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Keywords: Molecular Dynamics; Monte Carlo Simulation; Attosecond simulations; Subatomic Interactions; Subatomic Particle Simulations; Pilot-Wave Theory; Particle-Wave Nature



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Article

Subatomic Particle Simulations using Monte Carlo and Molecular Dynamics Algorithms to Simulate Stable Atom and Model Electronic Structures

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Abstract: A novel approach was presented in this study where molecular dynamics and Monte Carlo methods were applied to subatomic particles to simulate an atom using pseudo potentials. Pseudo potentials were developed for subatomic particles by conceptualizing them as conventional particles, exhibiting attractive and repulsive forces between them, ensuring the stability of an atom. A stable nucleus was formed at the center with electrons distributed around, resulting in the formation of an atom. Subatomic particle simulations impart a comprehensive perspective and a profound understanding of electron trajectories that correlates with atomic properties such as electron energies and atomic radius. These approaches intricately capture the impact of protons and neutrons motion in the nucleus on electron trajectories. Hydrogen and carbon atoms were considered, and their analyses were reported in this study. Time step for carbon atom simulation was calculated from dimensionless variables and found to be 1.67 attoseconds. The Pilot-wave theory was implemented to simulate the wave nature of subatomic particles in an atom. Electrons motion were guided by the interference pattern produced by the electron and proton aether medium waves. Molecular dynamics simulations on subatomic particles were implemented on an oxygen molecule, giving insights into electronic structures with electron trajectories shared by two atoms.

Keywords: attosecond simulations; subatomic interactions; subatomic particle dynamics; pilot-wave theory; particle-wave nature

1. Introduction

Coloumb's law plays an important role in modelling electrostatic interactions and forces acting on subatomic particles [1]. Coloumb interactions have been applied in many areas in the fields of Physics, Chemistry and Biology. This law was introduced in the 18th century and plays a pivotal role in understanding and modelling electrostatics in molecular simulations of atoms and molecules. They play a crucial role in modelling chemical reactions, formation of chemical bonds, material properties like electrical conductivity and magnetism, biomolecules like protein folding and DNA double-helix formation. Many studies presented the importance of coulomb interactions in modern applications in Carbon nanotubes [2] and Bilayer Graphene [3] systems. The standard perspective on protons and electrons is that protons bear a positive charge while electrons carry a negative charge and opposite charges attract while the same charges repel each other. This study provides a new perspective on looking into subatomic particles. Protons, electrons, and neutrons are treated as conventional particles in an atom, exhibiting attraction and repulsion between them, akin to the way atoms attract and repel each other in atomic simulations. This new approach leads to developing pseudo potentials for subatomic particles to model the subatomic particle interactions in an atom.

Monte Carlo (MC) simulations are a powerful tool in understanding molecular systems with their ability to sample from a large configurational space. MC simulations is based on generating random configurations for a system and accepting or rejecting them based on probability rules such as Metropolis criterion. In computational chemistry and physics for studying phase transitions, adsorption isotherms, and molecular conformations, Metropolis et al. laid the foundational work in

1953.[4] With advancements in computing power, more complex systems were simulated providing insights into phenomena at molecular scale. Gibbs ensemble Monte Carlo method developed by Panagiotopoulos in 1980, for phase coexistence systems by creating two or more boxes representing different phases gives insight for vapor-liquid equilibria, critical points and phase transitions.[5] Configurational-bias Monte Carlo method used for sampling long-chain molecules that improves the efficiency of exploring the conformational space.[6] In this study, Monte Carlo simulations were implemented to subatomic particles using pseudo potentials to simulate a carbon atom and properties like electron's energies and atom radius were analyzed.

Molecular Dynamics (MD) provides a fundamental approach to the dynamics of atoms and molecules at the atomic scale. This method was first introduced by Alder and Wainwright in 1959 [7] and evolved into an important tool in understanding the dynamics of atoms and molecules in complex systems. Its applications have diversified, ranging from biomolecular simulations [8] to materials science [9]. The introduction of enhanced sampling techniques by A. Laio and M. Parrinello [10] gives better insights into complex energy systems. In this study, molecular dynamics simulations were developed to model the subatomic particles interaction in within-atom. Results of these simulations for the hydrogen atom, carbon atom, hydrogen molecule and oxygen molecule were reported in this study. The time step for these simulations was calculated from dimensionless variables and found to be in the range of attoseconds. Many attosecond studies [11–15] have been carried out that provide insights into electron dynamics, ionization, and quantum mechanical phenomena occurring within atoms and molecules.

The Pilot-Wave theory, also known as de Broglie-Bohm theory [16,17] is an interpretation of quantum mechanics with a deterministic framework where particles are guided by guiding waves with hidden variables. Louis de Broglie [16] and David Bohm [17] developed this theory, giving insights into wave-particle duality. Recent studies [18,19] explore the theory of light of contemporary quantum research. In this study, Pilot-Wave theory was implemented on Hydrogen and Carbon atoms, where subatomic particles movement was affected by the presence of three-dimensional waves that were generated by the particles in an aether medium. Many studies [20–24] were carried out for Pilot Wave walkers where an oil droplet is allowed to bounce on the liquid surface and its motion is guided by the surface wave patterns like diffraction and interference.

In the remainder of this paper, simulation details were presented followed by results and discussion section. In the results and discussion section, Monte Carlo simulations for carbon atom followed by molecular dynamics simulations results of particle and wave nature were presented. Uranium atom, hydrogen molecule and oxygen molecule simulation videos were reported followed by conclusion section.

