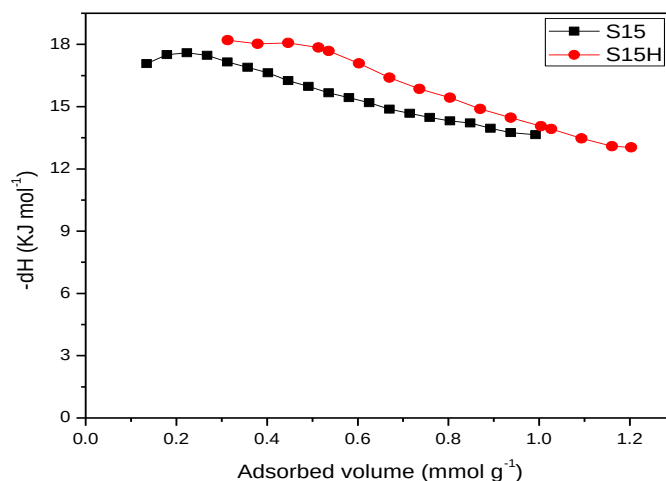


Figure 15. CO₂ adsorption enthalpies of SBA-15 precursory and hydroxylated silicas (S15 and S15H).

On the other hand, the CO₂ adsorption isotherms corresponding to the APTES functionalized substrates, confirm that the hydroxylated specimens better contribute to a more efficient and energetic CO₂ capture at different temperatures (Figs. 15 and 16). Also, the affinity toward CO₂ of amino-functionalized silica substrates (hydroxylated or not) is better than that depicted by the unfunctionalized solids.

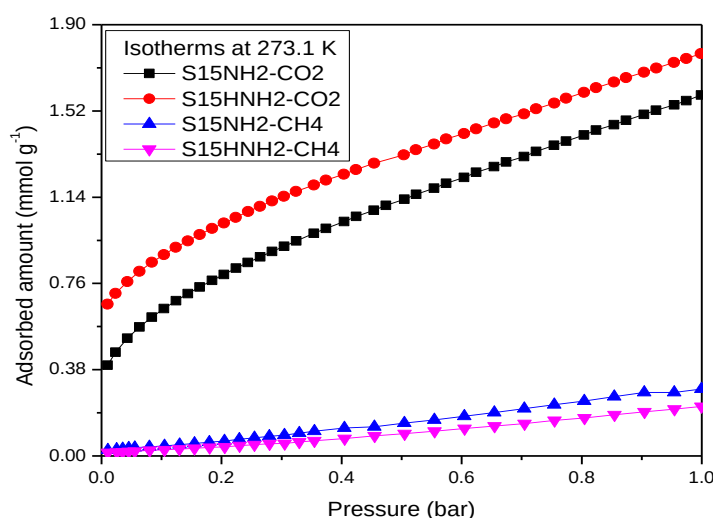


Figure 16. CO₂ and CH₄ adsorption isotherms at 273 K on APTES functionalised materials SBA-15 (pristine S15 and hydroxylated silica S15H).

Fig. 16 displays a comparison between CO₂ and CH₄ adsorption at 273 K on the precursory and hydroxylated amino-functionalized silica materials. Therein, it can be appreciated that the aforementioned substrates present a strong selectivity toward CO₂ rather than to CH₄. This could be due to the fact that the methane molecule is non-polar and it is not establishing a meaningful attractive interaction with the amine-functionalized SiO₂ surface. The adsorbed volumes showed by the silica samples toward CH₄ are too low for being reflected in a significant CH₄ uptake. These contrasting adsorption characteristics between CO₂ and CH₄ will indeed be very useful (i.e. selective) for achieving a fine separation of mixtures including these gases. Likewise, note the better affinity at 298 K towards CO₂ depicted by S15HNNH2 results to be 19.2% higher than the amount adsorbed by the S15NH2 matrix at 1 bar. However, even at 298 K and 0.15 bar, there is a difference between the two samples of 36.2% (**Fig. 17**). Even at 0.15 bar, S15HNNH2 shows a better performance at 313 K than the S15NH2 specimen, thus stressing this way that the thermos-alkaline treatment, previous to APTES surface functionalization, is very convenient in order to significantly raise CO₂ capture capacity with respect to the amount adsorbed by the APTES functionalized but not previously hydroxylated S15NH2 substrate. Therefore, the S15HNNH2 solid is more recommendable than the conventional (i.e. not hydroxylated) samples for pursuing enhanced CO₂ adsorption this effect being valid at 0.15 bar and temperature range of 318–348 K [36]. It is also important to mention that the CO₂ isotherms shown in **Fig. 17** depict low-pressure hysteresis loops at 298 K. However, this phenomenon is more accentuated for the S15HNNH2 specimen than for the S15NH2 material. This suggests that there may exist a stronger interaction between CO₂ and the hydroxylated surface (previously of being functionalized with APTES) than with the analogous but unfunctionalized sample [37].