2. Simulation Details

Monte Carlo Simulations Details: Conventional canonical (*NVT*) ensemble Monte Carlo method were performed on subatomic particles to form a stable atom assuming volume and temperature of atom is constant. MC simulation code was developed in Fortran. Carbon atom was considered with 18 subatomic particles (6 electrons, 6 protons and 6 neutrons), placed randomly in the simulation box of dimensionless length (L^*) of 120. Atom formation was observed in MC simulation with properties of atom like energies and radius were analyzed. Potential between subatomic particles were calculated based on equations 4, 5 and 6. Potentials of proton to electron, electron to electron, proton to proton, proton to neutron and neutron to neutron were considered in this simulation. No boundary conditions were applied in this simulation. No cut-off distance was used in calculating potentials between subatomic particles. Acceptance ratio of $e^{-(E_2-E_1)/C}$ was used in the MC simulation. Where C is the simulation tuning parameter that relates to temperature of the atom. E_2 and E_1 are the energies of atom for trail move and original configuration. Regular trail moves of electrons, placing randomly around the nucleus, gives poor sampling with low acceptance moves in the simulation.

To enhance the sampling, each trail move of electron consists of 10,000 additional samplings taken around the surface of the sphere with nucleus as center. All other particles were kept constant and the chosen particle was randomly sampled along the surface of the sphere around the nucleus as shown in Figure 1. Spherical coordinates were used making the 10,000 additional trail moves in each sampling. Electron trail move with dimensionless step length of 60 (L^*) and for protons and neutrons step length of 0.5 were considered. Nucleus center of mass was moved to origin at each accepted trail move. Both with and without center of mass adjustment gives same simulation performance and results. Atom formation and stabilization was observed in first 1000 MC cycles. Last 1000 MC cycles were considered for the analysis of energies and radius of atom.

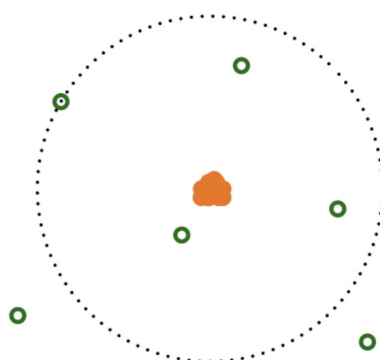


Figure 1. Electron's random sampling around the surface of sphere with nucleus as center for a Carbon atom. Dotted circle is surface of sphere (in 2D) with nucleus as center. Protons and neutrons are represented in solid orange circles, electrons are shown in green circles.

Molecular Dynamics Simulation Details: Simulation code was developed (in Fortran) for subatomic particles in an atom using molecular dynamics (MD) algorithm using Velocity Verlet integrator [25]. A dimensionless time step of 0.02 was used in MD simulations and the time step was calculated, from dimensionless variables, as 1.67 attoseconds. The initial configuration was created by placing protons and neutrons in the centre of the atom (nucleus) and the electrons placed randomly around the nucleus. Pseudo potentials were introduced in this study and are used to calculate force and potential acting on the subatomic particles at each time step. No boundary conditions were used for the simulation box and no cut-off was used for calculating the force acting on particles. All MD simulations were run for 16.7 picoseconds and results were collected for the last 0.2 picoseconds. All MD and MC simulation videos were captured using the VMD package [26].

Pseudo potentials were developed for subatomic particles on the principle of exhibiting repulsion (at smaller distances) and attraction (at larger distances) between them and between themselves. Considering this behavior for subatomic particles in MD code gives stable atom simulations. Pseudo potentials emerge by running MD simulations on subatomic particles by first considering the dimensionless variables for equations of force acting on subatomic particles. Two terms were considered for the force equation for subatomic particles similar to Leonard-Jones potential for molecular simulations. By trial-and-error method, these two terms were parameterized with the objective of keeping a stable atom and observing electron trajectories around the nucleus. Later, these force functions are integrated back to get the pseudo potentials for electrons and protons in an atom.

Following are the dimensionless force equations parameterized by using the trial-and-error method.

For electron to proton interaction,

$$f^* = 4 \left[\frac{1}{r^{*2.5}} - 8 \right] \tag{1}$$

for proton to proton, proton to neutron, neutron to neutron,

$$f^* = 4 \left[\frac{1.65}{r^{*0.4}} - 5 \right] \tag{2}$$

for electron to electron,

$$f^* = 4 \left[\frac{1}{r^{*0.2}} - 1.5 \right] \tag{3}$$

where f^* and r^* are dimensionless force and distance, respectively. More terms can be considered in these equations for complex trajectories of electrons around the nucleus. Integrating the above equations gives pseudo potentials for subatomic particles as follows:

for electron to proton:

$$V(r) = 4\epsilon \left[\frac{\sigma^{1.5}}{1.5r^{1.5}} + \frac{8r}{\sigma} \right] \tag{4}$$

for proton to proton, proton to neutron, neutron to neutron,

$$V(r) = 4\epsilon \left[\frac{5r}{\sigma} - \frac{2.75r^{0.6}}{\sigma^{0.6}} \right] \tag{5}$$

for electron to electron,

$$V(r) = 4\epsilon \left[\frac{1.5r}{\sigma} - \frac{r^{0.8}}{0.8\sigma^{0.8}} \right] \tag{6}$$

where $V(r)$ is the potential acting on particles, r is the distance between particles, ϵ and σ are pseudo potential parameters. ϵ and σ are parameterized using Carbon atom simulations, by its radius and Ionization Energy as shown in the following table.

Table 1. Pseudo potentials parameters, parameterized based on Carbon Atom Simulation

	ϵ (Joules)	σ (picometers)
Parameterized based on Carbon atom	7.837×10^{-23}	0.7777

The total potential energy of atom (force field potential function) was considered as sum of potential functions of electron to proton (eq. 4), electron to electron (eq. 6), proton to proton, proton to neutron and neutron to neutron (eq. 5).

3. Results and Discussion

3.1. Monte Carlo Simulations on Subatomic Particles in a Carbon Atom

Atom Formation was observed from randomly distributed particles in a simulation box as shown in Figure 2. All the protons and neutrons were came together and reached in the center of electrons distribution within 500 MC cycles. Simulation tuning parameter (C) was taken as 500 that shows reasonable acceptance moves in each MC cycle. Wide range of C values were considered and the energy of atom was found to be same in each case. Initial energy of the system is as high as 700×10^{-20} Joules and reached equilibrium by forming atom with average energy of atom at 290.7×10^{-20} Joules, as shown in Figure 3. This value of atom’s energy shows very good agreement with carbon atom’s energies of molecular dynamics simulation discussed in later sections. Radius of the fully formed atom was found to be 39.8 pm as shown in Figure 4. MC simulation of carbon atom formation and reaching equilibrium was shown in supplementary material video 1.

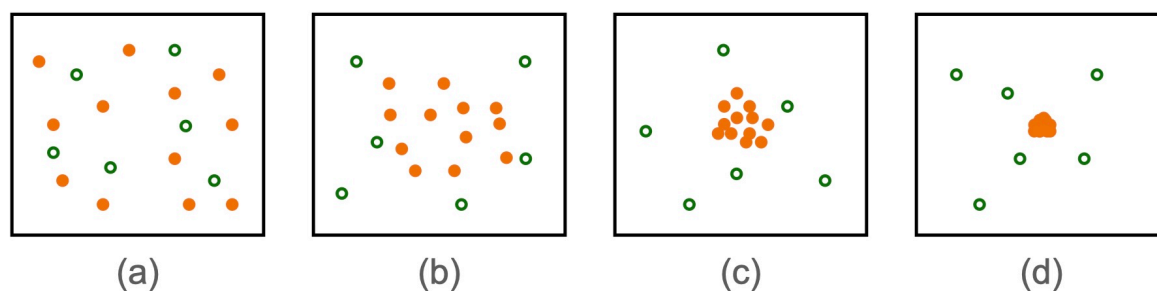


Figure 2. Carbon atom formation in Monte Carlo simulation. Protons and neutrons were represented in solid orange circles, electrons were shown in green circles.

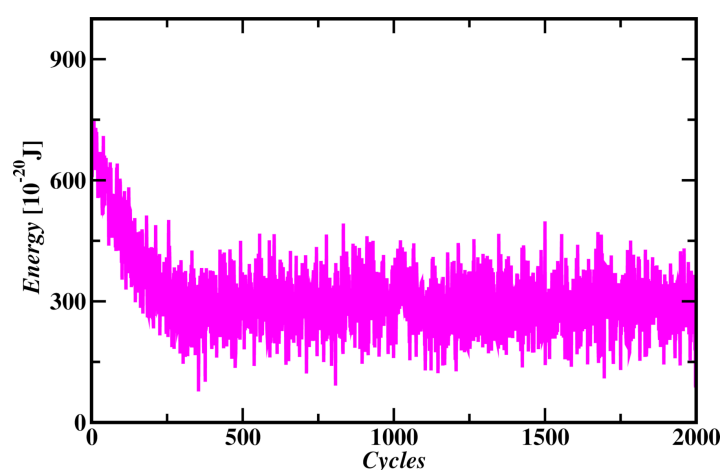


Figure 3. Energy changes in Monte Carlo simulation of a Carbon Atom for 2000 cycles.

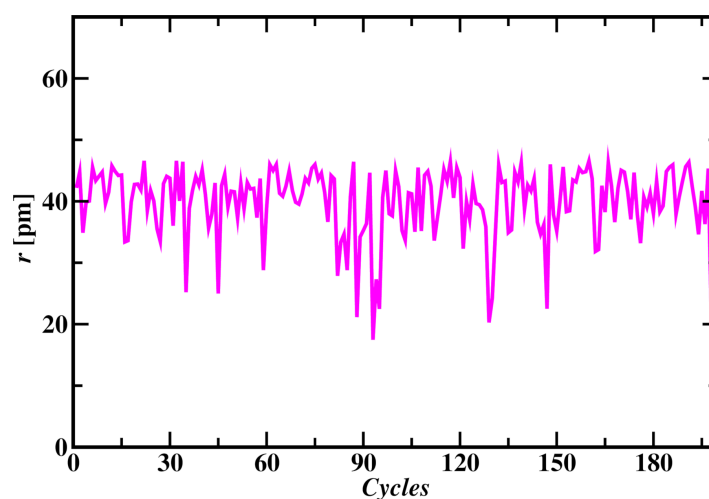


Figure 4. Radius of Carbon atom in MC simulation for the last 200 cycles.

3.2. Molecular Dynamics Simulations on Subatomic Particles in a Hydrogen and Carbon Atom

3.2.1. Hydrogen Atom Simulation

Subatomic particle dynamics simulation was carried out for the Hydrogen atom by placing the proton at the centre and placing an electron randomly around the proton. By the trial-and-error method, pseudo potentials were developed until the atom got stabilized and the electron rotates around the proton for an infinite time without losing its energy as shown in the supplementary material video 2. Electron revolves around the proton in three-dimensional elliptical and spherical orbits by changing its radius in cycles. The change in the radius of the Hydrogen atom is shown in Figure 5. In one cycle,

the radius of the Hydrogen atom changes from 0.22 pm to 1.33 pm. For every 34 time steps, this cycle repeats with the same pattern in a three-dimensional space.

The energy variation of electron while revolving around the nucleus was studied and shown in Figure 6. When the electron travels close to the nucleus, changes in potential and kinetic energy are observed. Potential energy curves vary with the radius of the atom and are directly proportional to the radius of the Hydrogen atom. Kinetic energy is inversely proportional to the radius of the atom. These profiles show that an electron near to proton has lower potential energy and higher kinetic energy and vice versa. Due to the attraction and repulsion behaviour of proton and electron interactions, the electron approaches the proton due to attraction and Potential energy decreases. Electron gets repelled when it goes very near to the proton and moves away with higher Kinetic energy. Potential energy decreases when the electron approaches the proton with an increase in Kinetic energy.