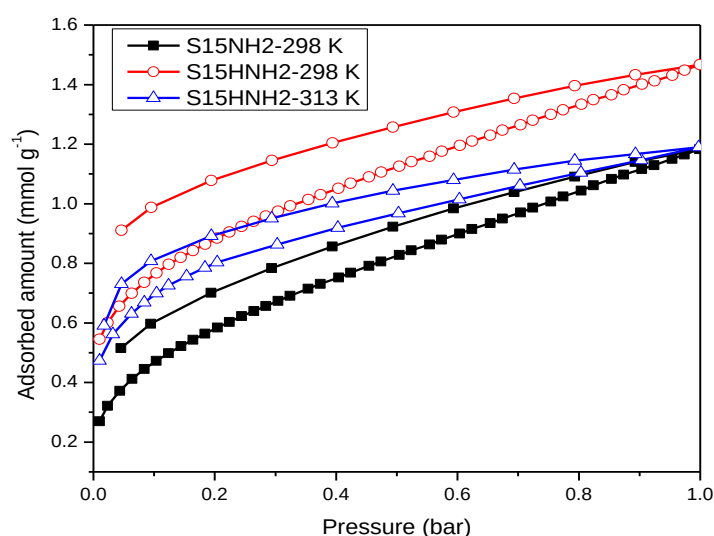


Figure 17. CO₂ adsorption isotherms at 298 K on APTES functionalised materials SBA-15 (pristine and previously treated SBA-15) and isotherm at 313 K for the sample SBA-15 hydroxylated and functionalized with APTES.

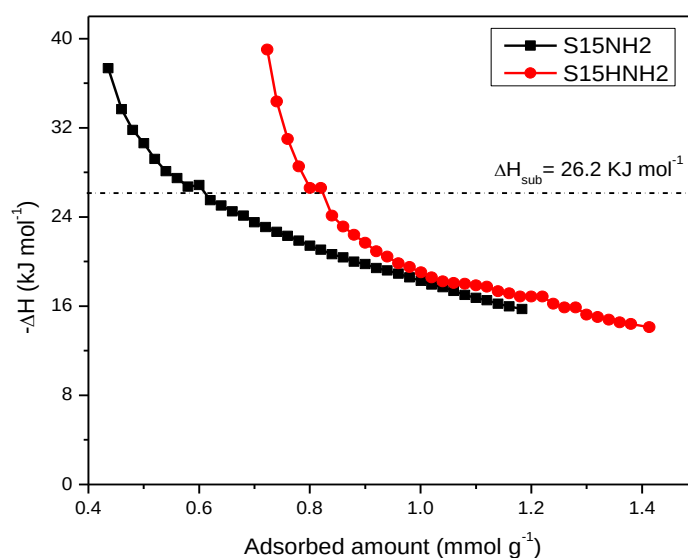


Figure 18. CO₂ enthalpy of adsorption of the APTES functionalized SBA-15 substrates ((pristine and previously treated SBA-15).

Finally, in order to know the forces that dominate the adsorbate-adsorbent interactions, a study of the isosteric enthalpy of adsorption was performed in relation to CO₂ adsorption at different temperatures; this study was carried out for the solids modified on the surface with APTES. The adsorption runs were made at the following temperatures: 263 K, 273 K, 288 K and 298 K; the determination of the isosteric enthalpy of adsorption was calculated from a Clausius-Clapeyron type equation (which simulated the quasi-solidification of CO₂ on the silica surface) **Fig. 18**. The results shown with respect to the enthalpy of adsorption of CO₂, for the APTES functionalized substrates,

emphasize that in these materials prevails an adsorption potential that strengthen the adsorbate-adsorbent forces, in comparison to the process carried out on the unfunctionalized silica solid. Then, the thermo-alkaline pretreatment and its subsequent functionalization with APTES lead to a higher capacity for CO₂ adsorption in comparison to that achieved by the amine-functionalized silica precursory specimen. The absolute $-\Delta H$ values result higher than the CO₂ enthalpy of sublimation (ΔH_{sub}) at 298 K and 1 bar (i.e. 26.2 kJ mol⁻¹). These results constitute an evidence of that a larger amount of CO₂ is being anchored on the surface of S15HNH2 solid than on that of the S15NH2 adsorbent. Even more, the S15HNH2 adsorbent establishes a stronger interaction with CO₂ than the unhydroxylated specimen extended up to at large surface coverages. For instance, for a CO₂ surface coverage of 0.72 mmol g⁻¹, an enthalpy of adsorption of -23.2 kJ mol⁻¹ is found for the S15NH2 specimen in contrast to the -39.2 kJ mol⁻¹ evolved by the S15HNH2 adsorbent.