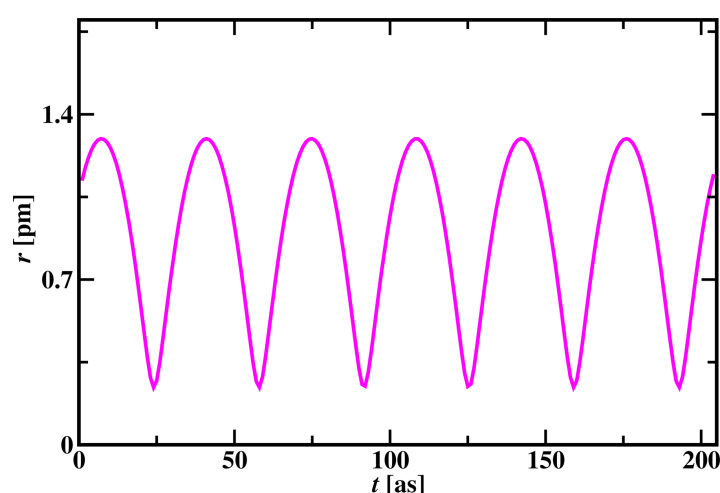


Figure 5. Radius of Hydrogen atom versus time in a subatomic particle dynamics simulation.

The average total energy of the electron in a Hydrogen atom is 436.05×10^{-23} J and is shown in Figure 7. Total energy decreases by 2.5×10^{-23} J when the electron approaches close to the nucleus. Total energy maintains constant through the revolution of an electron around the proton and a sharp increase in energy occurs when the radius is at minimum in each cycle. Electron's total energy reaches minima when it approaches the proton. There is a sharp increase in total energy of 4×10^{-23} J when the electron approaches r_{minimum} . After this sharp increase in energy, the electron starts to move away from the proton with the increase in total energy.

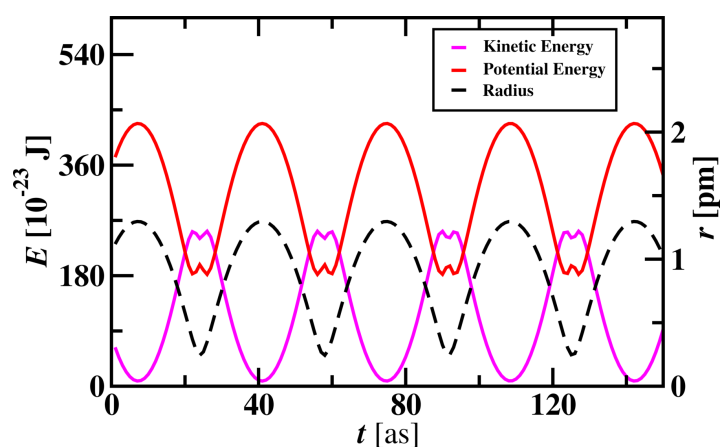


Figure 6. Energies of the electron and its radius variation with time in a Hydrogen atom. The radius of a Hydrogen atom is presented on the y-axis on the right.

The simulation of an electron orbiting around the nucleus was oscillating in one region of Phase Space configuration as shown in the supplementary material video 3. Figure 8 shows an example of an

electron phase space oscillating (represented by a black region) at the local region (local minima). To explore all the possible phase space configurations, the electron orbit has to be changed and has to be placed in other regions randomly around the nucleus. Properties of electron, proton and atom will be calculated when all the phase space configurations are captured in a simulation.

3.2.2. Randomness in Hydrogen Atom Simulation

To simulate all the possible phase space configurations for an electron revolving around the nucleus, the electron is placed randomly around the nucleus, with the probability, at each time step. x , y , and z coordinates of the electron changed with the probability, individually. An electron is placed randomly around the nucleus within the radius of 1.01 pm in x , y , and z coordinates. Changing the velocity of an electron randomly, by a very small value, shoots the electron as those random velocity values are not reasonable. Only random positions were given to the electron keeping its velocity same. Increasing the probability of placing an electron randomly around the nucleus and keeping the velocities the same, results in the higher radius of the Hydrogen atom.

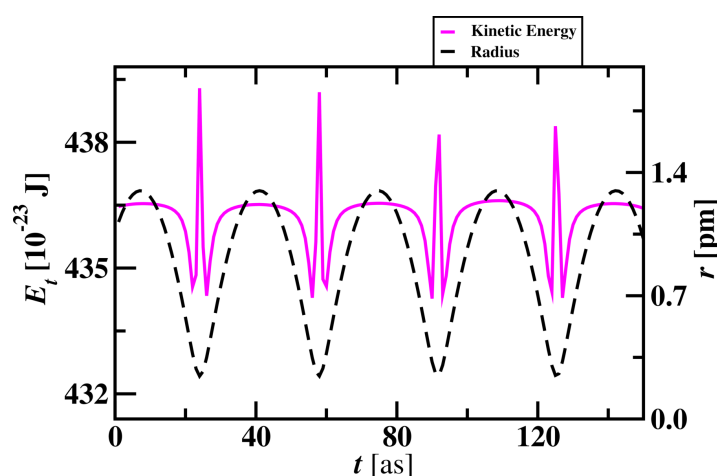


Figure 7. Total energy of the electron and its radius variation with time in a Hydrogen atom. The radius of a Hydrogen atom is presented on the y-axis on the right.

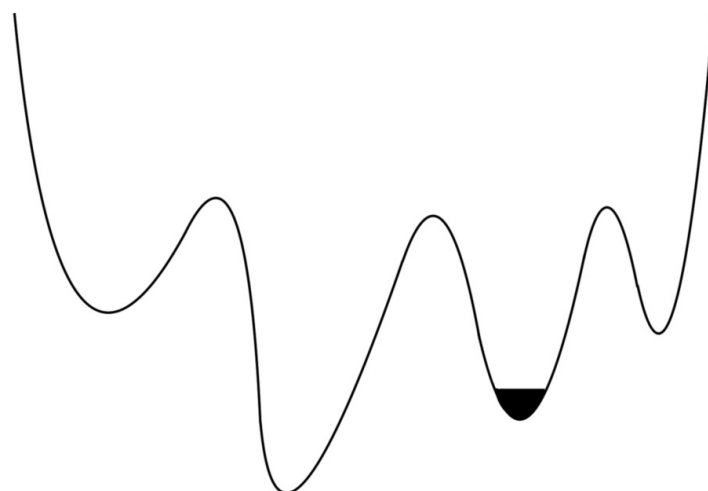


Figure 8. Electron's phase space configuration with local minima. The electron trajectory oscillates in one of the local minima (black region).

The simulation video of Hydrogen atom and electrons placed randomly in the simulation is shown in the supplementary material video 4. Randomness in the Hydrogen atom gives electron revolutions around the proton in all directions, unlike in normal Hydrogen atom simulation (supplementary material video 2) where the electron revolves only in one direction. This property of electron's

revolutions in all directions explores maximum phase space configurations. The average radius of the atom, total energy, potential energy and kinetic energy of the electron and the probability of randomly changing electron position are shown in Table 2. An increase in the probability of placing an electron in random positions increases its radius and energies up to a probability of 0.00001. Further increase in probability decreases its values. The analysis shows the maximum radius and electron energies at the probability of 0.00001. The experimental value of the Hydrogen atom radius (53 pm) is reflected between the probability of 0.00001 and 0.0001. The ionization energy of the Hydrogen atom is calculated by the total energy of the electron and is 31323×10^{-23} J for the probability of 0.00001. This property is less than the observed value of 217896×10^{-23} J. This difference in ionization energy is due to pseudo potentials being parametrized based on Carbon atom properties (discussed later in this section).

3.2.3. Particle Nature of Carbon Atom Simulation

Molecular dynamics algorithm applied to subatomic particles for Carbon atom simulation by placing six electrons around the nucleus of six protons and six neutrons. The effect of proton and neutron motion in the nucleus towards electron trajectories was seen clearly as shown in the supplementary material video 5. The trajectory of electrons revolving around the nucleus for 5,000 time steps is shown in Figure 9.

Table 2. Average Radius, Total Energy, Potential Energy and Kinetic Energy for Normal simulation and Randomness simulation in a Hydrogen Atom

Probability of Randomly Placing Elec- tron	Average Ra- dius (pm)	Average Total Energy, 10^{-23} J	Average Potential En- ergy, 10^{-23} J	Average Ki- netic Energy, 10^{-23} J
0 (Normal Hydrogen Atom)	0.8875	436.97	314.58	121.4
0.000001	3.00	1453.32	974.18	479.14
0.00001	65.32	31323.31	21066.33	10256.99
0.0001	41.77	20374.48	13472.27	6902.20
0.001	14.00	6997.81	4520.54	2477.34
0.01	6.34	4418.42	2051.65	2366.77

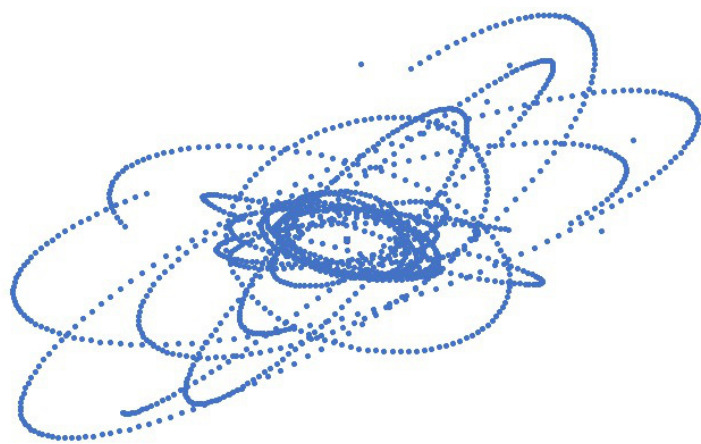


Figure 9. Electron trajectories around the nucleus of Carbon atom for 5000 time steps.

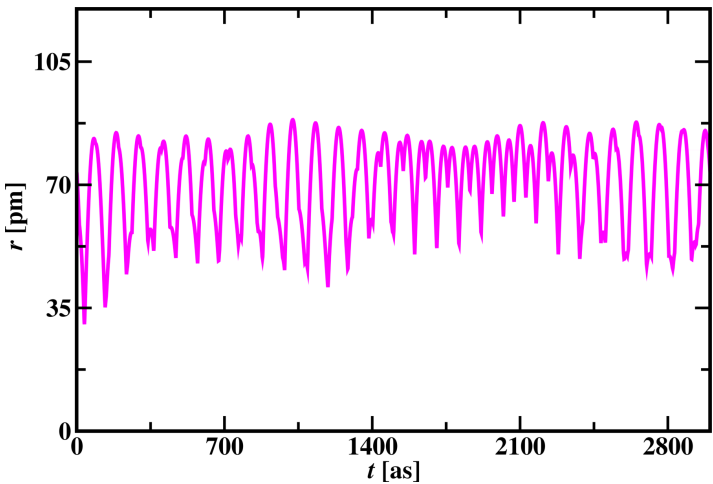


Figure 10. Change in Radius of the Carbon atom with time.

Unlike the hydrogen atom, the radius of the Carbon atom changes in an irregular pattern as shown in Figure 10. All the electrons orbit around the nucleus in cycles like the Hydrogen atom’s electron as shown in Figure S1 in the supplementary material. The valence electron is not constant and changes its position with core electrons and the core electron becomes a valence electron for a short period of time in a cycle.

The range of average velocity of electrons is found to be 835,747 m/s to 1,213,998 m/s in carbon atom, as shown in Table 3. The ratio of the velocity of electrons to the velocity of light is ranging from 0.0028 to 0.0040. The carbon atom’s average radius (r^*) is found to be 91.9 and σ in pseudo potentials is found to be 0.7777 pm. The ionization energy of the valence electron was calculated from the dimensionless total energy variable (23021.3) and ϵ in pseudo potentials were found to be 7.837×10^{-23} Joules. The dimensionless time step is taken as 0.02 in the simulation and time step was calculated to be 1.67 attosecond.

The average potential and kinetic energies of electrons (from core to valence) are shown in Figure 11. Core electron potential energies are lower when compared to valence electron, explaining core electrons are more stable than valence electron. Kinetic energy for core electrons is higher than valence electrons showing that core electrons orbit the nucleus with higher velocities than valence electrons. The total energy of electrons for core electrons is lower than for the valence electrons. The contribution of potential energy is higher than kinetic energy in total energy resulting in lower total energy for core electrons.