In addition, chemical reason for this adsorption behavior can reside on the evidence provided by the NMR spectra; the ideal case would be that three Si-OH silanol groups be required to substitute the three alkoxi groups of the APTES molecule. However, this occurs, in such a way that the remaining Si-OH species still be available for establishing more bondings with more alkoxi groups proceeding from other APTES molecules. That is, the attaching of one APTES molecule render new Si-OH groups for the attachment of other APTES molecules, see Fig. 19. The above behaviour can be attributed to the chemical nature of the adsorbent-adsorbate pair; there arise two possible chemical explanations for these phenomena to occur:

- a) The basic property of amino functionalized silica acts as a Lewis base that provides an electron pair to the CO₂ molecule, i.e. a coordination link is formed. The adsorbed CO₂ molecule then establishes additional coordination links with the hydrogen or oxygen atoms proceeding from adjacent APTES or with the hydroxyl groups of the silica substrate (see Fig. 19).

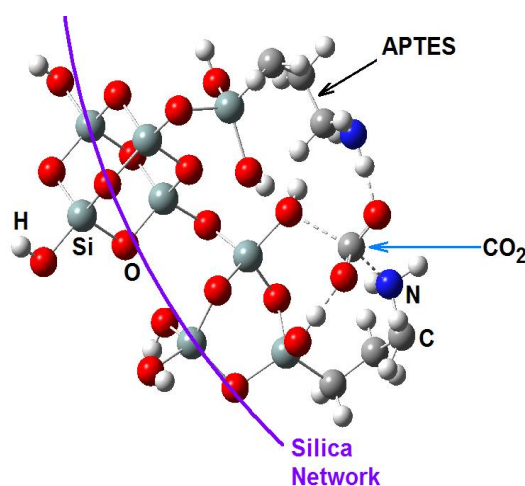


Figure 19. Hypothetical situation of the CO₂ molecule interacting with the nitrogens and hydroxyl groups of the APTES modified surface of SBA-15 substrate.

- b) When the surface population of Si-OH species is increased via the hydroxylation treatment, the APTES molecules exist over larger domains of the pore structure, therefore a larger amount of CO₂ molecules get captured. This last process could be carried out by the zwitterionic reaction that renders ammonium carbonate after the reaction of amine and CO₂

species [38]. The capture of CO₂ by neighboring (i.e. more concentrated) amine species results to be more stable than if only an individual, isolated amine species reacts with CO₂. A proton of one of the neighboring amines is transferred to the adjacent one; the former NH₂ group attracts the carbon dioxide to it due to its refurbished electron density. This being a reflection of this assertion, since therein a considerable enthalpy of adsorption is attained by the amine-functionalized and pre-hydroxylated SBA-15 species Fig 17.

6. Conclusions

These reports show that the soft alkaline hydroxylation is suitable for raise the species Si-OH in the conventional silica, which could be due breaking of Si-O-Si in groups Si-OH for the hydroxylation treatment, at the same time this involve loss of micropores and uniformed increase of mesoporous. The alkaline treatment, in the case of the S15 substrate, caused a pore width increment of 2.9 Å, it phenomenon is appropriate for our purposes. In all case, those new Si-OH groups allow the attachment of APTES in the present study but can helps to fixing other active chemical species over mesoporous substrates.

In the other hand, the studies of the amount of acid sites on the surface of the different adsorbents determined by the TPD-NH₃ desorbed and NRN confirm that the hydroxylation treatment employed induces the formation of a greater number of acid sites. In this way, the argument advanced from the NMR characteristics of the S15 and S15H materials are appropriate to anchor different amounts of APTES on their surface, even if every synthesis was carried out under the same conditions and employing the same APTES: solid molar ratio. Is likely due to the abundant presence of Q³ species, i.e. free or adjacent *silanol* groups, in spite of the fact that there existed a difference of 15.6 % in the amount of these species on each solid surface. Almost, the ²⁹Si NMR analysis reveals that is found in a semiquantitative way that 1.85 mmol and 3.3 mmol of APTES have been anchored on the surface of the S15NH₂ and S15HNH₂ samples, respectively. Furthermore, the fixing of APTES over these substrates preferably occurs through only one chemical union. Contrastingly, for the case of non-treated SBA-15 substrates, the APTES species remain fixed to the network through one or two chemical bonds.

Finally, the adsorption enthalpy of the APTES-functionalized substrates, confirms a higher CO₂ adsorption when substrates are previously treated with an alkaline solution. All the evidence herein shown suggests that interaction of CO₂ with the systems treated with alkaline solutions could be due to the existence of a higher amine group population or to the interaction combined with the existence of remnant S-OH groups coming from APTES.

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