Table 3. The average velocity of electrons in carbon atom and compared with the velocity of light

	Average (m/s)	Velocity	Average ity/Light Velocity
Electron 1	835747.25		0.0028
Electron 2	989235.65		0.0033
Electron 3	977732.33		0.0033
Electron 4	1213998.64		0.0040
Electron 5	1058082.83		0.0035
Electron 6	980335.11		0.0033

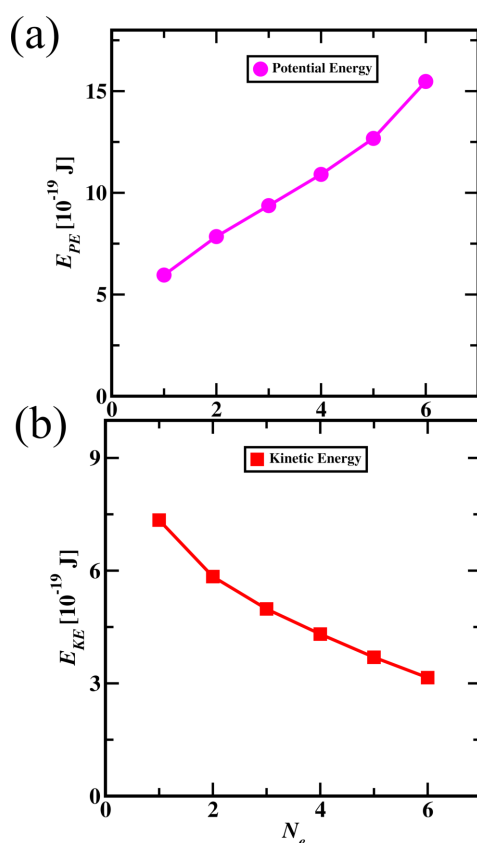


Figure 11. Potential (a) and Kinetic (b) energies of electrons in particle nature of Carbon atom. (On the x-axis, variable (N_e), the core electron numbered 1 and the valence electron numbered 6)

3.3. Wave Nature Subatomic Particle Dynamics Simulations

3.3.1. Wave Nature Simulation in Hydrogen Atom

Pilot-Wave theory was developed for subatomic particles to mimic the particle's wave nature in atom simulation. This approach is similar to walkers [20–24] in the two-dimensional surface wave where a silicon oil drop moves along the interference pattern of the wave produced by the medium. Here, particles produce three-dimensional waves in an aether medium called aether medium waves. Electrons are guided by the resulting interference pattern of aether medium waves produced by the presence of subatomic particles, as shown in Figure 12. Subatomic particles are assumed to produce three-dimensional waves on an aether medium, and their motion is guided by the resulting wave patterns.

In this study, these three-dimensional aether medium waves are considered as three-dimensional cos waves as a function of distance (r). The time-dependent cos wave function is not considered in this study and will be considered in future studies. This study assumes, the amplitude of the wave does not diminish over distance and propagates with constant wave amplitude, making the system nonlocal. The amplitude of aether medium cos waves is parameterized and reported here. Electron displacement was calculated as a result of the collision with aether medium waves. Proton movements are considered particle nature and the presence of proton's aether medium waves are considered. The collisions between aether medium waves and subatomic particles were considered as perfectly elastic where momentum and kinetic energy are conserved.

The velocity of an electron as guided by aether medium waves was derived as follows: The equation for the velocity of aether medium waves as a function of radius (r):

$$u_{Aether} = A \cdot \cos\left(\frac{r}{wl}\right) \quad (7)$$

where u_{Aether} is the velocity of aether medium waves, A is amplitude and wl is the wavelength of the waves.

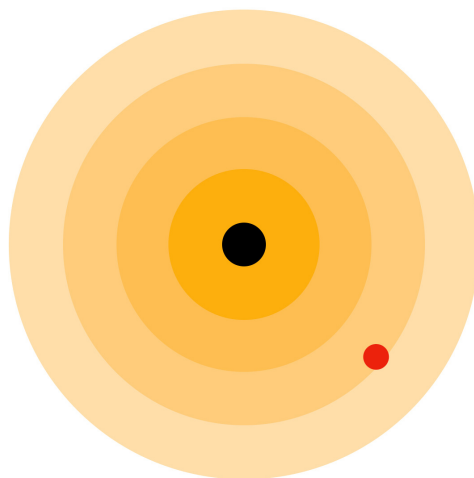


Figure 12. Proton (black circle) having Aether medium waves and electron (red circle) revolving around the proton.

For the resulting velocity of electron colliding with 3D aether medium waves, conservation of momentum and kinetic energies were applied as follows:

$$m_e \cdot u_e + m_{Aether} \cdot u_{Aether} = m_e \cdot v_e + m_{Aether} \cdot v_{Aether} \quad (8)$$

$$\frac{1}{2} m_e u_e^2 + \frac{1}{2} m_{Aether} u_{Aether}^2 = \frac{1}{2} m_e v_e^2 + \frac{1}{2} m_{Aether} v_{Aether}^2 \quad (9)$$

where m_e is the mass of an electron,

u_e is the velocity of the electron before collision

v_e is the velocity of the electron after collision

m_{Aether} is the mass of aether medium wave particles

u_{Aether} is the velocity of aether medium wave particles before collision

v_{Aether} is the velocity of aether medium wave particles after collision

The resulting equation for the velocity of electron after collision with aether waves is as follows:

$$v_e = \frac{u_e \left(1 - \frac{m_{Aether}}{m_e}\right) + 2 \frac{m_{Aether}}{m_e} u_{Aether}}{\left(1 - \frac{m_{Aether}}{m_e}\right)} \quad (10)$$

$$a_e = \frac{v_e - u_e}{\Delta t} = \frac{2 \frac{m_{Aether}}{m_e} u_{Aether}}{\left(1 - \frac{m_{Aether}}{m_e}\right) \Delta t} \quad (11)$$

where a_e is the acceleration of electron due to collision.

This acceleration of the electron due to the collision was calculated and added with acceleration acting due to electron-nucleus interactions (calculated using pseudo potentials developed in this study). In this study, the mass of aether medium wave particles (m_{Aether}) is taken as 30% of electron mass ($m_{Aether}/m_e = 0.3$). The variable wl (wavelength) is taken as 0.04777 which gives the wavelength of aether medium waves as 0.3 pm. Implementing wave interference interactions in the Hydrogen atom shows similar behaviour (supplementary material video 6) to the particle nature of the Hydrogen atom. As shown in Figure 13, the change in the radius of the atom shows a similar trend to the particle nature of electrons, as shown in the previous section.

Table 4. Properties of the Hydrogen atom for change in Aether medium wave amplitude

Aether Wave Amplitude	Average Radius (pm)	Average Potential En-ergy (10 ⁻²³ J)	Average Ki-netic Energy (10 ⁻²³ J)
0 (Particle Nature)	0.8875	314.58	121.4
0.0025	1.03	355.33	152.27
0.055	41.18	13288.73	8556.51
0.0554	239.94	77371.72	38671.99

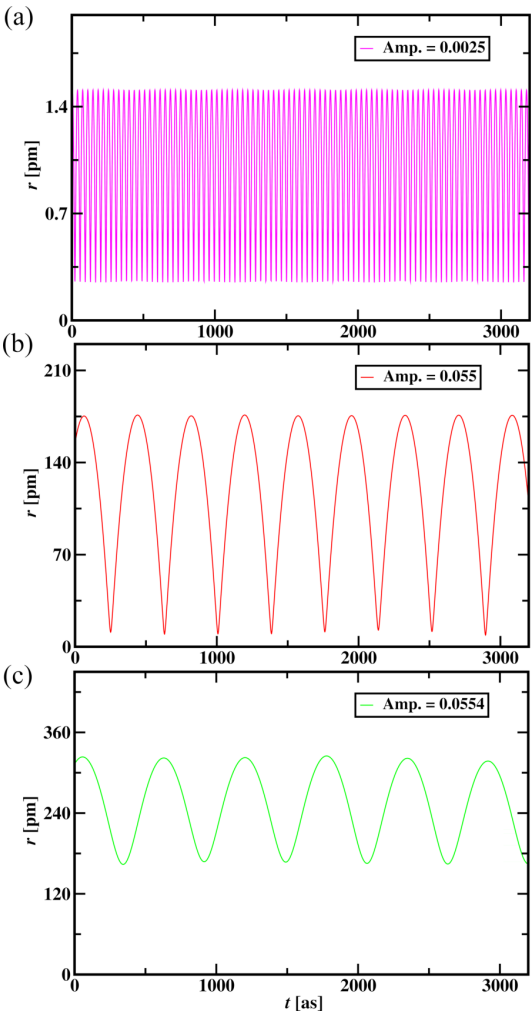


Figure 13. Radius of the Hydrogen atom versus time for the aether medium wave amplitude of 0.0025 (a), 0.055 (b), and 0.0554 (c) for 3200 time steps.

The presence of crest and trough of aether medium waves on the electron’s path was not reflected immediately but showed long-term effects on trajectories of the electron that further affect the radius of the Hydrogen atom as shown in Table 4. These wave collisions with electron show an effect in increasing its potential and kinetic energies. An increase in the amplitude of aether medium waves increases the radius of the Hydrogen atom. Also, an increase in the amplitude of aether medium waves increases the wavelength as shown in Figure 13. An increase in amplitude increases the potential and kinetic energy of the electron (Table 4). A small increase in the amplitude of 0.0004 shows a rapid increase in the radius of the Hydrogen atom and a rapid increase in the electron’s energy. Aether medium wave amplitude of 0.0025 for proton is considered further in this study as the radius of the atom is 1.03 pm, which is close to the particle nature of the Hydrogen atom of radius 0.89 pm.

By changing the wavelength, and amplitude of Aether medium waves, one can alter the electron trajectories around the nucleus. Future studies will include an investigation of these wave properties that lead electron trajectories to follow atomic orbitals around the nucleus.

3.3.2. Wave Nature of Carbon Atom Simulation

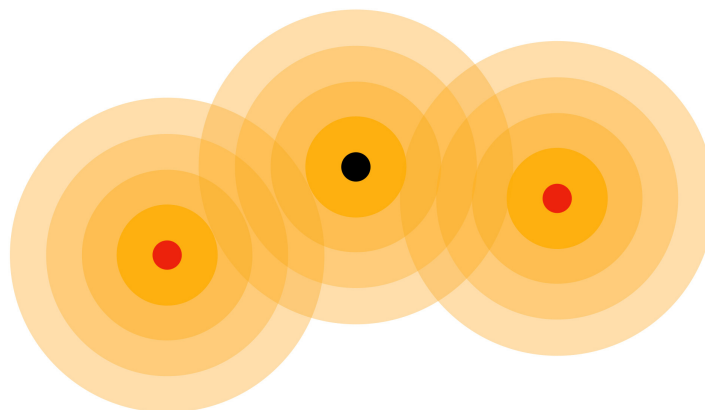


Figure 14. Proton (black circle) and electrons (red circles) have aether medium waves and communicate with interference patterns.

The electron's movement was considered by the potential acting on it and the presence of three-dimensional aether medium waves of protons, neutrons and electrons as shown in Figure 14. At each time step, past memory of the presence of aether medium waves was not considered and new aether medium waves and its interference pattern were considered. Acceleration acting on an electron is affected by the presence of protons, neutrons and other electrons which is given by summing up of equation 11 for all particles as follows:

$$a_j = \sum_{i \neq j, i=1}^{18} \frac{2 \frac{m_{Aether}}{m_e} u_{i,Aether}}{(1 - \frac{m_{Aether}}{m_e}) \Delta t} = \sum_{i \neq j, i=1}^{18} \frac{2 \frac{m_{Aether}}{m_e} A_i \cos(\frac{r_i}{wl})}{(1 - \frac{m_{Aether}}{m_e}) \Delta t} \quad (12)$$

The simulation of a carbon atom with a wave nature of electrons revolving around the nucleus is shown in the supplementary material video 7. The variables m_{Aether}/m_e (0.3) and wl (0.04777) are taken the same as the Hydrogen atom simulation. Parameterization of aether medium waves amplitude (A_i) for protons, neutrons and electrons for carbon atom was taken in two stages. The first stage is considering aether medium waves only for protons and neutrons and parameterizing wave amplitude (A_i) for them. In the second stage, aether medium waves for electrons were added and its wave amplitude (A_i) was studied. The first stage of parameterization of A_i was shown in Figure 15(a) and it shows a non-linear trend. Amplitude (A_i) for proton and neutron aether medium waves were taken as 0.0025, resulting in a minimum carbon atom radius. For the second stage parameterization, of A_i for electron aether medium waves, the carbon atom radius trend is complex as shown in Figure 15(b). Lower values of A_i showed a lower atom radius and a value of 0.00065 is taken resulting in an atom radius of 70 pm.

After parameterizing A_i based on the radius of the carbon atom for proton, neutron and electron, the radius of the carbon atom for particle nature and wave nature was shown in Table 5. Differences in particle and wave nature of Carbon atom for electron's Potential, Kinetic and Total energies are shown in Table 6 (Figure S2, S3, and S4 in the supplementary material). In both particle nature and wave nature, core electron energies are lower than valence electron energies. Electron energies (kinetic and potential energies) for particle nature carbon atom are higher than the wave nature of the Carbon atom. The differences in electron energies for particle and wave nature become less for valence electrons than for the core electrons.

Table 5. Difference in the radius of carbon atom in particle nature and wave nature

	Carbon Atom Particle Nature	Carbon Atom Wave Nature
Average Radius (pm)	71.48	70.00

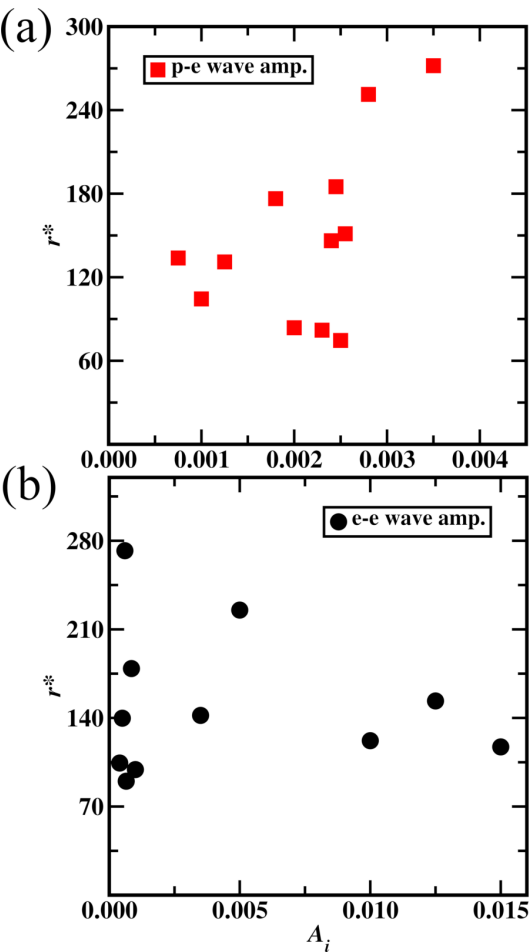


Figure 15. Radius of carbon atom variation with aether medium wave amplitude (A_i). (a) Proton and neutron aether medium wave amplitude, (b) electron aether medium wave amplitude.

Table 6. Difference in energies of carbon atom in particle nature and wave nature

	Potential Energy (10^{-23} J)		Kinetic Energy (10^{-23} J)		Total Energy (10^{-23} J)	
	Particle Nature	Wave Nature	Particle Nature	Wave Nature	Particle Nature	Wave Nature
Core Electron 1	59548.52	38095.44	73459.19	49875.38	133007.71	87970.82
Core Electron 2	78471.76	52407.55	58419.72	43061.08	136891.49	95468.63
Core Electron 3	93705.87	66293.05	49802.98	39838.54	143508.85	106131.59
Core Electron 4	109062.78	83036.78	43116.09	36906.08	152178.86	119942.86
Core Electron 5	126742.76	105527.48	36944.32	34075.01	163687.09	139602.49
Valence Electron	154750.70	150474.13	31521.12	29943.54	186271.82	180417.67

3.4. Subatomic Particle Dynamics Simulations for Molecules

Molecular dynamics simulation on subatomic particles of an oxygen molecule with two oxygen atoms along with a nucleus was developed and electron trajectories was shown in the supplementary material video 8. It was observed that the valence electrons of two atoms were shared between the two oxygen atoms. Analysis of this simulation and parameterizing pseudo potentials for molecules will be considered in future studies.

3.5. Uranium Atom Simulation

Uranium atom simulation was considered to see the advantages and limitations of this approach, as shown in the supplementary material video 9. Pseudo potentials have to be parameterized based on the Uranium atom properties and will be considered in future studies.

4. Conclusion

Monte Carlo and molecular dynamics simulations were implemented on subatomic particles to simulate a stable atom. Atom formation was observed in Monte Carlo simulation for the first 500 MC cycles. Pseudo potentials were developed to model subatomic particles interactions in stable atom simulations. The fundamental principle behind developing these pseudo potentials is that subatomic particles in an atom repel at small distances and attract each other at large distances. Pseudo potentials parameters were parameterized using carbon atom simulations and their properties. The time step for subatomic particle dynamics was calculated as 1.67 attoseconds which gives good agreement with electron dynamics in several attosecond science studies. Randomness in the Hydrogen atom for placing electron positions results in radius from 3 pm to 65.3 pm which gives good agreement with the experimental value. The Pilot-Wave theory was implemented for subatomic particles where electrons motion was guided by three-dimensional aether medium waves interference patterns. Subatomic particle simulations were carried out for Oxygen molecule that give stable molecular structure and trajectories have emerged for the electrons shared by multiple atoms.

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